

Review of the Human Health & Ecological Safety of Exposure to Recycled Tire Rubber found at Playgrounds and Synthetic Turf Fields

**Prepared for:
Rubber Manufacturers Association
Washington, DC**

**Prepared by:
ChemRisk, Inc.
Pittsburgh, PA**

July 17, 2008

Executive Summary

Increasingly, tires that reach the end of their serviceable lifetime are processed for beneficial reuse in novel applications. These include soil and surface amendments at athletic fields, playground and garden mulch, and bound surfaces at playgrounds and athletic facilities. These modern artificial surfaces reduce the likelihood of personal injury, provide uniform recreational playing surfaces, promote energy conservation, eliminate pesticide and fertilizer usage, and support waste recycling. Tires are manufactured with a variety of materials and additives to ensure optimum product safety, reliability and performance. Some tire ingredients are considered to be human health hazards at exposure levels several orders of magnitude greater than possible from contact with finished consumer products. Accordingly, athletes, parents and other stakeholders have expressed questions and concerns about the potential for adverse human health or ecological effects from the use of recycled tires in sport surface or playground materials.

The purpose of this report is to evaluate the health and ecological risks associated with the use of recycled tire rubber in consumer applications, particularly playgrounds and athletic fields. In doing so, a thorough review of available literature was conducted including studies from both advocates and opponents to the use of recycled tire materials.

An examination of the weight of evidence across all of the available studies was conducted to enable a comprehensive assessment of potential risk. As is true of all such studies, uncertainties and limitations to the health assessments that have been completed to date are recognized. However even recognizing such limitations, a review of available studies concludes that adverse health effects are not likely for children or athletes exposed to recycled tire materials found at playgrounds or athletic fields (Table 1). Similarly, no adverse ecological or environmental outcomes from field leachate are likely.

The reviewed studies considered the quantitative and qualitative aspects of exposure to classes of chemicals most likely to be inhaled, ingested or directly contacted during athletic or recreational use. While some of the ingredients used in tire manufacturing are considered potentially hazardous to human health at high doses, the potential for athlete or child exposure to these chemicals is very low. Tires are heated during manufacturing to generate physical and chemical reactions which bind the individual chemicals together such that they are inhibited from release into the environment. Studies which assessed exposure from breathing in indoor sporting environments where tire materials are used did not find appreciable adverse health effects. The same conclusion is applicable to outdoor settings, where particulate and gaseous phase air concentrations are expected to be 10 to 100 times lower, due to air dispersion and turbulence.

Uncertainties in the existing literature have been cited as areas of concern, resulting in confusion regarding the safety of recycled tire products, especially for children or other sensitive individuals. While these uncertainties, such as the lack of a temperature-emission rate relationship for outdoor ground rubber field installations or the lack of an extensive peer reviewed toxicology database for some compounds released from ground rubber from recycled tires, represent data gaps, the weight of the evidence indicates that these data gaps are not urgent or short term data needs. Although unique or significant health risks are

unlikely from use of recycled tires in sports or playing fields, research to affirm the continued safety of these products is planned and ongoing.

Based on a review of the currently available data, there are two reasonable long term research goals. Completion of these goals is not considered to be a short term or urgent data need, but would be useful in enhancing the quality of risk communication regarding play surfaces that use recycled tires. The two goals are assessment of fine particulate exposure at indoor and outdoor fields and assessment of outdoor airborne concentrations of volatile organic compounds as a function of temperature. The California Integrated Waste Management Board is currently considering completing research in each of these areas.

Of the exposure pathways and chemicals reviewed in this report, inhalation of respirable fine particulates, particularly at indoor fields, was identified as a candidate for additional characterization. Although ground rubber used in playing fields are typically one millimeter or larger in diameter, they were identified in one study as an appreciable fraction of the respirable fine particulate matter (PM_{2.5}) using a tracer molecule. Fine particulate load is expected to be low for most applications due to the processing and washing of the product which occurs during recycling. However, since adverse health outcomes are associated with fine particles, further characterization of PM_{2.5} in the raw material, as well as at indoor and outdoor fields, using a reliable tracer is recommended as a long term research objective. Although on-field outdoor PM_{2.5} levels and composition are not likely to differ from local background levels or pose a health risk, assessment of these levels is important for risk communication given the scientific consensus on adverse health outcomes associated with fine particles. If indoor spaces adhere to building codes and best practices defined by American Society of Heating, Refrigerating and Air-Conditioning Engineers (ASHRAE), no adverse health concern is expected due to PM_{2.5} levels.

The second research goal pertains to risk communication of human health inhalation risk for semi-volatile and volatile compounds. As is the case for many building materials, recycled tire rubber is not considered to be completely inert, as it contains certain volatile organic compounds that over time will diffuse into the air. These processes are slow and tend to be attenuated with time as the chemicals near the material surface are depleted. Although overall mass emission rates are quite low and diffusion limited, some stakeholders have expressed concern that solar radiation induced field temperatures as high as 160°F may cause levels of exposure exceeding that of indoor studies. While this is extremely unlikely to be the case, confirmatory field data would enhance understanding of the time-temperature emission profile, and could provide assurance that risk from inhalation of volatile organic compounds is below a level of concern.

Concerns have been expressed about ecological toxicity from zinc and the possibility of natural rubber allergy. Zinc is ubiquitous in the urban environment, and zinc leaching from artificial turf fields is not likely to pose unacceptable ecological risk. Surface water samples may easily be collected to address this issue if there are specific concerns about sensitive local species. The existing literature indicates that natural rubber sensitization or adverse allergic reactions are not likely from recycled tire materials, since liquid latex is not used in

making tires. Tires are made from natural rubber in bale form, which does not contain the same level of active proteins which may trigger allergenic responses, as found in liquid latex.

In Conclusion:

- The health and ecological risks associated with the use of ground rubber in consumer applications, particularly playgrounds and athletic fields, were evaluated through a thorough review of the literature;
- This review included studies from both advocates and opponents to the use of ground rubber;
- No adverse human health or ecological health effects are likely to result from these beneficial reuses of tire materials; and
- While these conclusions are supported by existing studies or screening risk assessments, additional research would provide useful supplemental data regarding the safety of recycled tire products and enhance the weight of evidence used in risk communication.

Table 1: Summary of Selected Human Health Assessments of Recycled Tire Rubber

Scenario	Classes of chemicals considered	Routes considered			Study Conclusions	Study Year and Citation
		Oral Ingestion or Hand-to-mouth transfer	Inhalation	Dermal		
Outdoor child playground usage	Metals, PAHs, VOCs, allergen	Literature data; simulated gastric digestion; wipe sampling of tile	--	Allergic skin sensitization based on standard guinea pig model	Acute ingestion of shreds unlikely to produce health effects; low chronic risk for hand-to-surface-to-mouth transfer; skin sensitization or reaction unlikely.	United States 2007 ⁽¹⁾
Indoor professional athlete use of artificial turf	PAHs	Literature review	Literature Review	Leaching studies; urine biomarker	No significant health risk for professional athletes; sufficient indoor ventilation recommended to control fine dust.	Netherlands 2007 ⁽²⁾
Artificial turf use	Nitrosamines	--	Air quality sampling and headspace analysis	--	Small quantities of nitrosamines emitted but not detectable in air; nitrosamine related health effects not likely.	Netherlands 2007 ⁽³⁾
Indoor artificial turf installation and amateur/professional athletic use	VOCs formaldehyde	--	Emission chamber test results paired with model small indoor gymnasium	--	Worst case indoor VOC and aldehyde concentrations do not pose a health concern for adult or child athletes; during field installation, an air exchange rate of at least 2 per hour is recommended for protection of worker health.	France 2007 ⁽⁴⁾
Indoor adult and child use of artificial turf	PAHs, PCBs, VOCs, phthalates, alkyl phenols	Ground rubber phthalate and alkyl phenol content	Indoor air quality sampling of gaseous and particulate phase compounds	Leaching studies	Chemical substances are released in very low quantities; based on worst case assumptions, use of artificial turf halls does not pose elevated risk; more information needed on natural rubber allergens.	Norway 2006 ⁽⁵⁾
Child use of public playgrounds	Organic extract of tire rubber	Genotoxicity testing			Extracts were not genotoxic and exposure potential in children deemed minimal; tire rubber at playgrounds does not pose a health hazard to children.	Canada 2003 ⁽⁶⁾

^aThe sponsors of this study have proposed a draft research plan to assess inhalation of particulates and volatile chemicals.

1.0 INTRODUCTION

A portion of tires that have reached the end of their serviceable lifetime are processed for beneficial reuse in athletic fields, playgrounds, and gardens. These include loose one to three millimeter particles used as soil and surface amendments, larger shreds for use as garden mulch, and bound surfaces at playgrounds and athletic fields. These modern artificial surfaces reduce the likelihood of personal injury, provide uniform recreational playing surfaces, promote energy conservation, eliminate pesticide and fertilizer usage and support waste recycling. Tires are manufactured with a variety of materials and additives to ensure optimum product safety, reliability and performance. Some tire ingredients are considered occupational hazards at high exposure levels. Accordingly, athletes, parents and other stakeholders have expressed questions and concerns about the potential for adverse human health or ecological effects from the use of recycled tires in sport surface or playground materials.

The purpose of this report is to evaluate the health and ecological risks associated with the use of ground rubber¹ from recycled tires in consumer applications, particularly playgrounds and athletic fields. In doing so, a thorough review of available literature was conducted including studies from both advocates and opponents to the use of recycled tire materials.

This report discusses the findings and limitations of key human health and ecological studies of ground rubber from recycled tires that have been completed to date. However even recognizing the limitations, the review of available studies concludes that adverse health effects are not likely for children or athletes exposed to recycled tire materials found at playgrounds or athletic fields (Table 1). Similarly, no adverse ecological or environmental outcomes from field leachate are likely.

The reviewed studies considered the quantitative and qualitative aspects of exposure to classes of chemicals most likely to be inhaled, ingested or directly contacted during athletic or recreational use. While some of the ingredients used in tire manufacturing are considered potentially hazardous to human health at high doses, the potential for athlete or child exposure to these chemicals is very low. Tires are heated during manufacturing to generate physical and chemical reactions which bind the individual chemicals together such that they are inhibited from release into the environment. Risk evaluations of potential exposures such as inhalation of natural rubber allergens or rubber particulate were not found in the available literature; therefore supplemental screening assessments were incorporated into the report where needed.

Various stakeholders have identified uncertainties in the existing literature as areas of concern, resulting in confusion regarding the safety of recycled tire products, especially for children or other sensitive individuals. While these uncertainties, such as the lack of a temperature-emission rate relationship for outdoor ground rubber field installations and the lack of an extensive peer reviewed toxicology database for some compounds from ground rubber from recycled tires represent data gaps, the weight of the evidence indicates that these

¹ While synthetically produced ground rubber is available, for the purposes of this report, unless otherwise noted, reference to ground rubber implies ground rubber derived from recycled tires.

data gaps are not urgent or short term data needs. Although unique or significant health risks are not likely from use of recycled tires in sports or playing fields, research to affirm the continued safety of these products is planned and ongoing, and may enable better communications on this topic.

2.0 DISPOSAL AND RECYCLING OF TIRES

The focus of this report is the use of ground rubber from ground scrap tires in sports field, running track and playground applications⁽⁷⁾. A number of methods are used to dispose of the tires discarded in the United States each year including recycling approximately 75 percent of the total disposed into useful products such as tire derived fuel (TDF), tire derived aggregate for civil engineering applications, infill for artificial turfs and as a cushioning ground cover in playgrounds⁽⁸⁻¹⁰⁾. Landfilling and tire piles have been discouraged by state and federal agencies because landfill caps can be compromised by tires rising to the surface and tire piles pose pest and fire risks, potentially requiring costly cleanups^(10, 11). Several states have implemented incentives for useful applications of waste tires including public reporting of waste tire fate in Arizona and a scrap tire recycling trust fund in Kentucky^(8, 12-14). The marketing of recycled ground rubber based products has been highly ranked in a list of environmental and economic preference for tire disposal, second only to using the tire for as long as possible before disposal⁽⁷⁾.

2.1 GROUND RUBBER PROCESSING

The recycling of used tires into ground rubber is a mature technology which requires complex machinery using either ambient - temperature or cryogenic processes. These multi-step processes result in a uniform product free of fiber or steel impurities^(7, 15, 16). For most applications, typical finished ground rubber diameters range from 0.5 to 10 mm⁽⁷⁾. Either process can be used to generate ground rubber for use as athletic field infill, with typical diameters between one to three mm⁽¹⁷⁾. In addition to inter-technology variation, there is likely to be variation in product characteristics within the same technology across various suppliers⁽¹⁸⁾.

In the ambient process, tire chips are ground by a sequence of consecutive granulators to produce ground rubber of varying size specifications with a yield of approximately 70 percent ground rubber and 30 percent steel and fiber^(7, 19). Steel and textiles are recovered using magnetic and vibration density separators. A spray or mist may be used for lubrication and to control particle generation rates. Respirable fine particles are generated during the mechanical shredding process, but are recovered to some degree in the latter stages by air pollution control devices such as cyclones or washing^(1, 15). In some applications, such as playground mats bound with polyurethane, roller mills are used to produce longer and rougher granulates which facilitate bonding⁽²⁰⁾.

In cryogenic recycling, liquid nitrogen is used to cool whole tires or chips to a temperature below -112 °F^(7, 19). At this temperature, the rubber is brittle like glass and size reduction is accomplished by crushing or breaking. Cryogenic recycling has been historically considered to result in a cleaner, less porous, and more uniform end product in fewer steps than ambient grinding, but the expense of liquid nitrogen is a consideration when comparing the two processes. As with the ambient process, steel and fibrous byproducts are recovered in the process. Because smaller size particles are more cost effective to produce than larger particles sizes, ground rubber products from cryogenic technology may have smaller nominal sizes than ground rubber products from ambient technology.

2.2 USES OF GROUND RUBBER

Ground rubber from recycled tires has a variety of uses including: rubber modified asphalt, molded products, athletic surfaces such as fields and tracks, reuse in tires/automotive products, construction, landscaping, and playgrounds^(7, 8). The benefits of ground rubber use in these applications are cost savings, improved performance, and increased safety and durability⁽⁸⁾. Ground rubber does not promote microbial growth. When used as a surface cover in playgrounds, it was shown to be more protective in preventing serious brain injury compared to pea gravel, sand and wood chips, saving an estimated \$6.6 billion per year in injury related costs^(8, 21-23). In landscaping uses, ground rubber resists compaction or decomposition over time when compared to wood mulch. Rubber modified asphalt is used on roads, highways, and bike, walking, and golf cart paths⁽⁸⁾.

Ground rubber is frequently used as infill for artificial turf athletic fields and the New York City Department of Parks and Recreation reports that artificial turf athletic fields are used 28 percent more often than a conventional sports field⁽²⁴⁾. Although the cost to install artificial turf fields can be more than conventional fields, artificial fields are estimated to have lower maintenance costs than grass fields⁽²⁴⁾. While frequency of injury does not differ between artificial and natural grass fields, the types of injuries that occur on each are very different. One study found that natural grass fields are associated with head and neural injuries, and ligament injuries whereas artificial turf fields were associated with noncontact injuries, surface and epidermal injuries, muscle trauma, and injuries at high temperature. Furthermore, natural grass field injuries generally require longer recovery times than do artificial turf field injuries⁽²⁵⁾. A separate study evaluated rotational and translational traction in rubber in-filled artificial versus natural turf fields and determined that natural grass has an increased rotational traction (often associate with more serious ligament injuries) when compared to artificial turf fields⁽²³⁾.

Some applications consist of ground rubber bound in a poured substrate, which is used at playground surfaces and running tracks⁽⁷⁾. As compared to loose rubber, it is easier to maintain and keep clean. The material is not moved or displaced during play but can have less shock absorbing potential than loose ground rubber⁽²²⁾.

3.0 RECENT PUBLIC HEALTH STATEMENTS, QUESTIONS AND CONCERNS

While the use of ground rubber in its applications provides for the recycling of scrap tires and can provide appreciable benefits over conventional materials, recent attention has focused on the possibility that ground rubber may cause an environmental or human health risk through these uses. Specific concerns are that particles of ground rubber may be inhaled or ingested; that dermal exposure may result in natural rubber allergy; or that VOCs and other chemicals such as PAHs may be emitted from ground rubber, resulting in negative impacts on human health or the environment^(6, 24). Many state and local governments, in response to public questions, have addressed the issue of the use of ground rubber in commercial applications. Included below are the conclusions and recommendations of these governing bodies.

New York State

Legislators in New York State have proposed a six-month moratorium on the installation of new synthetic turf fields until the benefits and disadvantages can be more thoroughly investigated in terms of children's health and water quality. While this is not specific to the use of ground rubber as fill in artificial turf fields, some of the concerns raised from ground rubber usage have influenced this decision⁽²⁴⁾.

New York City

New York City purchases the largest amount of synthetic turf compared to any other community in the United States⁽²⁴⁾. To address consumer concerns about the potential hazards associated with the use of artificial turf fields, the New York City Department of Health and Mental Hygiene (NYC DOHMH) published a fact sheet on artificial turf. In addition to providing information about the benefits of using artificial turf fields in comparison to natural grass fields, they address concerns regarding chemicals detected in ground rubber (PAHs, metals, VOCs), and natural rubber. The Department recognizes that while chemicals are detected in ground rubber, they are unlikely to pose a health risk based on currently available information, and furthermore are ubiquitously found in the urban environment from alternative sources. Lastly, the DOHMH refers to ongoing research to identify gaps in current knowledge regarding the health effects associated with artificial turf, but continues to recommend the use of artificial turf fields to consumers⁽²⁶⁾.

New Jersey

The New Jersey Department of Environmental Protection released a white paper reviewing the toxicity associated with the use of ground rubber from recycled tires in playgrounds and artificial turf fields. They conclude that there is "no obvious toxicological concern" associated with the intended uses of ground rubber in outdoor settings⁽²⁷⁾, while reserving conclusions about the potential for allergic reaction and natural rubber sensitization.

California

In 2007, the California Office of Environmental Health Hazard Assessment (OEHHA) released a risk assessment of the use of recycled waste tires in playgrounds and tracks with a specific focus on children as a susceptible population. This study included a thorough review of the literature related to chemical leaching from tire material and other relevant studies; an analysis of exposure and risk associated with oral ingestion of ground rubber; an analysis of

exposure via hand-to-mouth activity; an analysis of the potential for skin sensitization through dermal contact; ecotoxicity associated with recycled tire uses; and evaluation of head injuries related to different playground surfaces. The conclusions of this study indicate that there is little risk associated with exposure to recycled tire materials used in playgrounds or tracks⁽¹⁾.

Connecticut

The Connecticut Department of Public Health released a fact sheet addressing common questions regarding the health issues associated with artificial turf fields. In this fact sheet, the Department addresses chemical releases from infill material and routes of exposure. The Department suggests that, with respect to VOC emissions from turf fields, wind and temperature gradients should result in rapid dilution such that concentrations in the athlete's breathing zone are below levels of concern. Furthermore, they state that many of the chemicals emitted from the tire material are commonly found in urban and suburban environments from car exhausts, furnaces, consumer products, and foods. In conclusion, the fact sheet states that, based on current evidence, which is not without uncertainty, there is little risk to public health⁽²⁸⁾.

Concord, Massachusetts

The town of Concord Massachusetts hired an environmental engineering firm and a human health risk assessment expert to evaluate the potential human health risks associated with ground rubber in artificial turf fields. The expert reviewed literature and wrote a brief memorandum to the director of the Public Works Department in Concord. Much of the focus of the assessment was on PAHs, and the conclusions of the assessment were that there is little exposure to and thus little risk from PAHs or other chemicals associated with ground rubber used in artificial turf fields to the human population⁽²⁹⁾.

EPA Region 8

In response to Executive Order 13045, which instructs the EPA to investigate environmental or safety risks that may disproportionately affect children, and likely prompted by questions from consumers, regulators in Region 8 identified potential health hazards to children from playing on surfaces such as athletic fields that employ ground tire rubber^(13, 30, 31). EPA Region 8 representatives suggest that based on limited data and existing data gaps, the risk from the use of tire rubber at playgrounds and athletic fields is unknown with respect to pulmonary toxicity from particulate and fibers, systemic toxicity from inhalation of volatile organic compounds (VOCs) and heavy metals, and pulmonary sensitization to natural rubber. Furthermore, it is recommended that the EPA conduct a comprehensive risk assessment to include these endpoints, and initiate research to fill existing data gaps that may aid in this assessment⁽³¹⁾.

EHHI

In response to government issued statements regarding the safety of ground rubber used in consumer applications, Environment and Human Health, Inc. (EHHI), a Connecticut non-profit organization that conducts human health and environmental policy analysis, recently issued a report recommending a moratorium on the installation of fields or playgrounds that use ground-up rubber tires. These conclusions were based on limited testing which showed

that low levels of metals or organic compounds are leachable from tire rubber, extrapolation from occupational studies, and critique of relevant quantitative studies.

The concerns of EPA Region 8 and those publicized by EHHI are addressed in the literature review presented in Section 5.0 and discussion presented in Section 6.0.

4.0 OVERVIEW OF CHEMICALS USED IN TIRE MANUFACTURING

In order to understand the potential chemical risks associated with the use of ground rubber from recycled tires, it is necessary to review the tire manufacturing process, particularly in relation to the types of chemicals used, their potential for release from a finished tire and associated toxicities. In general a tire may consist of five primary components, namely: tread, sidewall, steel belts, body plies, and the bead. As such, tires are manufactured from many different materials including natural and synthetic rubber, textiles and steel. Depending on the specific function and performance of a tire, different rubber formulations based on different polymers, fillers and low molecular weight ingredients are necessary for the various tire components. The rubber components are made using chemically reactive and unreactive materials including:

Unreactive materials	Reactive materials
Polymer	Silanes (coupling agents)
Carbon black (filler)	Adhesives
Silica (filler)	Accelerators and vulcanizing agents (cross linking)
Mineral oil (plasticizer)	Sulfur (cross linking)
Resins	Stearic acid (activator)
Waxes	Retarders (cross linking)
Zinc oxide (activator)	Antioxidants
Processing aids (fatty acids, esters, glycol derivatives)	

The tire production process consists of three primary steps: preparation of the component materials, production of the components, and building of the tire (Figure 1). The majority of the chemical materials are added to the rubber mix during a step called compounding during preparation of the component materials.

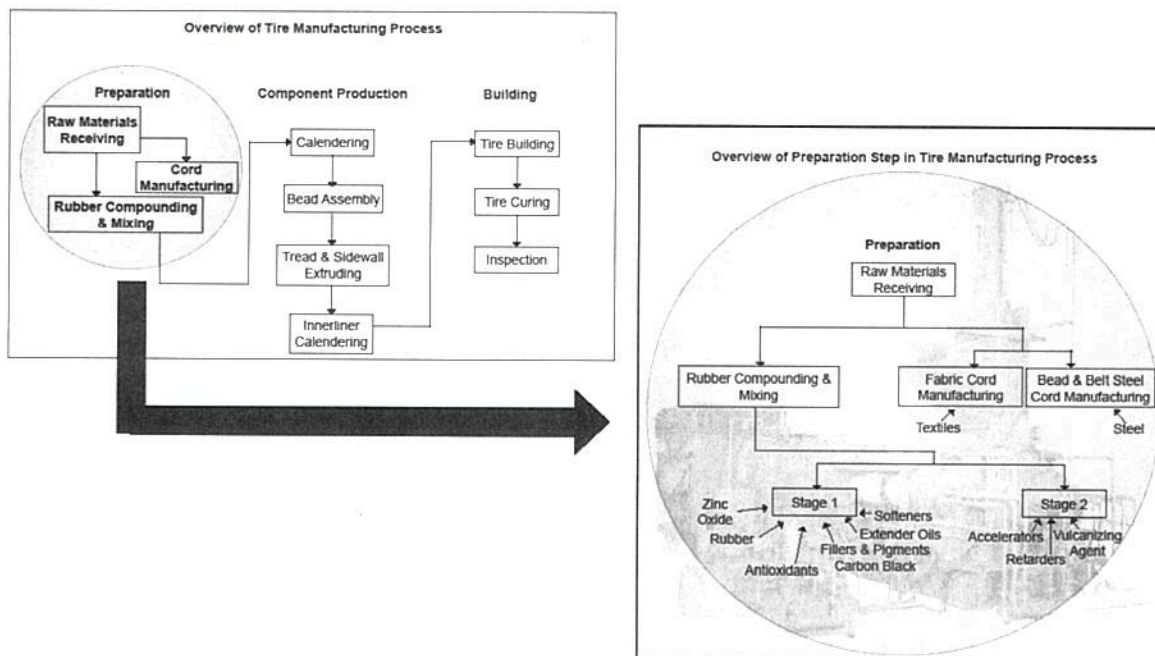


Figure 1: Overview of the Tire Manufacturing Process

During the tire making process, reactive materials are generally consumed during the curing process, so that little if any of these materials are found in the finished product^(32, 33). As such, many of the reactive chemicals which have been identified by some as a concern in tire rubber because of their classification as mutagens, carcinogens, or reproductive toxicants, are not present in the finished tire at significant concentrations. Therefore, it is incorrect to infer the toxic properties of the individual chemicals used in tire manufacturing are the same as the recycled tire ground rubber end-product.

4.1 HAZARD IDENTIFICATION IN TIRES

Even while reactive chemicals may not be available for release from end-product tires, the identification of chemicals used in tire manufacturing as mutagens, carcinogens, or reproductive toxicants has resulted in a significant amount of attention on safety from the use of recycled tires in applications such as playground surfaces and artificial turf athletic fields, particularly in light of the fact that one of the exposure populations is children. Much of the focus of this research has been on polynuclear aromatic hydrocarbons (PAHs), phthalates, and metals. Presented here is a brief summary of the classes of chemicals used in tire manufacturing that were the focus of investigators when assessing the safety of recycled tire materials used in consumer applications such as artificial turf fields and playgrounds.

4.1.1 Antioxidants

Antioxidants are added to the rubber compounding mixture to inhibit oxidative aging of the end-product rubber. Antioxidants are not consumed during vulcanization, but are consumed during product use. Common families of compounds used as antioxidants in tire manufacturing include quinolines, phenolic stabilizers and phenylenediamines. Antioxidants may be able to migrate within the vulcanized rubber, but have not been detected in leachate studies from waste tires or in highway runoff, indicating the likelihood of exposure to antioxidants from the use of recycled tire rubber in playgrounds and artificial turf fields is low⁽³²⁻³⁴⁾.

4.1.2 Accelerators and Vulcanizing Agents

Accelerators and vulcanizing agents are reactive chemicals used to promote or control the rate of sulfur vulcanization during tire curing. As reactive chemicals, they are wholly consumed during the curing process, and are not expected to be present in the end-product. Furthermore, individually these chemicals represent only a small component of the rubber compounding mixture, as they are added to the rubber compounding mixture up to one percent by weight^(32, 33). Consequently, exposure to these chemicals from the use of recycled tire rubber in playgrounds and artificial turf fields is likely to be negligible.

4.1.3 PAHs

Polynuclear aromatic hydrocarbons (PAHs) are found as impurities in aromatic extender oils which are used as plasticizers to provide elasticity and hardness to the finished tire. Therefore, recycled tires may contain PAHs⁽³⁵⁾, although recent legislation in the European Union which restricts the use of aromatic oils in tire manufacturing will result in fewer recycled tires that contain PAHs in the future⁽³⁶⁾. Some PAHs are recognized carcinogens by the International Agency for Research on Cancer (IARC), and several other regulatory bodies^(37, 38). As such, PAHs are often heavily regulated in terms of industrial emissions and clean-up levels⁽³⁹⁾. The predominant source of PAHs in the environment is fuel combustion, and on roadways, it is primarily associated with diesel fuel⁽⁴⁰⁾. Because of the perceived risk associated with PAHs, nearly all of the risk assessments evaluating the safety of ground rubber used in artificial turf fields and playgrounds have evaluated PAH exposures as an endpoint^(1, 4, 5, 35). PAHs, as a family, are also highly toxic to aquatic organisms.

4.1.4 Phthalates

Phthalates are plasticizers used at some tire manufacturing facilities to control elasticity of the end-product rubber⁽⁵⁾. The one phthalate that has received significant attention related to environmental health is di(2-ethylhexyl) phthalate (DEHP). While it may be used as a plasticizer in both synthetic and natural rubber products, it's most common use is in PVC plastics. DEHP is considered a probable human carcinogen by the U.S. EPA, although IARC concludes that the carcinogenicity of DEHP cannot be classified because the mechanism of carcinogenicity as demonstrated in rats and mice may not be relevant to humans. DEHP has also been identified as a suspected endocrine disrupter, as high acute exposures to DEHP can

induce alterations in sperm formation and fertility in both mice and rats. However, no reproductive effects have been observed at low level environmental exposures. Because of the perceived risk associated with DEHP, the detection of phthalates in ground rubber has drawn attention in relationship to the use of ground rubber in playgrounds and artificial turf fields⁽⁴¹⁾.

4.1.5 Metals

Zinc is the primary metal used in tire rubber compounding. Zinc, in the form of zinc oxide, is used as an activator of the vulcanization process^(32, 33). Zinc is an essential element to human health and is not typically regarded as a health hazard, although excessive zinc intake can result in electrolyte imbalance via interference with copper homeostasis^(42, 43). Zinc, like many other metals, has a low threshold for toxicity in aquatic species⁽⁴⁴⁾, and is therefore often the focus of leaching studies evaluating the potential for aquatic toxicity from the use of recycled tires in playground and artificial turf fields⁽⁴⁵⁾. While there are other metals found in whole tires, primarily in the steel belting of the tire, the ground rubber manufacturing process isolates and recovers these metals and therefore the recycled rubber is not a source of those metals in the environment^(1, 15).

4.1.6 Other

Because petroleum based oils containing volatile organic compounds (VOCs) are used during tire manufacturing, some VOCs may be present in end-product tires and ground rubber from recycled tires. It is expected that VOCs should off-gas from the tire after only a short time, due to high volatility, but these compounds have received significant focus in exposure and risk assessments of ground rubber uses, likely due to the toxicity associated with many VOCs (i.e. benzene and formaldehyde)^(1, 4, 5, 46, 47).

Certain proteins found in natural rubber are also detectable in small quantities in tires⁽⁴⁸⁾. Sensitization to these natural rubber proteins (i.e., natural rubber latex (NRL) proteins) through skin contact or inhalation can result in significant health hazards, such as severe allergy or asthma. Several groups have identified allergy as an endpoint of concern, based on limited information regarding natural rubber allergen concentrations in air as a result of the use of ground rubber in athletic fields and playgrounds^(1, 5).

5.0 SUMMARY OF HUMAN AND ECOLOGICAL RISK STUDIES OF RECYCLED TIRE PRODUCTS

This section provides a review of the literature associated with human health and ecological studies of useful applications of recycled tires. While the use of ground rubber is most pertinent here, findings associated with other recycled tire products (i.e. tire shreds) may also be relevant and are also discussed briefly.

5.1 IMPACTS ON HUMAN HEALTH

5.1.1 Oral Exposure to Ground rubber

Oral exposure to ground rubber or associated chemicals may occur through multiple means: ingestion of ground rubber (intentional or incidental); hand-to-mouth activity; and intake of drinking water contaminated by chemicals leached from ground rubber. The existing literature evaluating oral exposure to components of ground rubber addresses each of these issues.

5.1.1.1 Oral Ingestion of Ground rubber

Oral ingestion of ground rubber, either intentional or incidental, is unlikely to represent a major exposure pathway. However, consideration of this pathway is necessary, especially in the case of children who may consume ground rubber or pieces of poured rubber at playgrounds. The California OEHHA assessed the potential risk to children from this pathway⁽¹⁾. In the OEHHA analysis, the toxicity assessment was conducted using data from published literature of leachate from tire shreds as well as a human bioavailability study. In the first analysis using the leachate data from the literature, OEHHA conservatively assumed that the highest concentration of each chemical detected in the leachate would be available for ingestion. Additionally, a single dose estimate of individual chemical constituents from ingestion of 10g of ground rubber (in a 15 kg child) was determined based on the leaching concentrations and risk quantified using a hazard index (acute screening value/dose estimate). This approach is consistent with U.S. EPA guidance which recommends assessment of acute exposure for a pica child using an ingestion rate of 10g per day. Where no acute screening value was available, a subchronic or chronic screening value was used for comparison. Where the dose was lower than a subchronic or chronic screening level, OEHHA concluded that acute health effects were unlikely. This is a reasonable approach, as acute effects most frequently require much higher doses than do chronic effects. Of those chemicals identified to leach from tire materials, 17 were unable to be characterized in terms of risk due to either absence of a screening criteria or insufficient available information to calculate dose. Hazard indices were calculated for 24 chemicals, but only zinc exceeded a hazard index of 1.0. The hazard index for zinc was 5.167 based on an average daily intake of 1.55 mg/kg. Zinc, however, is an essential element to the diet, and has a tolerable upper intake level of 7 mg/kg for a 3-year old child⁽⁴²⁾. Furthermore, the leaching value used to estimate dose for zinc (2.3 mg/g tread) was 2.6-2,300-fold higher than results from other studies. Therefore, OEHHA concluded that the risk associated with zinc leaching is likely overestimated.

In addition to acute health risk, long term risk for developing cancer was estimated for those chemicals in the leachate that were considered carcinogens by the State of California. Those substances that were evaluated for carcinogenic risk included arsenic, cadmium, lead,² benzene, trichloroethylene, aniline, and naphthalene. Dose estimates were calculated using the same exposure assumptions as defined above (10g single exposure) but averaged over a 70 year lifespan. Considerations were made for the increased susceptibility of children to mutagenic carcinogens by multiplying cancer risk by 3, as recommended by the U.S. EPA⁽⁴⁹⁾. Total cancer risk from ingestion of ground rubber based on available leaching studies in the literature was 1.2×10^{-7} , well below the acceptable limit of 10^{-6} .

In order to more accurately predict leaching from ingestion by humans, OEHHA conducted a simulated stomach leaching study, wherein 40 grams of ground rubber were leached using a simulated gastric fluid, in order to replicate the environment of the stomach. The simulated gastric fluid was subsequently analyzed for chemical constituents. The non-cancer acute hazard indices and cancer risks were then recalculated using these leachate concentrations and the previous exposure assumptions. The non-cancer hazard index for all leachable chemicals was below 1.0, with the exception of aniline (1.062). Leaching of zinc into the gastric juices yielded a concentration of zinc nearly $1/18^{\text{th}}$ that of the estimate used for determination of risk from the tire shred leaching studies, indicating this value is an overestimate, and thus risk from zinc is likely to be very low. Of the chemicals detected in the simulated gastric leachate, five were considered carcinogens by the State of California (arsenic, cadmium, cobalt, lead, aniline) and therefore theoretical excess cancer risk estimates were made. None presented an increased risk for cancer based on the dose estimate, and the cumulative cancer risk was 3.7×10^{-8} . This is well below the acceptable risk level of 10^{-6} , as determined by the EPA, and is one-third of the estimate based on tire shred leaching values obtained from the literature.

In estimating non-cancer and cancer risk based on literature studies and the gastric leaching experiment, the OEHHA used a conservative approach in determining bioavailability of the chemical following leaching. They assumed that 100 percent of all of the chemicals were available for uptake into the systemic circulation. Therefore, it is likely that cumulative risk estimates, while low, are actually overestimates of risk associated with ingestion of ground rubber.

This study, while the best available for investigating the risk associated with ingestion of ground rubber, has been criticized by EHHI⁽⁵⁰⁾. Because some chemicals lacked criteria values for comparison, EHHI suggests the risk may actually be higher as it was not possible to assess risk from those chemicals. Furthermore, they criticize the use of an acute exposure estimate to estimate lifetime cancer risk. Recommendations for estimating soil intake in children (which is assumed to be similar to intake of rubber) suggest that children may ingest up to 10g of soil one or two days per year, a behavior expected to discontinue as the child ages⁽⁵¹⁾. Supplemental chronic risk estimates based on a child's typical incidental ingestion

² Arsenic, cadmium, and lead are not expected to be present in native tire tread based on composition, but may become entrained in the tread rubber upon contact with the road surface and are thus detectable in ground rubber from used tires.

rate of 100 mg/day, as prescribed by the U.S. EPA's Child-Specific Exposure Factors Handbook, indicate that regular exposure (e.g., regular play on ground rubber filled athletic fields) to ground rubber for the length of one's childhood does not increase risk of cancer above levels considered by the state of California to be *de minimus* (i.e. a lifetime excess cancer risk of 10^{-6}) or pose a likelihood of non-cancer effects (i.e. hazard index less than one).⁽⁵¹⁾ Consideration of additional exposure through adulthood (based on total child through adult-hood upper bound residential tenure of 30 years), indicates that chronic adverse health effects are unlikely under any scenario. (See Attachment II for risk calculations).

Incidental ingestion following inhalation of non-respirable particulate also represents a possible exposure pathway. However, as this exposure scenario is likely to result in little ingestion relative to the intentional ingestion scenario as described above, the associated risks would be appreciably lower. In addition to the detailed California assessment based on acute intake, the Norwegian health assessment concluded that chronic incidental ingestion of 0.5 to 1 g/match ground rubber containing phthalates and alkyl phenols does not pose an elevated health risk⁽⁵⁾. Therefore, the California evaluation of acute exposures was reasonably health protective for this exposure scenario.

5.1.1.2 Hand-to-Mouth Activity

In order to estimate exposure to chemicals from ground rubber via hand-to-mouth activity, a wipe sampling study was initiated by the California OEHHA. In this study, the OEHHA used rubber tiles made from recycled tire material (often ground rubber in a poured substrate). A steel weight was placed atop a wipe and dragged across the rubber tile three times along the same 12 foot path. The wipe samples were then evaluated for chemical constituents. Five chemicals were detected at levels above background: zinc, and four PAHs (chrysene, fluoranthene, phenanthrene, and pyrene). In order to estimate oral exposure via hand-to-mouth activity, several factors need to be considered: surface area of body in contact with playground surface; frequency of hand-to-playground contact; frequency of hand-to-mouth activity; efficiency of chemical transfer from hand to mouth; and frequency of playground use. Using previously established values for these variables⁽⁵¹⁻⁵⁹⁾, estimations of oral exposure via hand-to-mouth activity were derived for those five chemicals detected above background levels and risk assessed. For non-cancer effects, screening criteria values were several fold higher than ingested dose estimates, indicating a low risk from oral exposure via hand-to-mouth activity. Estimation of carcinogenic risk for those chemicals identified as carcinogens (chrysene only) resulted in a cancer risk of 2.9×10^{-6} . As a note, chrysene was only detected in the wipe survey from a playground that used a bottom layer of recycled tire and top layers of EPDM rubber (ethylene propylene diene monomer rubber). Chrysene was not found in the wipe survey of a playground surface that only used recycled tire material. Therefore, any increased risk associated with exposure to chrysene via hand-to-mouth activity at playgrounds is not attributable to the use of ground rubber from recycled tires in poured rubber applications.

5.1.1.3 Leaching into Drinking Water

While most studies evaluating the leaching of chemical constituents into water sources have focused on impact on ecological systems, a few have addressed the issue of whether leaching of recycled tire material may impact drinking water, and thus present a human health risk. Of these studies, most focus on civil engineering applications of tire material, such as use in soil absorption systems or roadside leaching fields. While the physical characteristics of the shreds used in these applications are very different from that of ground rubber, the ability of chemicals to be extracted by water is likely to remain similar, as the compositions of ground rubber and shreds are similar. Only the National Institute for Public Health and the Environment (RIVM) in the Netherlands has evaluated the potential for leaching from artificial turf fields using ground rubber infill to impact drinking water.

The RIVM study, which focused on zinc loading into water and soil from the use of ground rubber in artificial turfs, suggested that the risk to human health from zinc leaching will be negligible as concentrations of zinc in groundwater should fall well below drinking water standards for zinc⁽⁴⁵⁾. Analyses of the impact of the use of tire shreds in civil engineering applications on groundwater concentrations of metals and other contaminants have conflicting conclusions. In field studies performed by the Minnesota Department of Transportation, drinking water standards were exceeded for barium, cadmium, chromium, lead, and PAHs, where as a similar study from the Wisconsin Department of Transportation only found exceedances for lead and barium^(60, 61). However, these studies have been criticized for not maintaining proper controls⁽⁶²⁾. A well controlled study from the University of Maine indicated that primary drinking water standards (health protective) for metals were not exceeded due to the use of tire shreds, while secondary standards for iron and manganese (based on aesthetics) were exceeded⁽⁶³⁾. Humphrey et al. were unable to detect VOCs and SVOCs in groundwater below tire shred applications, and thus concluded that tire shreds have a negligible impact on groundwater quality at neutral pH⁽⁶⁴⁾. Based on these studies, it is unlikely that leaching of recycled tire material will represent a health risk for humans from ingestion of drinking water due to use in athletic fields, civil engineering applications, or other applications.

5.1.1.4 Other potentially Relevant Studies

A study conducted by the Danish Ministry of the Environment assessed health risks associated with play in sandpits lined with used tires⁽³⁵⁾. Migration studies were performed to determine what chemicals moved from the tire rubber into the sand, and thus were available for intake through ingestion of sand by children. Several PAHs and phenylenediamines (used as antioxidants in tires) were detected in the sand, although it was noted that the PAH profile was not identical to that in the tires and was considered to originate from atmospheric deposition from alternative sources of PAHs. Nonetheless, a risk assessment (using MOS approach) was conducted based on ingestion of 10g of sand, five times a week for half a year. It was assumed that 100 percent of the substances in the sand were able to be absorbed into the body upon ingestion. Margins of safety for ingestion from

chemicals detected in the sand (fluoranthene; 6PPD; IPPD; pyrene; benzo(a)pyrene;) ranged from 10,000 to greater than 1,000,000, indicating there is a very low likelihood of risk to children from ingestion of sand in tire-lined sand boxes. While this study evaluates the health risk associated with whole tires used in playground applications, it is not without relevance when understanding the risks associated with ground rubber from recycled tires used in playground applications. When normalized by surface area, both whole tires and ground rubber will contain similar chemical profiles, and thus migration of these chemicals from the rubber matrix into sand or other surrounding media (e.g., soil) would be similar.

5.1.1.5 Conclusions About Oral Studies

Collectively, studies evaluating endpoints in both children and adults indicate that there is low risk associated with the use of recycled tires in playgrounds or athletic fields to human health from oral exposure pathways. Such pathways include incidental or intentional ingestion of ground rubber, hand-to-mouth activity in children following contact with rubberized surfaces, and drinking of contaminated water. Other relevant studies evaluating safety associated with alternative tire uses in playgrounds supports this conclusion.

5.1.2 Inhalation Exposure to Ground rubber

Another potential pathway for exposure to ground rubber is inhalation, including chemicals off-gassing from the surfaces (playground, artificial turf fields, etc.) and inhalation of particulate matter (and subsequent chemical exposure via interaction in lung) entrained in the ground rubber product.

5.1.2.1 VOCs

As suggested by EHHI, one of the primary concerns associated with the use of ground rubber is the potential for volatile organic compounds (VOCs), and possibly semi-volatile organic compounds (SVOCs) to off-gas, especially with the high temperatures that rubber-containing surfaces can achieve in outdoor environments⁽⁵⁰⁾. In support of this argument, EHHI cites a study conducted by the Connecticut Agricultural Experimental Station that evaluated the chemical composition of the head space above 0.25g of ground rubber in a 2 mL bottle heated to 60 °C.⁽⁴⁶⁾ Four chemicals were identified in the headspace: benzothiazole, butylated hydroxyanisole, n-hexadecane, and 4-t-octyl-phenol.

EHHI suggests that much of the toxicity data regarding these four chemicals is lacking. Furthermore, they identified butylated hydroxyanisole as a recognized carcinogen⁽⁵⁰⁾. Butylated hydroxyanisole is recognized as a carcinogen by the NIH and IARC^(38, 65). However, of those carcinogenicity studies performed in animals with butylated hydroxyanisole, only those utilizing an oral treatment regimen have resulted in tumor formation. Tumors are limited to the stomach or forestomach⁽⁶⁵⁾. Route of administration is often an important consideration in mechanism of carcinogenesis (e.g. chromium), and in the absence of a connection between inhalation of butylated hydroxyanisole and cancer, the implication by EHHI that the detection of butylated hydroxyanisole in the head-space above ground rubber presents a cancer risk is unfounded.

The study conducted by the Connecticut Agricultural Experimental Station lacks a defined relationship between the findings and exposure to a human receptor population. In defining risk from inhalation (or any other exposure pathway), it is necessary to base risk estimates on likely air concentrations in an exposure scenario (such as at a playground or athletic field that uses ground rubber). A few organizations have measured air concentrations of VOCs under real-world conditions in order to more accurately predict risk based on off-gassing from ground rubber. Both the Environmental French Agency (ADEME) and the Norwegian Institute of Public Health and the Radium Hospital ("Norwegian study") investigated VOC concentrations above artificial turf fields^(4, 5).

In the ADEME study, miniaturized artificial turf fields were built and maintained at a constant temperature (23 ± 2 °C)⁽⁴⁾. Samples were collected in the airspace at day 0, day 1, day 3, and day 28 and analyzed for VOCs and aldehyde emissions (including formaldehyde). Total VOCs at day 0 were approximately $1600 \mu\text{g}/\text{m}^3$, but decreased to $134 \mu\text{g}/\text{m}^3$ by day 28, indicating an appreciable decrease in total VOCs over time. This data was subsequently used in an exposure assessment which modeled exposure during field installation or athletic activity on the indoor field. Of the 112 substances identified in the samples, quantitative exposure estimates and health risks were calculated for 16 (based on available toxicity criteria). Four population groups were identified (workers installing surfaces; professional athletes/coaches; amateur athletes; spectators at sporting events) and both acute and chronic exposure scenarios considered. The authors concluded that based on these exposure scenarios, VOC and aldehyde emissions from artificial turf floorings do not pose a health risk in any of the exposure groups, with the exception of workers installing artificial surfaces in small and poorly ventilated areas.

In the Norwegian study, air samples were collected at three indoor artificial turf fields, two of which (Manglerud and Valhall) used recycled tire rubber and SBR rubber, respectively, for infill⁽⁵⁾. In Manglerud, 234 chemical compounds were detected, of which 29 were able to be identified. Total VOC concentration was $716 \mu\text{g}/\text{m}^3$. During a second sampling period, total VOC concentrations were $230 \mu\text{g}/\text{m}^3$. In Valhall, mean total VOC concentrations were $234 \mu\text{g}/\text{m}^3$. In estimating risk, VOC concentrations from Valhall were used in order to establish a worst-case scenario, as chemical concentrations at this location were consistently two to three times higher when compared to the other locations. Exposure estimates and risk were calculated for four exposure scenarios: adults, juniors, older children, and children using the facility for training. Risk from acute exposure was determined to be negligible. While risk cannot be estimated based on total VOCs, risk can be determined for specified VOCs with toxicity criteria for inhalation (toluene; benzene; benzoic acid; xylenes; styrene; formaldehyde; limonene; benzothiazole). Margins of safety based on non-cancer NOAELs for all of these chemicals, with the exception of formaldehyde, exceed 100, and in most cases are greater than 10,000 for all exposure scenarios. Only benzene was considered for carcinogenic risk, although that too was within the range of acceptable risk.

While the authors from both of these studies consider the indoor scenario a "worst case" scenario, neither of these studies considered temperature variation in the field. In fact, in the

ADEME study, a temperature-controlled scenario was employed. As such, conclusions from these studies have been criticized by EHHL. Volatilization of chemicals is a temperature-dependent process, and surface temperatures at outdoor fields may reach as high as 160 °F. However, surface temperatures of this magnitude are not particularly remarkable as asphalt, which is another common surface used for recreational purposes such as basketball courts, also achieves similar maximum temperatures⁽⁶⁶⁾. Although chemical emission rates increase with temperature, the increase in volatile organic emissions from rubber is much less than that implied by theoretical vapor pressure relationships. The reason for the discrepancy is that as the ground rubber surface is depleted of VOCs, subsequent emissions are limited by the slow rate of chemical diffusion to the surface of the rubber. This process is much less dependent upon temperature than solid to vapor phase partitioning equilibrium. For example, over a temperature range of 67 to 160 °F, the vapor pressure of benzothiazole increases by a factor of almost 40⁽⁶⁷⁾. However, based on a study of a synthetic rubber athletic track, total VOC emissions are estimated to increase by a factor of only 2 over the same range based on a curve fitted to field flux chamber measurements⁽⁶⁸⁾.

The rolled sheets used in the synthetic rubber track are expected to have a similar temperature-emission profile as ground rubber. Therefore, dilution in the outdoor environment as well as source depletion from the surface of the ground rubber appreciably reduces the likelihood of VOC emissions posing a hazard. The Norwegian and ADEME indoor studies are clearly representative of worst case inhalation exposure concentrations, as the increased dilution outdoors is expected to be many times more important than the increase in emission rates with temperature. At one of the Norwegian fields, it was specifically noted that natural ventilation (i.e. open windows and hatches) was employed. Had mechanical ventilation been employed, it is likely exposure concentrations would have been lower. Another important observation is that outdoor emission rates are expected to decrease appreciably with age of the field due to surface depletion of the volatile chemicals, as shown in the synthetic rubber track and ADEME studies. A research plan including assessment of surface temperature being considered by the California Integrated Waste Management Board will be helpful in confirming this conclusion⁽⁶⁹⁾.

Total VOC levels detected in these two studies fall within the range of other indoor air spaces. In a study investigating VOC concentrations in 750 homes in the United States, Wallace et al. detected total VOC levels exceeding 1000 µg/m³ in more than half of the homes.⁽⁷⁰⁾ Similar findings were reported in a study measuring total VOCs in newly manufactured and site-built homes in the U.S.⁽⁷¹⁾. A similar study of home environments in Germany detected a geometric mean total VOC concentration of 584 µg/m³⁽⁷²⁾. However, in the Norwegian study, it was concluded that rubber granulate was an important contributor to the total VOCs in the hall⁽⁵⁾. Therefore, while the total VOC levels in these buildings may be comparable to other indoor environments, the chemical makeup of the VOC mixture is likely to be different. Furthermore, sports arenas, such as those evaluated in this study, are subject to more demanding requirements for ventilation than are homes, and comparisons to homes or other indoor air spaces may not be appropriate.

To date there have been no studies evaluating VOC emissions from outdoor turf fields or playgrounds using ground rubber surfaces. However, the State of California, in their next

phase of research, plans to measure VOC concentrations above outdoor turf fields and the influence of ambient temperature on these concentrations⁽⁶⁹⁾. Chang et al. measured emissions of VOCs at breathing height from athletic tracks made of synthetic rubber, and evaluated impact of temperature and aging on VOC emissions⁽⁶⁸⁾. Hexanal, 2-methyl furan, toluene + octane, and methyl isobutyl ketone (MIBK) were the dominant compounds emitted from the synthetic rubber track. MIBK was unique to the synthetic rubber track, in comparison to those tracks using polyurethane based surfaces. With aging of the track, VOC emissions decreased. Emissions did not vary substantially by temperature, especially in comparison to track age. While the rubberized surface in this study is not ground rubber (although it may be made of poured ground rubber), this is the only study available which evaluates outdoor VOC concentrations associated with synthetic rubber athletic surfaces and may provide a useful surrogate for understanding the emissions from athletic surfaces and playgrounds in the absence of data from these applications of ground rubber. No exposure estimates or risk calculations were determined based on results from this study. However, total VOC concentration at breathing height above the track was $0.39 \mu\text{g}/\text{m}^3$. This is several orders of magnitude lower than detected in the indoor scenario, which based on the exposure scenarios used in the ADEME and Norwegian studies, did not pose any risk to human health.

In summary, VOC emissions from rubberized surfaces in athletic fields or playgrounds are unlikely to pose a human health risk based on the available data. The authors of the Norwegian study note that absence of toxicity criteria for some of the chemicals detected does not mean these chemicals cannot constitute a health risk, but that rather, based on currently available data, no cause for concern based on VOC emissions exists.

5.1.2.2 Particulate

Particulate matter, including airborne dust, is generated in all indoor and outdoor environments from a variety of sources such as agriculture, power plants, industrial facilities, on-road and off-road vehicles and forest fires⁽⁷³⁾. Particulate matter is a complex mixture of solid inorganic and semivolatile organic chemicals and aqueous materials and is found in a range of sizes described by an aerodynamic diameter. Examples of particulate matter are soot, smoke, elemental and organic carbon, nitrates, sulfates, acids, bacteria, fungi, spores, pollen, dust, and tire wear materials. Fine particles less than one to three μm in diameter generally originate from combustion sources or precursor gases whereas larger coarse particles are considered primary particles emitted directly from specific sources⁽⁷³⁾. Fine particles are generally not derived from primary particles due to the amount of energy that would be required to generate such fine particles from larger pieces of rubber. In each environment the levels of particulates are influenced by the level of air dispersion or ventilation, the rate of particle release or suspension and the physical configuration of the space or area. With regard to potential human risk, health scientists assess both the bulk physical characteristics of the particles (i.e. total mass, surface roughness and geometry of inhaled particulate) as well as the particulate phase chemical composition (i.e. concentrations of individual chemicals).

Although validated relationships between specific sources of particulate matter in outdoor ambient air and health outcomes are not available, long term exposure to fine and coarse

particulate matter is associated with death in older adults with cardiopulmonary disease⁽⁷³⁻⁷⁵⁾. However, the mechanisms relating the characteristics of particulate matter to specific health effects are poorly understood. Research suggests that chemical composition is a minor contributor to particulate matter toxicity because similar dose-response relationships are observed across the world and a wide range of particulate compositions⁽⁷⁶⁾. Proposed fine particle respiratory damage mechanisms include penetration and accumulation in the interstitial spaces of the lungs, tissue damage by aggressive chemicals such as acids and catalytic effects and oxidant formation attributable to trace metals within the lungs^(77, 78). Consistent with this research, systemic toxicity (i.e. whole body) attributed to trace inorganic or organic compounds found within particulate matter is expected to be low. In outdoor settings, the U.S. EPA generally considers evaluations of the soil direct contact pathway to be protective of fugitive dust inhalation exposures, as soil screening levels are typically several orders of magnitude lower (i.e. more stringent) for the oral route versus the inhalation rate⁽⁷⁹⁾. With the exception of hexavalent chromium, routine evaluation of residential or commercial/industrial fugitive dust exposure is not recommended unless unusual heavy truck traffic or annual average wind speeds well above national averages are expected. Therefore, individual chemical risks attributable to airborne ground rubber are expected to be low. With respect to ground rubber recreational field installations, limited airborne particulate data are available, but upper bound total mass and individual chemical particulate exposures can be assessed using data collected at indoor Norwegian artificial turf fields (addressed below)⁽⁵⁾.

5.1.2.2.1 Total Respirable Airborne Particulate Exposure

Two characteristics of ground rubber are likely to limit the magnitude of fine particle or airborne dust release and subsequent exposure. First, during rubber recycling, fiber and dust removal is typically accomplished using air classifiers or other equipment^(7, 15, 16). Second, foot traffic is unlikely to generate appreciable quantities of new particulates during field use due to the high amount of energy that would be required to generate small respirable particles⁽¹³⁾. However, it is unknown the degree to which coarse and fine particles created or entrained are removed in processing of recycled rubber. In two ambient scrap-tire shredding facilities located in France and Taiwan, ambient levels of respirable dust were 230 to 1250 $\mu\text{g}/\text{m}^3$ indicating the potential for particle generation during processing⁽¹⁵⁾. Regardless of the underlying particulate content of the ground rubber, turbulent air dispersion in outdoor settings and precipitation wash-off are expected to appreciably attenuate on-field particulate concentrations relative to indoor settings. For settled dust, the two primary resuspension processes in air are abrasion of surfaces by applied mechanical force by foot traffic, wheels or other implements and dust particle entrainment by turbulent air currents at high wind speeds (i.e. greater than 12 mph)⁽⁸⁰⁾. Based on a review of the literature and a simple screening calculation, the primary resuspension process for ground rubber particles used in fields or playgrounds appears to be surface disturbance by foot traffic^(5, 80).

In air, total suspended particulate matter (PM) is defined as aggregated molecules or particles which typically range in aerodynamic diameter from 0.01 to 100 μm (one micrometer is one millionth of a meter). For health assessment, the operational definition (i.e. indicator) of particulate matter is typically based on the cut-point of 50 percent collection efficiency for a sampler that contains a size fractionator. The common metrics include PM_{10} based on an

aerodynamic diameter cutpoint of 10 μm , $\text{PM}_{2.5}$ based on an aerodynamic cutpoint diameter of 2.5 μm and $\text{PM}_{10-2.5}$ representing the difference between the two size fractions.

Particles are considered to be thoracic if they penetrate anywhere within the lung airways or gas exchange region, whereas particles are considered to be respirable if they deposit exclusively in the gas-exchange or pulmonary region of the deep lung. Particles greater than 100 μm are too large to remain suspended in air, whereas particles larger than 10 μm are not considered to be respirable, as they are not deposited on the non-ciliated portion of the lungs. Particles less than 10 μm are characterized by slow gravitational settling velocities which in the presence of air turbulence impede the rate of settling back to the ground after initial release. Therefore, pulmonary risk is primarily attributed to particles with characteristic aerodynamic diameters less than 10 μm . Particles with diameters between 2.5 and 10 μm accumulate in the lung and are considered coarse particles and regulated in the United States based on acute risk⁽⁸¹⁾. Particles with diameters less than 2.5 μm , or $\text{PM}_{2.5}$, are considered fine particles, and considered to pose greater health risk than PM_{10} due to their ability to penetrate deeper into the lung and are regulated based on both chronic and acute health risk⁽⁸¹⁾. Epidemiological studies have shown associations between ambient particulate concentrations and adverse health indicators such as increased mortality and chronic respiratory disease or secondary cardiovascular effects⁽⁷³⁻⁷⁵⁾.

For open sources, such as dirt roads or playing fields, fugitive dust is generated when mechanical disturbances suspend granular material exposed to the atmosphere⁽⁸⁰⁾. Data regarding outdoor emissions of particulate from ground rubber playing surfaces was not identified in the literature. However, using a simple “unlimited reservoir” model which assumes that wind erosion suspends an unlimited reservoir of erodible particles from an unobstructed open field or playing surface with a nominal grain diameter of 3-mm (Attachment I), the estimated PM_{10} and $\text{PM}_{2.5}$ concentrations from wind erosion are unlikely to exceed 0.1 $\mu\text{g}/\text{m}^3$ ^(79, 82).

In contrast to the low particulate levels generated by wind erosion, the authors of a study of three indoor Norwegian turf halls concluded that fine particulate associated with ground rubber with a nominal diameter of approximately 3 to 4 mm may be readily suspended by regular field use⁽⁵⁾. The study assessed two fields constructed with ground rubber infill derived from recycled tires including a newly installed field and a field approximately one year old. The source of the airborne particulate is likely to have been resuspension of existing fine and coarse particles by the mechanical force generated by use and maintenance of the field. In contrast to outdoor settings, air dispersion and dilution in indoor settings is limited by mechanical ventilation rates or natural ventilation induced by infiltration or open doors and windows. Additionally, particle washoff by precipitation is likely to reduce outdoor particle levels on the field over time. Therefore, particulate levels of ground rubber caused by disturbance of the field are likely to be on average at least an order of magnitude lower in outdoor settings.

Measured air concentrations of PM_{10} in the indoor fields containing ground rubber from recycled tire treads ranged from 30 to 40 $\mu\text{g}/\text{m}^3$ and $\text{PM}_{2.5}$ concentrations ranged from 17 to 18 $\mu\text{g}/\text{m}^3$. Based on the use of N-cyclohexyl-2-benzothiazolamine (NCBA) as a marker, the

portion of PM_{10} specific to ground rubber was approximately $9 \mu\text{g}/\text{m}^3$, or 23 to 30 percent of total PM_{10} . For the $PM_{2.5}$ fraction, the concentration attributable to ground rubber was 7 to $9 \mu\text{g}/\text{m}^3$ or 35 to 50 percent of total $PM_{2.5}$. The total indoor particulate concentrations were similar to levels measured in other urban indoor settings. For example, a recent survey of 17 Los Angeles homes and residents found indoor and personal $PM_{2.5}$ concentrations of 17.6 and $17.7 \mu\text{g}/\text{m}^3$, respectively⁽⁸³⁾. A survey of particulate levels in offices and schools found geometric mean $PM_{2.5}$ concentrations of $8 \mu\text{g}/\text{m}^3$ (offices) to $13 \mu\text{g}/\text{m}^3$ (schools)⁽⁸⁴⁾. Geometric mean PM_{10} concentrations were between $12 \mu\text{g}/\text{m}^3$ (offices) to $46 \mu\text{g}/\text{m}^3$ (schools).

NCBA is an impurity of one of the most frequently used accelerators in the tire industry and is considered have reasonable reproducibility when used as a marker but its quantification requires a complex analytical method^(85, 86). In comparison to the indoor fields, 7.5 percent of PM_{10} at an urban Switzerland curb side sampling location was attributed to tire wear particles using the same NCBA marker⁽⁸⁷⁾. The fraction of PM_{10} attributed to tire wear particles was 2 percent at an urban background site. The levels of PM_{10} attributable to ground rubber measured at the Norwegian fields appear to be similar in magnitude levels attributed in ambient air near roadways or tunnels. Based on a variety of markers, typical ambient tire wear particle concentrations of PM_{10} or total suspended particulate are $2\text{-}5 \mu\text{g}/\text{m}^3$ for roadways and $10\text{-}20 \mu\text{g}/\text{m}^3$ for tunnels.⁽⁸⁷⁾ Research to date has shown a highly variable distribution between fine ($< 2.5 \mu\text{m}$) and coarse ($> 7 \mu\text{m}$) in airborne roadside tire wear particles⁽⁸⁷⁾.

The U.S. EPA has established standards for PM_{10} and $PM_{2.5}$ which are protective of human health including sensitive subpopulations such as children⁽⁸⁸⁾. With regard to assessing indoor air quality, the American Society of Heating Refrigeration and Air-Conditioning Engineers standards (ASHRAE 62.1-2007) adopt the U.S. EPA's National Ambient Air Quality Standards as one of the appropriate evaluation metrics, including the PM_{10} and $PM_{2.5}$ standards. The PM_{10} standard is a 24-hour average of $150 \mu\text{g}/\text{m}^3$ and the corresponding 24-hour $PM_{2.5}$ standard is $35 \mu\text{g}/\text{m}^3$. In addition, an annual average $PM_{2.5}$ standard of $15 \mu\text{g}/\text{m}^3$ has been established (3-year average of weighted annual mean). The particulate levels measured at these fields are specific to the ventilation conditions and pre-existing fine particle content of the ground rubber. Although detailed information regarding ventilation was not provided, the authors indicated that ventilation was induced by opening 8 roof hatches and 16 windows at one of the fields. Based on the observed maximum ground rubber PM_{10} and $PM_{2.5}$ concentrations of $9 \mu\text{g}/\text{m}^3$, indoor installations of ground rubber are unlikely to result in exceedances of the 24-hour EPA standard for $PM_{2.5}$ and PM_{10} when fields are ventilated in accordance with recommended design standards and background outdoor ambient particulate concentrations comply with the standard.

For short term exposure such as athletic field usage, the 24-hour $PM_{2.5}$ standard is the best metric by which to assess potential health effects. However, in order to also qualitatively evaluate the chronic $PM_{2.5}$ exposure, an annual average $PM_{2.5}$ exposure concentration was calculated based on the maximum portion of indoor $PM_{2.5}$ attributable to ground rubber of $9 \mu\text{g}/\text{m}^3$ and adjustments to account for less than 24 hour exposure time and higher inhalation rate during vigorous activity. The exposure time adjustment is based on worst case

assumption of indoor field use for 2 hours per day, 5 days per week for 25 weeks per year or $(2 \text{ hours} \times 5 \text{ days/week} \times 25 \text{ weeks/year}) / (24 \text{ hours} \times 7 \text{ days/week} \times 50 \text{ weeks/year}) = 0.028$. The inhalation rate adjustment factor accounts for higher inhalation rate during field usage and was set equal to the heavy activity adult short term inhalation rate of $3.2 \text{ m}^3/\text{hour}$ divided by the average male and female long term inhalation rate of $0.55 \text{ m}^3/\text{hour}$, equal to 5.8. Therefore, the worst case adjusted annual average $\text{PM}_{2.5}$ concentration attributable to field use would be $9 \text{ }\mu\text{g}/\text{m}^3 \times 0.028 \text{ annual field hours/total annual hours} \times 5.8 \text{ vigorous activity inhalation rate/long term inhalation rate}$, or $1.5 \text{ }\mu\text{g}/\text{m}^3$, or 10 percent of the chronic U.S. $\text{PM}_{2.5}$ standard. Accordingly, indoor field containing ground rubber are unlikely to result in personal exposure exceedances of the annual average $\text{PM}_{2.5}$ standard when fields are ventilated in accordance with recommended design standards and background outdoor ambient particulate concentrations comply with the standard. Particulate exposure to ground rubber at outdoor fields is expected to be at least an order of magnitude lower than indoor settings.

Although the levels of fine and course particulate generated at ground rubber fields is not likely to pose a health concern, more study is required to evaluate outdoor fields and to assess variability in particulate generation rates between indoor and outdoor fields. Given that $\text{PM}_{2.5}$ and PM_{10} are ubiquitous in the atmosphere, the characterization of background levels and use of rubber tracer molecules to assess the fraction of particulate matter generated from the field are key considerations for future studies. Additionally, for fields situated near high density traffic areas, an important consideration is the rubber contribution from tire wear particles versus ground rubber from the field. On April 22, 2008, the California Integrated Waste Management Board presented a draft scope of work to perform volatile organic chemical and inhalable rubber particulate ($\text{PM}_{2.5}$) sampling to assess exposure potential and study the influence of temperature and intensity of field disturbance⁽⁶⁹⁾. This study will provide useful supplemental data on particulate concentrations and human exposure for fields using ground rubber.

5.1.2.2.2 *Particulate Phase Chemical Exposure*

For ambient conditions, particulate phase chemical exposures are typically low, with potential human risk several orders of magnitude lower than potential incidental oral ingestion or dermal risk⁽⁷⁹⁾. The authors of the Norwegian artificial field study assessed several particulate exposure scenarios including adults, juniors and children. Dose was calculated based on concentration in the rubber granulate and the PM_{10} concentration (PCBs, PAHs, phthalates, alkyl phenols) and based on the measured air concentration (PAHs and phthalates). As expected based on past experience, daily chemical uptakes were low. For example, a worst case daily phthalate uptake of 47,000 pg/kg resulted in child and adult scenario margins of safety of 23,000 to 80,000. For all chemical classes assessed, it was concluded that the chemical compounds present did not pose an elevated health risk.

In addition to the compounds qualitatively assessed, ground rubber dusts contain a complex mixture of various inorganic and organic compounds, such as benzothiazoles, aromatic amines and unidentified compounds. Therefore, the potential respiratory risk of the entire ground rubber particle must also be assessed. One relevant animal study has been identified

where rubber tread particulate mammalian toxicity was assessed using rats exposed in a whole body inhalation chamber⁽⁸⁹⁾. This study was originally published in Japanese but has been translated to English. Four groups of animals were exposed to 0, 100, 300 and 1,000 $\mu\text{g}/\text{m}^3$ of particulate generated from studded tire tread and was sacrificed at one or 1.5 years. Dose dependent accumulation of rubber particulate in the lungs and lymph node was observed with increasing concentrations of silica and aluminum. No relationship was observed between animal survival and concentration or duration of exposure. Mild fibrosis was observed in the lungs at the 1000 $\mu\text{g}/\text{m}^3$ dose group, but not in the controls. Exposure related tumors were not observed.

The rodent no observed effect level (NOAEL) for this study was 300 $\mu\text{g}/\text{m}^3$. In order to estimate a margin of safety based on the NOAEL from this study, a human equivalent concentration must be extrapolated from the rat-based NOAEL in order to account for differences in respiratory system dynamics between humans and rats. Using the RDDR program provided with the RfC guidelines developed by the EPA, a dosimetric adjustment factor of 1.5 was determined to convert the rat NOAEL for the particle distribution used in this study (MMAD = 2.3 μm , $\sigma_g=1.6$)^(89, 90). Therefore, the human equivalent chronic NOAEL based on the results of this study is 450 $\mu\text{g}/\text{m}^3$.

Based on the quantitative chemical assessment described above, the risk of systemic toxicity is expected to be low. The maximum concentration of rubber PM_{2.5} and PM₁₀ of 9 $\mu\text{g}/\text{m}^3$ observed in the indoor Norwegian study is a factor of 50-times lower than the human equivalent NOAEL based on the mammalian study. Using the same upper bound exposure parameters from above, including the heavy activity inhalation rate and indoor field use for 2 hours per day, 5 days per week for 25 weeks per year, the corresponding margin of exposure, or human NOAEL divided by dose, is 460. Margins of exposure greater than 100 typically indicate an acceptable level of exposure and low level of health concern⁽⁹¹⁾. It should be noted there are some limitations to this study including a lack of systemic toxicity evaluation and use of particulate generated only from snow tires. Another uncertainty of this approach is the use of rodent study based on tire wear particulate as a surrogate for exposure to trace particulate present in bulk recycled tire rubber. Additional *in vivo* (i.e. living organism) whole particle animal toxicity tests would be useful in supplementing the findings of this study.

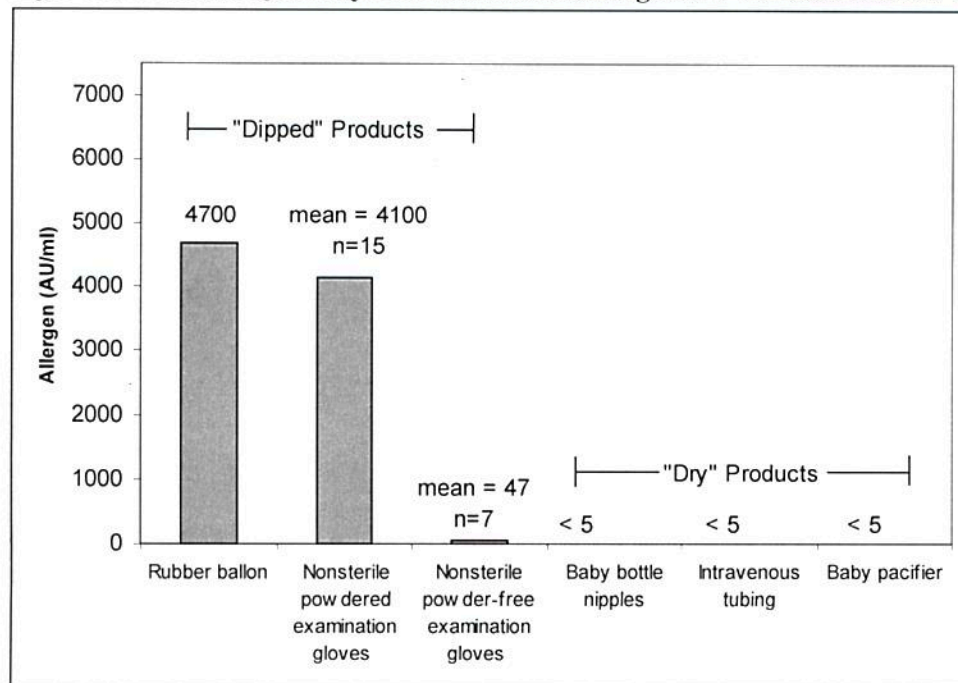
5.1.2.2.3 *Natural Rubber Allergy and Asthma from Particulate Inhalation*

One of the concerns with regard to exposure to rubber-containing particulate is the risk for the development of natural rubber allergy and associated asthma. Natural rubber contains proteins thought to induce allergy or hypersensitivity to natural rubber-containing products^(92, 93). Therefore, there is a concern due to the severity of natural rubber allergies that exposure to recycled tire material may lead to natural rubber allergy, and in the case of inhaled natural rubber-containing particles, asthma. While a recent publication indicates that exposure to particles in ambient air (from traffic sources) does not pose an asthma risk from exposure to natural rubber associated proteins from tire tread⁽⁹⁴⁾, this question has not been formally addressed with respect to the use of ground rubber from whole tires. Approximately 20

percent of tire treads produced contain natural rubber (primarily in the truck market)⁽⁹²⁾, and natural rubber is a constituent of both passenger and truck tire casings⁽³³⁾.

For natural rubber products, the induction of Type I (immediate hypersensitivity) allergic response is mediated by the human IgE antibody. The potential to induce allergy in sensitized and non-sensitized populations is dependent on the level of natural rubber protein antigens in a product, typically assessed using an *in vitro* IgE binding assay or *in vivo* skin prick test. With regard to the Type I allergic responses associated with the specific proteins in natural rubber⁽⁹⁴⁾, it is important to evaluate the two main types of natural rubber products used in the manufacture of products. Products based on solid natural rubber, such as tires or footwear are processed differently than more elastic products based on untreated natural rubber latex (NRL), such as surgical gloves or balloons. The production of treated solid, or bale natural rubber, requires intensive heating which decreases the levels of proteins by several orders of magnitude and the chemical additives used further decrease the bioavailability of the remaining protein. In contrast, dipped products are based on raw natural rubber latex with little pre-treatment, retaining many of the antigenic proteins from the raw material when sufficient washing or chlorination treatments are not applied. For example, dry rubber and dipped rubber extracts tested using the *in vitro* IgE binding assay demonstrate that the levels of allergen were up to 1000 times lower in dry versus dipped products (Figure 2). As expected, dry rubber products do not elicit skin reactions when tested and are generally considered free of the protein allergy problems reported for dipped products⁽⁹⁴⁾.

Figure 2: Relative Quantity of Extractable Allergens in Various Rubber Products⁽⁹⁵⁾



There is no evidence based on occupational exposures to dry rubber products to support the hypothesis that this form of rubber is a potent allergen. First, if the level of allergen in tire products were above a clinically relevant threshold, it would be reasonable to expect a high incidence of natural rubber allergy in the tire industry, especially in tire retreading or processing where buffing and grinding leads to airborne tire dust particulate. However, no such case reports or studies have been published in the literature⁽⁹²⁾. Additionally, a natural rubber IgE reaction has not been found in a survey of 208 workers from 9 different rubber manufacturing companies in the Netherlands⁽⁹²⁾.

Tire tread particle extracts have been assessed for binding with serum IgE from latex-sensitive patients. Miguel et al. measured the natural rubber latex (NLR) allergens in radial tires, truck tires, and recap waste treads from truck tires at levels of 3.48, 1.31 and 0.6 µg protein/g tire tread respectively⁽⁴⁸⁾. Based on this data and a No-Adverse Effect Level of 55 to 100 ng/m³ determined from a controlled study in latex-sensitive population⁽⁹⁶⁾, Finley et al. concluded that the weight of evidence indicates that natural rubber in tire particulates are not a significant contributor to asthma⁽⁹⁴⁾.

Of note is a recent EPA action under the Toxic Substances Control Act (TCSA), which denied a petition to prohibit use and distribution of natural rubber latex adhesives with a total protein content greater than 200 µg/g dry weight based on ASTM D-1076-06⁽⁹⁷⁾. In its denial of the petition, EPA stated that a regulation requiring reduced protein content would be unlikely to reduce natural rubber allergy in the general population. The EPA also cited the governmental evaluation, including a June 2004 Consumer Product Safety Commission assessment that found that while many consumer products contain natural rubber, there are few documented cases of reaction to these products. Of the case reports showing an association, most were associated with medical products. EPA concluded that the CPSC evaluation suggests that risks associated with natural rubber are “relatively insubstantial”.

Exposure to allergens from the use of ground rubber in CRM asphalt is also unlikely. As suggested by Liu et al, any allergens that may be present in ground rubber are likely to remain in the pavement matrix⁽⁹⁾. The conclusion that ground rubber and other recycled tire uses do not pose a threat to the development of natural rubber-associated allergies or respiratory disease, despite the presence of natural rubber in tire compositions, is further supported by an absence of occupational natural rubber allergies in the tire industry⁽⁹²⁾.

5.1.3 Dermal Exposure to Ground rubber

Exposure to ground rubber through dermal contact may occur through the use of ground rubber in playground applications and athletic fields. In addition to the concern of natural rubber allergy from the presence of natural rubber protein in some tire formulations, some of the chemicals used in tire manufacturing are thought to induce allergic contact dermatitis.⁽⁹⁸⁾ Furthermore, allergic contact dermatitis has been demonstrated in employees working in rubber manufacturing facilities⁽⁹⁹⁾. As such, some argue that there is a potential for allergenic response via dermal contact to ground rubber⁽¹⁰⁰⁾.

The California OEHHA conducted a skin sensitization test to evaluate the potential for allergic response due to dermal contact with rubberized playground surfaces⁽¹⁾. In this study elicitation of an allergenic response in the guinea pig, (a standard model for identifying human contact skin sensitizers) from exposure to materials (including ground rubber) used in playground surfaces was evaluated⁽¹⁰¹⁾. Test samples were applied to the animal's skin during three six-hour induction exposures each separated by one week. Following the induction exposure regimen, the animals were challenged with the test samples for six hours and evaluated 24 and 48 hours later for signs of erythema. A second challenge was initiated one week later. None of the rubber containing material, including the ground rubber, initiated an allergic response or indicated sensitization⁽¹⁾. While this study was intended to evaluate the potential for development of allergy in response to the use of recycled tires in playground surfaces, it too is applicable to dermal contact with ground rubber used in athletic fields as both are similar products in terms of chemical composition and contact surface area (not particle surface area) would determine toxicity.

In addition to contact allergy, Environment and Human Health, Inc. (EHHI) has raised concerns related to chemical leaching through skin from dermal contact with ground rubber as a potential mechanism of toxicity⁽⁵⁰⁾. However, because the ground rubber is unlikely to adhere, the prolonged contact required for uptake through the skin, which provides a reasonable barrier to many chemicals, is not likely. As such, uptake of chemicals is unlikely to result in systemic toxicity from dermal contact with ground rubber. This is supported by work performed by the Norwegian Institute of Public Health that evaluated the potential for dermal uptake of PCBs, PAHs, phthalates, and alkyl phenols from skin contact with rubber particles from artificial turf fields.⁽⁵⁾ In this analysis, 100g of rubber were leached in 1L of water over 48 hours in order to determine what is extractable from the rubber matrix. From this experiment, they determined that the leaching potential for PCBs, PAHs, phthalates, and alkyl phenols are 0.8×10^{-6} percent, 1×10^{-6} percent, 30×10^{-6} percent, and 5×10^{-6} percent, respectively. Using the leaching potential and assuming 100 percent uptake through skin absorption, the exposure estimates from dermal exposure to rubber granulate for adults, juniors, and children using athletic fields employing ground rubber is very low (for all chronic exposure scenarios, daily intake is less than 100 parts per trillion (ppt)). These concentrations, while exceedingly low, assume 100 percent of the chemical that can leach from the particle into aqueous solution is able to be absorbed by the skin. Because the skin provides a natural barrier to absorption of chemical toxicants, 100 percent bioavailability is unlikely in intact skin⁽¹⁰²⁾. In the case of phthalates it is known that only 5 percent of the compound will be absorbed into the systemic circulation⁽⁵⁾. In fact, based on a biomarker study in soccer players exposed to "intensive skin contact" with rubber infill, detection of biomarkers for PAHs was not increased over background, indicating that bioavailability of these compounds is low via the dermal pathway⁽²⁾.

A similar study evaluating the potential for chemicals (PAHs and phenylenediamines) in sand originating from tire barriers (as used in sandpits) to migrate through the skin was conducted by the Danish Ministry for the Environment⁽³⁵⁾. Four compounds (fluoranthene, pyrene, 6PPD, and IPPD) were able to migrate into artificial sweat from the sand. Based on these results, a risk assessment for exposure via this pathway was completed based on 200 cm² of exposed skin (child's thighs) and daily one hour exposure. Margins of safety for all

chemicals evaluated ranged from 10,000 to greater than 1,000,000, indicating negligible risk from this exposure scenario.

In summary, the results from these studies of dermal exposure indicate that the dermal pathway represents a low health risk from the use of recycled rubber products in playgrounds or artificial turf fields.

5.1.4 Other Toxicity Studies Regarding Ground rubber

In addition to the assessments described above, there are a few studies in the literature that investigate the impact of ground rubber on other endpoints of toxicity. Multiple researchers have investigated the potential for extracts from rubber materials to induce genetic changes in *in vitro* systems. Birkholz et al. performed an extraction of tire rubber with dichloromethane and evaluated mutagenicity of the pooled extract in *Salmonella typhimurium* with and without metabolic activation with human S9 (pooled liver enzyme fractions). This assay is regularly used as a screening level genotoxicity test, and has become a standard component of mutagenicity testing battery.⁽¹⁰³⁾ In none of the tests was the extract genotoxic to *Salmonella*⁽⁶⁾. Gualtieri et al. evaluated DNA damage to A549 cells, a human lung cell line, in response to tire debris organic extract (TDOE) using the Comet assay and detected a dose dependent increase in damage to the DNA.⁽¹⁰⁴⁾ However, the Comet assay as used is a non-specific DNA damage assay that is difficult to replicate, sensitive to physical changes in the environment, and does not provide specific information regarding the mutagenic potential of the extract itself.

In both of these studies, extracts were performed using dichloromethane, but the Gualtieri study utilized particulate ranging from 10 to 80 μm , as this study was intended to evaluate the potential for lung damage in response to inhalation of tire particulate. Therefore, while the composition of the rubber products in each study (tire debris in Gualtieri and ground rubber in Birkholz) may be similar, the total surface area for extraction is much higher in the Gualtieri study allowing for greater quantity of the chemicals to be extracted. As a note, organic extraction does not represent a reasonable extraction method for mimicking lung exposure to humans. Rather, organic extraction allows for a worst case scenario in terms of exposure to organic constituents. Although several of the rubber compounding materials may be extractable using harsh solvents such as dichloromethane, few organic compounds can be extracted using water. Thus, genotoxicity screens using organic extracts must be viewed with caution, as relevance to human exposure scenarios is unclear and overestimation of genotoxic potential from organic constituents is likely. While the results from these two studies appear to be contradictory, the dissimilarities in study approach and endpoints of interest make comparison between the studies difficult. Further research may be required to fully characterize the mutagenic potential associated with exposure to ground rubber.

In a study of occupational exposures in scrap-tire shredding facilities, airborne particulate collected in two scrap-tire shredding plants was subject to a mutagenicity screen in *Salmonella*⁽¹⁵⁾. The particulate was extracted using acetone, the extract analyzed for chemical composition and tested for mutagenicity with and without S9. The extracted chemicals did not exhibit mutagenic activity in any of the strains tested in the absence of S9. The addition of S9 increased frame-shift mutations, but not base-pair substitution mutations. Based on chemical structure and known mutagenic activity of compounds used in rubber manufacturing, vulcanization stabilizers (and degradation products such as N-nitrosamines) and PAHs may contribute to the mutagenic potential of the particulate matter generated during scrap-tire shredding. The authors caution that without understanding the quantities of particulate generated and the ability of the body to absorb chemicals through the particulate, conclusions regarding the mutagenicity of these particles *in vivo* are premature. As

suggested earlier, methodology utilizing organic extraction is not the best model for anticipating mutagenic effects in the human lung. In fact, organic extraction is likely to exaggerate the mutagenic potential of organic constituents, and therefore the findings from this study may not be relevant to human exposure scenarios.

In response to a concern that artificial turf fields may increase *Staphylococcus aureus* infections, a comparison study was initiated at Penn State University to evaluate microbial populations in rubber-infilled artificial turf fields versus natural grass fields. Total microbial numbers were *lower* in synthetic turf systems when compared to natural grass fields. *Staphylococcus aureus* was not found on any of the playing surfaces. One explanation offered is that the surface temperatures associated with rubber-infilled artificial turf fields, which are much higher than natural grass fields, are not conducive to the growth of many infectious microbes, including *S. aureus*⁽²³⁾. This finding is somewhat contradictory to studies suggesting that play on artificial turf surfaces may be a risk factor for the *S. aureus* infections⁽¹⁰⁵⁻¹⁰⁷⁾. However, artificial turf fields are more abrasive than natural grass fields, and as a result, athletes are more prone to epidermal injuries such as cuts or abrasions⁽²⁵⁾. Therefore, transmission of microbes through locker room activities (towel or equipment sharing, for example) could result in a higher likelihood of skin penetration and subsequent infection.

5.2 IMPACTS ON THE ENVIRONMENT

In considering the hazards associated with the use of ground rubber in commercial applications, such as playgrounds or athletic fields, ecological endpoints are a necessary consideration. A standard aquatic toxicity battery as recommended by the EPA includes evaluating lethality or growth inhibition in algae, invertebrates (often *Daphnia magna* or *Ceriodaphnia dubia*), and fish, although the approach for estimating aquatic toxicity of solids is not straightforward⁽¹⁰⁸⁾. Other international regulatory bodies (OECD, Health Canada) employ similar recommendations, but sometimes use different test species. The method used in much of the existing literature addressing the toxicity of tire shreds, ground rubber, or other tire-related material (wear particles, etc.) includes using a leachate of the rubber product and treating the test species.

Birkholz et al. leached 250g of both fresh and aged ground rubber from tires in one liter of water and treated bacteria (*Vibrio fischeri*), algae, microcrustaceans (*Daphnia magna*), and fish (*Pimphales promelas*) with the resulting leachate⁽⁶⁾. While leachates from the fresh ground rubber were toxic to all species investigated, aging of the ground rubber resulted in a nearly 60 percent reduction in toxicity. Further reduction in toxicity occurred with the addition of nutrients, sewage seed, and five days of aeration. They conclude that while undiluted leachate from fresh tire rubber may pose a moderate threat to aquatic toxicity, environmental aging will attenuate this toxicity such that the risk is not appreciable. Further, they state that surface runoff from playgrounds or athletic fields containing ground rubber is likely to be diluted by larger bodies of water (in which the aquatic species dwell), which should eliminate the possibility that even fresh ground rubber is an ecologic hazard⁽⁶⁾.

Sheehan et al. evaluated the toxicity of samples (and serial dilutions thereof) collected from the aforementioned field study in Maine⁽⁶³⁾ to *Ceriodaphnia dubia* and *Pimphales promelas*⁽¹⁰⁹⁾. It was noted, however, that the shreds used in the field study contained exposed steel belts at the cut edges of the shreds. Metallic material from steel belts is removed from ground rubber during production. Survival and reproductive capacity of *C. dubia* was negatively impacted by tire shreds placed below the water table over control, but not from that placed above the water table. Furthermore, it was expected that *C. dubia* toxicity would be reduced to that equivalent to background upon a 2- to 4-fold dilution of leachate. It was suggested that the demonstrated toxicity was related to the concentration of iron (and possibly other metals), which are likely attributable to the presence of steel belting in the shreds.

ADEME, in coordination with ALIAPUR and Fieldturf Tarkett (a manufacturer of artificial turf field surfaces), assessed the environmental impact of the use of ground rubber in outdoor artificial turf fields⁽⁴⁾. In this study, ground rubber infilled artificial turf fields were built atop a lysimeter and water collection system and treated with simulated rain (one year of rainfall). Percolates were collected weekly, combined, and analyzed at 1, 2, 3, 6, 9, and 11 months. The percolates were then used to treat *Daphnia magna* and *Pseudokirchneriella subcapitata* (soft water algae). Results from this study indicate that these species were not affected by the percolates from the rubber-infilled artificial turf fields.

The Laboratory of Ecological Risk Assessment in the Netherlands (RIVM) assessed leaching of zinc from rubber infilled artificial turf fields⁽⁴⁵⁾. They estimated zinc loads in soil, groundwater and surface water based on leaching results from both laboratory and field experiments utilizing both fresh and aged ground rubber. Based on these studies, they conclude that zinc leaching (and thus load) increases with aging. The predicted zinc loads to each compartment were compared to environmental risk criteria for soil, groundwater and surface water and found to exceed these criteria in all three environmental compartments, indicating that, based on this study, the use of rubber-infilled artificial turf fields presents an ecological risk. To address the uncertainties in this analysis, RIVM recommends a series of studies to: investigate the impact of aging of rubber in constituent releases to the environment; monitor drainage water from artificial turf fields utilizing rubber as an infill component; perform bioassays with drainage water; and to construct a miniaturized artificial turf field with a lysimeter to provide insight on emission and mobility of zinc under actual field conditions. The results of the above studies can provide useful information to improve the modeling and more accurately estimate risk to the environment.

The Norwegian Institute for Water Research, based on a leaching study conducted previously that collected run-off from artificial turf fields, modeled local concentrations of metals, PAHs, phthalates and other rubber-affiliated chemicals in surface water and sediment to estimate PEC/PNEC ratios, a measure of ecologic risk⁽¹¹⁰⁾. The risk assessment performed in this study was specific for local environments (i.e. surface runoff from artificial turfs in nearby streams). The PEC/PNEC ratio exceeded 1.0 (indicating a potential for ecologic risk in local environments) for octylphenol (2.9), total PAHs (1.13), and zinc (40) in surface water. In sediment, only octylphenol and zinc result in PEC/PNEC ratios greater than one. However, the leachate studies that provide the environmental concentrations for this study

were determined based on a laboratory leaching study (recycled ground rubber placed in water), and were not collected based on a field study (or under simulated field conditions). The authors suggest that, while the results indicate an ecological risk, further work is required in order to more definitively characterize risk in a more realistic setting. In addition, they state that the ecological effects are likely to elicit an impact locally only, and that over the course of the year, the limited runoff is not expected to be an important source of pollution when compared to other potential sources.

In addition to studies evaluating the potential for ground rubber products to induce toxicity in aquatic species, multiple research groups have investigated the possibility that particulate generated from the use of tires may impact aquatic organisms⁽¹¹¹⁻¹¹⁴⁾. While these particles are demonstrably different from the ground rubber used in playgrounds and artificial turf fields, especially in relation to particle size and total surface area, the results from these studies are relevant to the question of ecological risk related to the use of tires in such commercial applications. Gualtieri et al. observed lethality in *Daphnia magna* following treatment with leachates from tire tread debris⁽¹¹¹⁾. Wik and Dave, in a series of experiments, demonstrated similar findings, with toxicity to *Daphnia magna* from a range of high concentrations of leachate from tire debris^(112, 113). Care must be taken in extrapolating the results of these studies to ground rubber used in playgrounds or artificial turf fields. The average size of the particulate used in these studies (to generate the leachate) is several orders of magnitude smaller than ground rubber used in these applications. As a result, the total surface area and relative leaching potential is much larger for these studies. Furthermore, in the Gualtieri study, a harsh organic solvent was used to extract the tire debris, which results in a different profile of chemicals in the leachate than would result from using water, an environmentally relevant leaching medium⁽¹¹¹⁾.

Additional studies evaluating the impact of other tire-related material (whole tires, scrap tire fill, etc.) were performed by several researchers. The Minnesota Pollution Control Agency conducted a general vegetation survey on roads containing or lacking scrap tire fill that indicated no difference between the two road types⁽⁶¹⁾. In Connecticut, a group of researchers collected anecdotal evidence regarding the impact of tires used as energy absorbing bumpers on fresh water lake docks which indicated that the tires have little effect on the water, fish, or plant life in the lake⁽⁹⁾. The Canadian Water Research Institute prepared contaminated water by submerging whole passenger tires in natural ground water with continuous aeration. The contaminated water was used to treat fish (rainbow trout and fathead minnows). The leachate was 100 percent lethal within 48 hours to the rainbow trout, but no toxicity was demonstrated in fathead minnows. Zinc was identified as the toxic constituent⁽¹¹⁵⁾.

The ecological risk associated with the use of ground rubber in playgrounds and athletic fields has been investigated in several studies by evaluating the impact of leachates from ground rubber or other tire-related material (i.e. tire shreds) on aquatic life. In the majority of these studies, zinc was identified as the most likely toxic constituent. While lethality was observed in several species in many of the studies, aging of the rubber material and dilution from natural systems in which the species live is likely to prevent toxic effects demonstrated

from ground rubber leachates. Therefore, the use of ground rubber in athletic fields and playgrounds is unlikely to represent an ecological risk.

5.3 CONCLUSIONS FROM LITERATURE REVIEW

The literature surrounding the safety of ground rubber in uses such as playgrounds or artificial turf fields is, collectively, quite thorough in addressing potential concerns from the consumer standpoint. Each likely exposure pathway has been investigated, and in many cases deemed to be an unlikely risk to either human or ecological receptors. In many cases, authors focused on children as a susceptible subpopulation, and yet risks remained low. The current literature does not provide a compelling argument for discontinuation of the use of ground rubber products in playgrounds or athletic fields from the standpoint of either human or ecological risk. Furthermore, there are significant benefits associated with the use of ground rubber in these applications. Many of the criticisms that remain focus on the absence of toxicity information relating to some of the chemicals associated with ground rubber. However, due to the sheer volume of chemicals (both natural and synthetic) that are found in consumer products, a complete toxicity profile for all chemicals for which humans are exposed is a goal requiring many decades of future study. The vast number of synthetic and natural chemicals has motivated health scientists to develop tiered and hierarchical approaches to safety assessment. The following section details the approaches used for chemical safety assessments of whole products, citing examples from both natural and synthetic products.

6.0 SCIENTIFIC APPROACH TO CHEMICAL, SITE AND PRODUCT SAFETY ASSESSMENT

The U.S. Environmental Protection Agency has established an overall framework for assessing the nature and extent of site-specific health risks as part of the National Oil and Hazardous Substances Pollution Contingency Plan (NCP)⁽¹¹⁶⁾. A comprehensive evaluation of human health risk involves several key components. The first step is the collection and evaluation of data relative to human health and the identification of substances for risk characterization. An exposure assessment is then performed to assess the magnitude, frequency and duration of exposure, typically with a characterization of typical and reasonable maximum exposure. As part of an exposure assessment, the pathways of exposure (e.g. oral, dermal, inhalation), exposure concentrations and characteristics of the exposed population are used to calculate intake.

In parallel to the exposure assessment, a toxicity assessment is completed to describe the types of adverse health effects and dose-response relationship which describes the relationship between magnitude of exposure and adverse effects. The process of characterizing the nature and extent of strength of evidence of causation, as well as determining whether the agent can cause a specific adverse health effect is termed hazard identification. The quantitative use of toxicity information to relate the administered dose to incidence of adverse outcomes in humans at different exposure levels is termed dose-response evaluation. Although most natural and anthropogenic settings are characterized by complex mixtures of inorganic and organic chemicals, many of which are not fully studied, site risk assessments are primarily based on currently existing toxicity information developed for specific chemicals.

The outcome of the exposure and toxicity assessment is summarized in the risk characterization. One of the important features of a risk characterization is that both qualitative and quantitative statements regarding potential for noncancer or cancer risks are developed. Another purpose of the risk characterization is to evaluate uncertainties and to address the need for further characterization.

High quality general human health and ecological evaluations of recycled tire rubber products which conform to the U.S. EPA risk assessment framework have been completed, and these assessments have concluded that these products present a low likelihood of adverse health effects^(1, 4, 5). Recently, the Bainbridge Island School District located in Washington State requested an initial site-specific assessment of potential human health risks associated with the installation of a synthetic turf field based on recycled tire rubber⁽¹¹⁷⁾. This assessment was consistent with the U.S. EPA risk assessment framework and considered exposure concentration, route of chemical exposure, duration of exposure and chemical potency. The assessment identifies the important distinction between the composition of a product and the potential environmental exposure. For many consumer products, the component chemicals are not accessible to humans (e.g. the lead used inside cathode ray tube computer monitor) while in other instances the chemicals are accessible but absorbed by the body at different rates (e.g. the age dependent internal uptake of lead in paint chips). Analytical methods which monitor unventilated headspace or total chemical composition

dissolved in strong acid are useful for hazard identification but unusable for assessment of exposure, which is a critical step in the risk assessment process. Exposure scenarios representative of upper bound Pacific Northwest exposures were assessed in a child sport play scenario and teenager sport play scenario. The risk and exposure assessments were based on key chemical compounds determined based on a review of the literature and paired with conservative (i.e. likely to overestimate risk) assumptions of 261 days/year exposure frequency, high exertion breathing rates for 3 hours per day and use of indoor concentrations as a surrogate for outdoor concentrations. The assessment was consistent with other generic evaluations of recycled tire rubber and concluded some chemicals leach or volatilize from the recycled product in small amounts, but the weight of evidence indicated that the carcinogenic and noncarcinogenic risk for inhalation, dermal adherence and incidental ingestion pathways were minimal.

Although comprehensive health assessments of ground rubber based fields have been completed which are consistent with the EPA risk assessment framework, there are additional considerations when evaluating the chemical composition of a discrete consumer product. Chemicals in the environment are derived from natural sources such as plant or animal metabolism, forest fires or weather or from synthetic sources during chemical manufacture. There have been over 39 million organic and inorganic compounds identified from synthetic or natural sources in the scientific literature since 1957 with each of these compounds assigned a unique identifier by the American Chemical Society termed a CAS number.⁽¹¹⁸⁾ In the quantitative assessment of potential human health risk, the current state of knowledge precludes individual assessment of each of these compounds. For example, as of April 25, 2008, there were only 544 substances with peer reviewed quantitative toxicity factors listed in the U.S. EPA Integrated Risk Information System (IRIS) database, or 0.001 percent of the total substances identified since 1957.⁽¹¹⁹⁾ Similarly, U.S. EPA Region III maintains a database of toxicity values which includes additional provisional data but includes only 377 compounds⁽¹²⁰⁾.

The absence of toxicity factors for each possible compound does not imply that a framework to rigorously assess human safety of complex products does not exist. In contrast, there are a variety of tools health scientists use to assess product safety, many of which rely on hierarchical approaches, human epidemiology and evaluation of indicator compounds for which toxicity is well characterized. One example of such an approach is the U.S. EPA's Voluntary Children's Chemical Evaluation program, which identified 23 compounds for detailed assessment based on data which showed exposure had occurred based on human blood, breast milk or exhaled breath⁽¹²¹⁾. For these compounds, sponsoring companies were asked to identify all of the sources of exposure that contributed to the observed body burdens. In this program, a tiered approach was used to assess data needs for both potential hazard and exposure. Another example is in the assessment of disinfection byproducts created during drinking water treatment, where U.S. EPA has identified and cataloged more than 600 halogenated and other byproduct chemicals⁽¹²²⁾. Based on a peer review, 252 of these compounds were detected in drinking water in various studies. Of the chemicals detected in drinking water, only 30 were considered to have sufficient toxicity data and 209 were evaluated for cancer potential using theoretical structural activity relationships. Compounds that show high potential for toxicity were considered for further animal or other testing.

Many of the compounds were considered to be of low priority for further study due to the low likelihood of adverse health effects. The qualitative hierarchical treatment of potential chemical risk is an essential and key step in the assessment of real world consumer products, including food, many of which are comprised of complex mixtures.

Although exposures to complex mixtures are frequently associated with synthetic, or human made chemicals, there are many examples of natural products for which individual chemical assessment is not plausible. For example, a detailed chemical analysis of natural products such as roasted coffee reveals an extensive list of over 1,000 compounds, the majority of which traditional quantitative risk assessment is not possible⁽¹²³⁾. Of the 30 compounds tested for rodent carcinogenicity, 21 were positive, resulting in approximately 10 mg of rodent carcinogen per cup of coffee. One of the compounds detected, the carcinogenic PAH benzo(a)pyrene, is a common byproduct of cooking. However, most people generally consider coffee to be an extremely safe product when consumed in moderation based on the characteristics of the product. Coffee is not necessarily a risk factor in human cancer. Rather, this example shows that natural compounds that are carcinogens in high dose rodent tests are ubiquitous in the human diet, at levels often far exceeding synthetic chemical exposure.

The most abundant semi-volatile organic compound identified in ground rubber head space analyzed by the Connecticut Agricultural Experiment Station was benzothiazole. EHFI specifically noted the lack of information regarding benzothiazole was a severe limitation of the existing research on recycled tire rubber exposure. However, EHFI failed to recognize that the Joint FAO/WHO Expert Committee on Food Additives (JECFA), which is an international scientific expert committee administered by the Food and Agriculture Organization of the United Nations (FAO) and the World Health Organization (WHO), considers ingestion of benzothiazole to be a safe food additive when used as a flavoring agent and is considered to be generally recognized as safe (GRAS) by the Flavor and Extract Manufacturers Association (FEMA)⁽¹²⁴⁾. Further, there are many natural dietary sources of this compound such as fresh apple, sour cherry, butter, wine, tea and cooked scented rice⁽¹²⁵⁾. This example illustrates the ubiquity of chemicals in our diet as well as the importance of comprehensive evaluation of health hazards.

The University of California – Berkeley maintains a Carcinogenic Potency Database (CPDB), which catalogs 6,500 chronic, long-term animal studies on approximately 1,500 chemicals. This research group, including the creator of the Ames test, a mutagenic biological assay screening method, has identified several key points regarding synthetic versus natural chemical exposure which are essential for reliable assessment of product health effects, costs and benefits. Natural chemical exposure is far broader and much greater in magnitude than synthetic chemical exposure, yet exposure to natural chemicals has not been systematically evaluated. For example, 99.99 percent of dietary pesticides (totaling 5,000 to 10,000 compounds) are estimated to be naturally produced by plants for protection against fungi, insects or animal predators⁽¹²³⁾. Accordingly, public and regulatory perception of carcinogenic hazards, which emphasize synthetic chemicals, is not properly aligned with true human exposure. Given the level of these natural pesticides, dietary human exposure to known rodent carcinogens is frequent and high in magnitude. Assessment of the potential

health risk of exposure to natural compounds should not reduce the level of study of synthetic chemicals. However, knowledge of the ubiquitous presence of natural and synthetic compounds (many of which are carcinogenic in rodent studies at high doses) is useful in understanding the tiered and hierarchical scientific process which must be used by health scientists to assess food and consumer product safety.

In food or consumer products, many inorganic, volatile and semi-volatile compounds will be detected for which detailed toxicity assessments have not been completed. In these instances, three types of assessments are performed. First, whole product safety is assessed using animal data. For example, in the assessment of white spirits solvent (mineral spirits), guinea pigs were the most sensitive of five species based on continuous inhalation exposure for 90 days⁽¹²⁶⁾. Mineral spirits are complex products derived from crude oil of variable raw composition and whole product testing is essential in understanding human health risk. Next, key individual chemicals of known toxicity are evaluated. In the case of mineral spirits, scientific consensus dictates that an individual exposure or risk assessment be performed for trace aromatic compounds such as 1,2,4-trimethylbenzene and benzene. Finally, human epidemiological information is considered, typically from controlled studies or occupational exposure assessments which report short and long-term neurological, target organ specific, irritation and other effects. In some instances, reliable human epidemiological data may not be available due to the difficulty in controlling for confounding exposures or lack of knowledge regarding historical dose or non-occupational dose. However, even in these instances, qualitative case reports regarding respiratory irritation or dermal sensitivity may be available.

EHHI recently issued a report recommending a moratorium on the installation of fields or playgrounds that use ground-up rubber tires based on limited testing which showed that low levels of metals or organic compounds are leachable from tire rubber, extrapolation from occupational studies, and critique of relevant quantitative studies. While the creation of a long term research program for recycled tire rubber products may be appropriate, the weight of evidence and range of studies that have been performed to date does not support EHHI's conclusion that use of existing fields should be limited or that planned fields should not be installed. EHHI's criticisms of existing studies fail to acknowledge the spectrum of valid qualitative and quantitative methodologies which have been traditionally employed to evaluate many of the useful, but chemically complex, consumer products where human contact occurs on a daily basis. Specific examples of EHHI criticisms that could generically be applied to other common products include surface temperature (comparable to upper bound outdoor asphalt basketball court temperatures of 160 °F), leachable organic chemicals lacking toxicity factors (comparable to several hundred semivolatile and volatile compounds found in roasted coffee) or the potential for unacceptable levels of zinc in the rubber tire mulch leachate (comparable to zinc leached from galvanized residential cistern rainwater collection systems)⁽⁵⁰⁾. The concerns publicized by EHHI represent a viewpoint that is unsupported by the current scientific consensus, or weight of evidence, as well as the views of the majority of governmental agencies.

As can be seen from these examples, criticisms of ground rubber which question the safety of the product based solely on the absence of comprehensive peer reviewed toxicity database

for every possible detected organic compound are quite misleading. Scientific health and safety assessment of natural and processed food and food additives, as well as consumer products is necessarily based on a holistic and hierarchical approach which synthesizes a number of different types of information to inform an assessment of product safety. Such assessments ensure that beneficial products are available to the public, and that use of these products will not result in unacceptable adverse human or ecological effects.

7.0 CONCLUSIONS

Based on a review of the available studies, there is a low likelihood of adverse health effects for children or athletes exposed to recycled tires found at playgrounds or athletic fields (Table 1). There were no short-term or urgent research needs identified upon consideration of the weight of evidence presented in the current literature. However, additional research could be useful in better defining and communicating potential risks. One such area is assessment of fine particulate exposure at ground rubber installations and assessment of outdoor airborne concentrations of volatile organic compounds as a function of temperature. The California Integrated Waste Management Board is currently considering completing research in each of these areas. Based on the range of questions and concerns among various stakeholders, another area of potential inquiry could be site-specific assessment of zinc concentrations in local ecosystems. Although some studies have suggested that more information is needed regarding the potential for natural rubber allergy after contact with recycled tire products, no evidence was found to support the hypothesis that tires, which are made from natural rubber in bale form, are likely to cause adverse allergic reactions.

REFERENCES

1. California Integrated Waste Management Board, *Evaluation of Health Effects of Recycled Waste Tires in Playground and Track Products*. 2007, Integrated Waste Management Board: Sacramento, CA.
2. Hofstra, U., *Environmental and Health Risks of Rubber Infill: Rubber crumb from car tyres as infill on artificial turf*. 2007, INTRON.
3. van Bruggen, M., E.M. van Putten, and P.C.J.M. Janssen, *Nitrosamines released from rubber crumb* 2007, RIVM: Bilthoven, the Netherlands.
4. Moretto, R., *Environmental and health assessment of the use of elastomer granules (virgin and from used tyres) as filling in third-generation artificial turf*. 2007, ADEME/ALIAPUR/FIELDTURF TARKETT. p. 1-27.
5. Norwegian Institute of Public Health and the Radium Hospital, *Artificial turf pitches – an assessment of the health risks for football players*. 2006, Norwegian Institute of Public Health and the Radium Hospital: Oslo. p. 1-34.
6. Birkholz, D.A., K.L. Belton, and T.L. Guidotti, *Toxicological evaluation for the hazard assessment of tire crumb for use in public playgrounds*. J Air Waste Manag Assoc, 2003. **53**(7): p. 903-7.
7. Reschner, K. *Scrap Tire Recycling- A Summary of Prevalent Scrap Tire Recycling Methods*. 2006 [cited; Available from: <http://www.energymanagertraining.com/tyre/pdf/ScrapTireRecycling.pdf>].
8. Stutz, J., et al., *Recycled rubber products in landscaping applications*. 2003, Tellus Institute Resource and Environmental Strategies: Boston, MA. p. 1-19.
9. Liu, H.S., J.L. Mead, and R.G. Stacer, *Environmental effects of recycled rubber in light-fill applications*. Rubber Chem & Tech, 2000. **73**(3): p. 551.
10. U.S. EPA. *Management of Scrap Tires*. 2008 [cited April 28, 2008]; Available from: <http://www.epa.gov/garbage/tires/basic.htm>.
11. Rubber Manufacturers Association. *RMA Briefing Sheet: The use of scrap tire as playground material*. 2008 [cited 2008 4/21]; Available from: http://www.rma.org/scrap_tires/scrap_tire_markets/.
12. U.S. EPA. *Tire Fires*. 2007 [cited April 28, 2008]; Available from: <http://www.epa.gov/garbage/tires/fires.htm>.
13. Anderson, M.E., et al., *A case study of tire crumb use on playgrounds: risk analysis and communication when major clinical knowledge gaps exist*. Environ Health Perspect, 2006. **114**(1): p. 1-3.
14. Arizona Department of Environmental Quality, *Waste Tire Report*. 2002, Arizona Department of Environmental Quality: Phoenix, AZ. p. 1-9.
15. Chien, Y.C., et al., *Assessment of occupational health hazards in scrap-tire shredding facilities*. Sci Total Environ, 2003. **309**(1-3): p. 35-46.
16. Texas Natural Resources Conservation Commission (TNRCC). *TNRCC Information: The Many Uses of Rubber*. 2008 [cited 2008 4/28]; Available from: <http://www.tceq.state.tx.us/assets/public/permitting/registration/tires/crumb.pdf>
17. Recipneu. *Memo: Cryogenic rubber infill for artificial grass sports fields: Comprehensive description of its superior technical features*. 2006 [cited 2008 4/28]; Available from: <http://www.tufsd.org/documents/budget/turf/fields/MEMOpercent20->

- percent20Criogenicpercent20rubberpercent20infillpercent20forpercent20Artificialper
cent20Grasspercent20percent204percent20Sept.pdf.
18. Sportex. *Ambient vs. cryogenic: Do rubber processing methods affect synthetic turf system quality?* 2008 [cited 2008 4/28]; Available from: http://www.sportexe.com/PDF/101/TT_Rubber.pdf.
 19. Clean Washington Center (CWC). *Best Practices in Scrap Tires & Rubber Recycling: Ambient versus Cryogenic Grinding.* 1998 [cited 2008 4/28]; Available from: www.cwc.org/tire_bp/t_bp_pdf/2-03-04.pdf
 20. Weibold, R. *Rubber granulates for the purpose of infilling artificial turf.* eximlink Ltd. 2005 [cited 2008 4/28]; Available from: http://eximlink.com/wp-content/uploads/2007/10/eximreport_on_artificial_turf_infill.pdf.
 21. Amme, R.C., et al., *Scrap tires in full swing: a recent study demonstrates the benefits of playground surfaces covered with recycled crumb rubber.* Resource Recycling, 2003: p. 1-2.
 22. Spencer, A.-M., *What lies beneath: learn the pros and cons of a variety of surfacing solutions.* , in *Parks & Recreation.* 2005.
 23. McNitt, A.S. *Evaluation of Playing Surface Characteristics of Various In-Filled Systems.* 2008 April 9, 2008 [cited 2008 4/20]; Available from: <http://cropsoil.psu.edu/mcnitt/infill.cfm>.
 24. Claudio, L., *Synthetic turf: health debate takes root.* Environ Health Perspect, 2008. **116**(3): p. A116-22.
 25. Meyers, M.C. and B.S. Barnhill, *Incidence, causes and severity of high school football injuries on FieldTurf versus natural grass.* Am J Sports Medicine, 2004. **32**(7): p. 1626-1638.
 26. New York City Department of Health and Mental Hygiene. *Artificial Turf Fact Sheet.* 2008 [cited April 15, 2008]; Available from: <http://www.nyc.gov/html/doh/html/code/code-turf.shtml>.
 27. Ledoux, T., *Preliminary Assessment of the Toxicity from Exposure to Crumb Rubber: its use in Playgrounds and Artificial Turf Playing Fields.* 2007, New Jersey Department of Environmental Protection, Division of Science, Research and Technology: Trenton, NJ. p. 1-2.
 28. Connecticut Department of Public Health, *Fact Sheet Artificial Turf Fields: Health Questions*, Environmental & Occupational Health Assessment Program, Editor. 2007: Hartford, CT. p. 1.
 29. Lamie, P. *Memorandum to: Richard Reine, Director Concord Public Works. Rubber Crumb Health Risk Evaluation.* April 24. 2007 [cited 2008 4/28]; Available from: http://www.concordma.gov/pages/ConcordMA_publicworks/rubber.pdf.
 30. Executive Order #13045, *Protection of Children From Environmental Health Risks and Safety Risks.* 1997.
 31. U.S. EPA Region 8, *Region 8 Crumb Rubber Research and Recommendation.* 2008.
 32. Lawrence Livermore National Laboratory, *Effect of waste tires, waste tire facilities and waste tire projects on the environment*, Department of Energy, Editor. 1996.
 33. U.S. EPA, *AP-42 Section 4.12, Manufacture of Rubber Products*, United States Environmental Protection Agency (U.S. EPA), Editor. 1997.
 34. Japanese Automobile Tyre Manufacturers Association, *JATMA Report No. 15-8A-0098.* 1998.

35. Nilsson, N.H., A. Feilberg, and K. Pommer, *Emissions and evaluation of health effects of PAH's and aromatic mines from tyres*. 2005, Danish Ministry of the Environment.
36. European Commission, *Directive 2005/69/EEC (OJ L323, 9/12/2005, p. 51)*. 2005.
37. International Agency for Research on Cancer, *Some non-heterocyclic polycyclic aromatic hydrocarbons and some related exposures*. 2005, IARC Monographs on the Evaluation of Carcinogenic Risk to Humans.
38. National Toxicology Program, *Report on Carcinogens, Eleventh Edition*, Department of Health and Human Services, Editor. 2005.
39. Agency for Toxic Substances and Disease Registry, *Public Health Statement for Polycyclic Aromatic Hydrocarbons (PAHs)*. 1995.
40. Takada, H., T. Onda, and N. Ogura, *Determination of polycyclic aromatic hydrocarbons in urban street dusts and their source materials by capillary gas chromatography*. Environ Sci Technol, 1990. **24**: p. 1179.
41. Agency for Toxic Substances and Human Services, *Toxicological profile for di(2-ethylhexyl)phthalate*, Department of Health and Human Services, Editor. 2002.
42. The National Academies, *Dietary Reference Intakes for Vitamin A, Vitamin K, Arsenic, Boron, Chromium, Copper, Iodine, Iron, Manganese, Molybdenum, Nickel, Vanadium, and Zinc*. 2001, Washington, D.C.: National Academies Press.
43. U.S. EPA, *Zinc and compounds*, Integrated Risk Information System, Editor. 1992, National Center for Environmental Assessment.
44. Smolders, E. and F. Degryse, *Fate and effect of zinc from tire debris in soil*. Environ Sci Technol, 2002. **36**(17): p. 3706-10.
45. Verschoor, A.J., *Leaching of zinc from rubber infill on artificial turf (football pitches)*. 2007, RIVM: Bilthoven, the Netherlands.
46. Mattina, M.I., et al., *Examination of crumb rubber produced from recycled tires*. 2007, The Connecticut Agricultural Experiment Station Department of Analytical Chemistry: New Haven, Connecticut.
47. The Swedish Chemicals Inspectorate (Kemi), *Synthetic turf from a chemical perspective - a status report*. 2006, KEMIKALIENSPEKTIONEN Sundbyberg. p. 1-31.
48. Miguel, A.G., et al., *Latex allergens in tire dust and airborne particles*. Environ Health Perspect, 1996. **104**(11): p. 1180-6.
49. U.S. EPA, *Supplemental guidance for assessing cancer susceptibility from early-life exposures to carcinogens, External Review Draft*. 2003.
50. Brown, D.R., *Artificial Turf - Exposures To Ground-Up Rubber Tires - Athletic Fields - Playgrounds - Gardening Mulch*, N. Alderman and S. Addiss, Editors. 2007, Environment & Human Health, Inc.: North Haven, CT. p. 1-40.
51. U.S. EPA, *Child-Specific Exposure Factors Handbook, Interim Report*. 2002.
52. U.S. EPA, *Memorandum: Acephate-sensitivity analysis for turf risk assessment [Case #819371, PC Code 103301, DP Barcode D276433]*, Office of Prevention Pesticides and Toxic Substances, Editor. 2001.
53. U.S. EPA, *A probabilistic exposure assessment for children who contact CCA-treated playsets and decks*. 2005.

54. Freeman, N., et al., *Characterizing indoor-outdoor activity patterns of young children*, in *Annual Meeting of the International Society of Exposure Analysis*. 2005: Tucson, AZ.
55. Freeman, N.C.G., et al., *Quantitative analysis of children's microactivity patterns: The Minnesota Children's Pesticide Exposure Study*. *J Exposure Anal Environ Epidemiol*, 2001. **11**: p. 501-509.
56. Beamer, P., et al., *Analysis of a child's mobility on their micro-level activity pattern*, in *Annual Meeting of the International Society of Exposure Analysis*. 2004: Philadelphia, PA.
57. Black, K., et al., *Children's mouthing and food-handling behavior in an agricultural community on the US/Mexico border*. *J Expo Anal Environ Epidemiol*, 2004. **15**: p. 244-251.
58. AuYeung, W., et al., *Young children's hand contact activities*, in *Annual meeting of the International Society for Environmental Epidemiology*. 2004: Johannesburg, South Africa.
59. AuYeung, W., et al., *Young children's mouthing behavior: an observational study via videotaping in a primarily outdoor residential setting*. *J Children's Health*, 2004. **2**: p. 1-25.
60. Edil, T.B., J.K. Park, and J.Y. Kim, *Effectiveness of Scrap Tire Chips as Sorptive Drainage Material*. *J. Envir. Engrg*, 2004. **130**(7): p. 824-831.
61. Twin City Testing Corporation and A. Ronchak, *Waste tires in sub-grade road beds*, in *Environmental Study of the Use of Shredded Waste Tires For Roadway Sub-grade Support*. 1990, Minnesota Pollution Control Agency: St Paul, MN. p. 1-44.
62. Sengupta, S. and H. Miller, *An Evaluation of Recycled Tire Shreds as a Substitute for Gravel in Residential Soil Absorption Systems*, in *Civil Engineering Practice*. 2004. p. 33-52.
63. Downs, L., et al., *Water quality effects of using tire chips below the groundwater table*, in *Department of Civil and Environmental Engineering*. 1996, University of Maine: Orono, ME.
64. Humphrey, D.N. and E.K. Lynn, *Five Year Study of the Water Quality Effects of Tire Shreds Placed Above the Water Table*. 2000, Transportation Research Board: Washington DC.
65. International Agency for Research Cancer, *Butylated hydroxyanisole*. 1986, IARC Monographs on the Evaluation of Carcinogenic Risks to Humans.
66. Diefenderfer, B., I. Al-Qadi, and S. Diefenderfer, *Model to predict pavement temperature profile: development and validation*. *Journal of Transportation Engineering*, 2006. **132**(2): p. 162-167.
67. Steele, W., et al., *The thermodynamic properties of benzothiazole and benzoxazole*. 1991, Bartlesville, Oklahoma: National Institute for Petroleum and Energy Research.
68. Chang, F.H., et al., *Emission characteristics of VOCs from athletic tracks*. *J Hazard Mater*, 1999. **70**(1-2): p. 1-20.
69. California Integrated Waste Management Board. *Draft Scope of Work: Evaluation of the safety of artificial turf fields containing crumb rubber from recycled tires*. 2008 [cited 2008 4/24]; Available from: www.ciwmb.ca.gov/agendas/mtgdocs/2008/04/00023252.doc.

70. Wallace, L., E. Pellizarri, and C. Wendel, *Total volatile organic concentrations in 2700 personal, indoor, and outdoor air samples collected in the US EPA Team studies*. Indoor Air, 1991. **1**: p. 465-477.
71. Hodgson, A.T., et al., *Volatile organic compound concentrations and emission rates in new manufactured and site-built houses*. Indoor Air, 2000. **10**: p. 178-192.
72. Hoffman, K., et al., *The German environmental survey 1990/92 (GerES II): Sources of personal exposure to volatile organic compounds*. Journal of Exposure Analysis and Environmental Epidemiology, 2000. **10**: p. 115-125.
73. U.S. EPA, *Review of the National Ambient Air Quality Standards for Particulate Matter: Policy Assessment of Scientific and Technical Information*. EPA-452/R-05-005a. 2005, Washington D.C.: Office of Air Quality Planning and Standards (OAQPS).
74. Dockery, D., et al., *An association between air pollution and mortality in six US cities*. New England Journal of Medicine, 1993. **329**(24): p. 1753-1759.
75. Pope III, C., et al., *Particulate air pollution as a predictor of mortality in a prospective study of U.S. adults*. American Journal of Respiratory Critical Care and Medicine, 1995. **151**(3): p. 669-674.
76. Harrison, R. and J. Yin, *Particulate matter in the atmosphere which particle properties are important for its effects on health?* Sci Total Environ, 2000. **249**(1): p. 85-101.
77. Mauderly, J., L. Neas, and R. Schlesinger. *PM monitoring needs related to health effects*. in *Atmospheric observations: helping build the scientific basis for decisions related to airborne particulate matter. Report of the PM Measurements Research Workshop Chapel Hill, NC, July 22-23*. 1998. Cambridge, MA: Health Effects Institute.
78. Oberdorster, G., et al., *Association of particulate air pollution and acute mortality: Involvement of ultrafine particles?* Inhalation Toxicology, 1995. **7**(1): p. 111-124.
79. U.S. EPA, *Supplemental guidance for developing soil screening levels for superfund sites*. OSWER 9355.4-24. 2002, Washington D.C.: Office of Solid Waste and Emergency Response.
80. U.S. EPA, *AP 42, Fifth Edition, Volume I, Chapter 13: Miscellaneous Sources*. 1995, Washington D.C.: Office of Air Quality Planning and Standards (OAQPS).
81. Trimbach, J. *Oil in the rubber - new oils for an old product*. in *German Rubber Conference*. 2006. Nuremberg, Germany.
82. Hayes, T.D., et al., eds. *Gas Research Institute: Management of Manufactured Gas Plant Sites. Vol I, II*. 1996, Amherst Scientific Publishers: Amherst, MA.
83. Suh, H., P. Koutrakis, and S. Ebelt, *Detailed characterization of indoor personal particulate matter concentrations. Prepared for California Air Resources Board. Final Report Contract No. 00-302*. 2004, Boston, MA: Harvard School of Public Health.
84. Ligman, B., et al., *Airborne particulate matter within school environments in the United States*. Proceedings of Indoor Air, 1999. **IV**: p. 255-261.
85. Kumata, H., et al., *Historical Trends of N-Cyclohexyl-2-benzothiazolamine, 2-(4-Morpholinyl)benzothiazole, and Other Anthropogenic Contaminants in the Urban Reservoir Sediment Core* Environ Sci Technol, 2000. **34**(2): p. 246-253.

86. Kumata, H., et al., *Benzothiazolamines as tire-derived molecular markers: sorptive behavior in street runoff and application to source apportioning*. Environ Sci Technol, 2002. **36**(4): p. 702-8.
87. Luhana, L., et al., *Measurement of non-exhaust particulate matter*. 2004, Deliverable 8 of European Commission DG TrEn 5th Framework PARTICULATES Project.
88. U.S. EPA. *Particulate Matter: PM Standards*. 2007 [cited 2008 4/24]; Available from: <http://epa.gov/particles/standards.html>.
89. Masano, I., *Report of the health influence of dusts by studded tires - experiment of long-term exposure to dusts in rats*. National Environmental Protection Department, Special Pollution Research Information, 1988. **23**(5): p. 35-42.
90. U.S. EPA, *Methods for derivation of inhalation reference concentrations and application of inhalation dosimetry*. 1994.
91. The Scientific Committee on Cosmetic Products and Non-food Products Intended for Consumers, *The SCCNFP's Notes of Guidance for the Testing of Cosmetic Ingredients and Their Safety Evaluation*. 2003, European Commission.
92. Vermeulen, R., G. Doekes, and H. Kromhout, *Latex allergy risk among the general population due to traffic-related airborne dust?* Epidemiology, 2000. **11**(1): p. 92.
93. Turjanmaa, K., *Allergy to natural rubber latex: A growing problem*. Ann Med, 1994. **26**: p. 297-300.
94. Finley, B.L., D.R. Ownby, and S.M. Hays, *Airborne Tire Particles in the Environment: A Possible Asthma Risk from Latex Proteins?* Human and Ecological Risk Assessment, 2003. **9**: p. 1505-1518.
95. Yunginger, J., et al., *Extractable latex allergens and proteins in disposable medical gloves and other rubber products*. J Allergy Clin Immunol, 1994. **93**: p. 836-842.
96. Laoprasert, N., et al., *Inhalation challenge testing of latex-sensitive health care workers and the effectiveness of laminar flow HEPA-filtered helmets in reducing rhinoconjunctival and asthmatic reactions*. J Allergy Clin Immunol, 1998. **102**(6 Pt 1): p. 998-1004.
97. Environmental Protection Agency, *Natural Rubber Latex Adhesives; Disposition of TCSA Section 21 Petition*. EPA-HQ-OPPT-2008-0273; FRL08368-4. 2008. p. 32573-32577.
98. Taylor, J.S. and Y. Leow, *Cutaneous reactions to rubber*. Rubber Chem & Tech, 2000. **73**: p. 427-485.
99. Chaiear, N., *Health and Safety in the Rubber Industry*. Rapra Review Reports, ed. Rapra Technology Limited. 2001.
100. Sullivan, J.P., *An Assessment of Environmental Toxicity and Potential Contamination from Artificial Turf using Shredded or Crumb Rubber*. 2006.
101. U.S. EPA, *Health Effects Test Guidelines, OPPTS 870.2600, Skin Sensitization*. 1998.
102. Rozman, K.K. and C.D. Klaassen, *Absorption, Distribution, and Excretion of Toxicants*, in *Cassarett and Doull's Toxicology The Basic Science of Poisons*, C.D. Klaassen, Editor. 2001, McGraw-Hill: New York.
103. Ames, B.N., et al., *Carcinogens are mutagens: a simple test system combining liver homogenates for activation and bacteria for detection*. PNAS, 1973. **70**(8): p. 2281-2285.
104. Gualtieri, M., et al., *Toxicity of tire debris extracts on human lung cell line A549*. Toxicology in Vitro, 2005. **19**: p. 1001-1008.

105. Borchardt, S.M., *Outbreak of Methicillin-Resistant Staphylococcus aureus Skin Infections Among High School Athletes in Illinois*, in *Illinois Infectious Disease Report*. 2005, Illinois Department of Public Health,. p. 1.
106. Begier, E.M., et al., *A high-morbidity outbreak of methicillin-resistant Staphylococcus aureus among players on a college football team, facilitated by cosmetic body shaving and turf burns*. Clin Infect Dis, 2004. **39**(10): p. 1446-53.
107. Center for Disease Control and Prevention, *Methicillin-resistant Staphylococcus aureus infections among competitive sports participants- Colorado, Indiana, Pennsylvania, and Los Angeles County, 2000-2003*. Morbidity and Mortality Weekly Report, 2003. **52**: p. 793-5.
108. U.S. EPA, *Catalog of Standard Toxicity Tests for Ecological Risk Assessment*, Office of Solid Waste and Emergency Response, Editor. 1994.
109. Sheehan, P.J., et al., *Evaluating the risk to aquatic ecosystems posed by leachate from tire shred fill in roads using toxicity tests, toxicity identification evaluations, and groundwater modeling*. Environ Toxicol Chem, 2006. **25**(2): p. 400-11.
110. Kallqvist, T., *Environmental risk assessment of artificial turf systems*. 2005, Norwegian Institute for Water Research: Oslo. p. 1-19.
111. Gualtieri, M., et al., *Impact of tire debris on in vitro and in vivo systems*. Part Fibre Toxicol, 2005. **2**(1): p. 1.
112. Wik, A. and G. Dave, *Environmental labeling of car tires--toxicity to Daphnia magna can be used as a screening method*. Chemosphere, 2005. **58**(5): p. 645-51.
113. Wik, A. and G. Dave, *Acute toxicity of leachates of tire wear material to Daphnia magna-- variability and toxic components*. Chemosphere, 2006. **64**(10): p. 1777-84.
114. Zheng, M., et al., *Source apportionment of PM_{2.5} in the southeastern United States using solvent-extractable organic compounds as tracers*. Environ Sci Technol, 2002. **36**(11): p. 2361-2371.
115. Day, K.E., et al., *Toxicity of leachate from automobile tires to aquatic biota*. Chemosphere, 1993. **27**.
116. U.S. EPA, *Risk assessment guidance for Superfund Volume I. Human health evaluation manual (Part A) Interim Final*. EPA/540/1-89/002. 1989, Washington D.C.: Office of Emergency and Remedial Response.
117. Johns, D.M., *Initial evaluation of potential human health risks associated with playing on synthetic turf fields on Bainbridge Island*. 2008, Seattle, WA: Windward Environmental LLC.
118. American Chemical Society. *CAS Registry (SM) and CAS Registry Numbers*. 2008 [cited 2008 4/26]; Available from: <http://www.cas.org/expertise/cascontent/registry/regsys.html>.
119. EPA, U.S. *Integrated Risk Information System (IRIS): A-Z List of substances sorted by date*. 2008 [cited 2008 4/26]; Available from: http://cfpub.epa.gov/ncea/iris/index.cfm?fuseaction=iris.showSubstanceList&list_type=date&view=all.
120. EPA, U.S. *EPA Region III Human health risk assessment: risk based concentration table*. 2008 [cited 2008 4/26]; Available from: <http://www.epa.gov/reg3hwmd/risk/human/index.htm>.

121. U.S. EPA. *Voluntary Children's Chemical Evaluation Program (VCCEP): Basic Information*. 2007 [cited 2008 4/26]; Available from: <http://www.epa.gov/oppt/vccep/pubs/basic.htm#basic3>.
122. Woo, Y., et al., *Use of mechanism-based structure-activity relationships analysis in carcinogenic potential ranking for drinking water disinfection by-products*. Environmental Health Perspectives, 2002. **110 Supplement 1**: p. 75-87.
123. Gold, L., T. Slone, and B. Ames, *Natural and synthetic chemicals in the diet: a critical analysis of possible cancer hazards*, in *Food Safety and Food Quality. Issues in Environmental Science and Technology 15.*, R.E. Hester and R.M. Harrison, Editors. 2001, Royal Society of Chemistry: Cambridge, UK. p. 95-128.
124. World Health Organization. *Summary of Evaluations Performed by the Joint FAO/WHO Expert Committee on Food Additives*. 2003 [cited April 28, 2008]; Available from: http://www.inchem.org/documents/jecfa/jecval/jec_189.htm.
125. Burdock, G.A., ed. *Encyclopedia of Food and Color Additives*. 1997, CRC Press: New York. p. 255.
126. World Health Organization. *Environmental Health Criteria 187: White Spirit (Stoddard Solvent)*. 1996 [cited 2008 4/26]; Available from: www.inchem.org/documents/ehc/ehc/ehc187.htm.

Attachment I: Calculation of outdoor airborne ground rubber concentration from wind dispersion as PM-10.

EPA 2002. Equation E-4 ⁽⁷⁹⁾			
Area of site	8,094	m ²	Large 2-acre field.
Area of site	2.0	acres	Unit conversion.
A (default)	16.2302	unitless	EPA 2002. ⁽⁷⁹⁾ Equation E-4.
B (default)	18.7762	unitless	EPA 2002. Equation E-4.
C (default)	216.108	unitless	EPA 2002. Equation E-4.
Q/C_{Wind}	73.69941495	g/m ² -s per kg/m ³	EPA 2002. Equation E-4.
V (no vegetation)	0	percent	EPA 2002. Equation 4-5.
U _m (default)	4.69	m/s	EPA 2002. Equation 4-5.
z _o (plowed field)	1	cm	Hayes et al. (eds) 1996. ⁽⁸²⁾ Figure C-3-3.
u* (3-mm diameter mode)	1	m/s	Hayes et al. (eds) 1996. Figure C-3-1.
u _t (threshold velocity)	16.38	m/s	Hayes et al. (eds) 1996. Equation C-3.
x	3.09		Hayes et al. (eds) 1996. Equation C-4.
F(x)	0.0034		Hayes et al. (eds) 1996. Figure C-3-2.
C _{wind} = 1/PEF	1.1E-11	kg/m ³	EPA 2002. Equation 4-5.
C _{wind} = 1/PEF	0.01	ug/m ³	Unit conversion.

**Attachment II: Gastric digestion supplemental childhood chronic ingestion assessment:
cancer risk and non-cancer hazard quotient**

Cancer Risk - Screening Assessment

Chemical	Age range considered	Exposure duration (years)	Gastric Digestion Concentration (g/g)	Slope Factor (mg/kg-day) ⁻¹	Slope Factor Source	Excess Lifetime Cancer Risk ^a
Arsenic	3 to 70	Child (16 years)	0.031	9.45	OEHHA	2E-07
Cadmium		+	0.014	0.38	OEHHA	4E-09
Lead		Adult (14 years)	0.71	0.0085	OEHHA	5E-09
Aniline		=				
		30 years total	33.5	0.0057	OEHHA	1E-07
Total						4E-07

Hazard Quotient - Screening Assessment

			Gastric Digestion Concentration	Oral RfF (mg/kg-day)		Maximum Hazard Quotient
Chemical	Age range considered	Exposure duration (years)	(g/g)		RfD Source	
Antimony	3 to 70	Child (16 years) + Adult (14 years) = 30 years total	0.55	0.0004	IRIS	0.002
Arsenic			0.031	0.0003	IRIS	0.0002
Barium			0.44	0.2	IRIS	0.000003
Cadmium			0.014	0.001	IRIS (food)	0.00002
Chromium			0.285	0.003	IRIS	0.0002
Cobalt			0.25	0.02	NCEA P	0.00002
Copper			8	0.04	HEAST	0.0003
Lead			0.71	0.00067	OEHHA	0.002
Molybdenum			0.09	0.005	IRIS	0.00003
Nickel			0.135	0.02	IRIS	0.00001
Selenium			0.09	0.005	IRIS	0.00003
Vandium			0.048	0.001	IRIS	0.0001
Zinc			130	0.3	IRIS	0.0007
Aniline			33.5	0.007	NCEA P	0.008
Captan			2.5	0.13	NCEA P	0.00003
Total					0.013	

- a An age-dependant adjustment factor of 3 for ages 3 to 15 was used to estimate risk, as was done in the initial risk assessment performed by OEHHA. However, this adjustment factor is recommended by the U.S. EPA **only** in cases where the mode of action of the chemical is definitively mutagenic. While this may not be the case for all chemicals considered here, the risk calculation here was modeled after the initial OEHHA risk assessment.

Sample Calculation: Cancer

$$\text{Risk} = \text{CS} \times \text{IR} \times \text{EF} \times \text{ED} \times \text{FI} / (\text{AT} \times \text{BW}) \times (\text{SF} \times \text{ADAF})$$

Chemical	Age range	Concentration (mg/g) ^a	Ingestion rate (g/day) ^b	Exposure Frequency (day/year) ^c	Exposure duration (years)	Fraction Ingested from Sourced	Averaging Time	Average Body Weight (kg) ^d	Slope Factor (mg/kg- day) ^{-1e}	Age- Dependant Adjustment Factor ^f	Risk
		CS	IR	EF	ED	FI	AT	BW	SF	ADAF	R
Arsenic	3 to 5	0.000031	0.1	129	3	100%	25550	17.5	9.45	3	8E-08
	6 to 10	0.000031	0.1	129	5	100%	25550	27.5	9.45	3	8E-08
	11 to 15	0.000031	0.1	129	5	100%	25550	47.5	9.45	3	5E-08
	16 to 18	0.000031	0.1	129	3	100%	25550	65	9.45	1	7E-09
	19 to 70	0.000031	0.05	129	14	100%	25550	70	9.45	1	1E-08
Total					30						2E-07

Notes:

a Based on results from gastric ingestion study by OEHHHA

b Based on EPA recommendation for soil ingestion rate (100 mg/d)

c Exposure frequency is 5 days per week x 4.3 weeks per month x 6 months per year equal to 129 days per year.
This frequency is based on EPA recommendations for outdoor recreational activity and is also used in RIVM and Norwegian oral ingestion risk assessments for ground rubber.

d EPA Children's Exposure Factor Handbook

e California Integrated Waste Management Board, 2007

f An age-dependant adjustment factor of 3 for ages 3 to 15 was used to estimate risk, as was done in the initial risk assessment performed by OEHHHA. However, this adjustment factor is recommended by the U.S. EPA in their Supplemental Guidance for Assessing Susceptibility from Early-Life Exposure to Carcinogens (2005) only in cases where the mode of action of the chemical is definitively mutagenic. While this may not be the case for all chemicals considered here, the risk calculation here was modeled after the initial OEHHHA risk assessment.

Sample Calculation: Non-Cancer

$$HQ = CS \times IR \times EF \times ED \times FI / (AT \times BW) / (RfD)$$

Chemical	Age range	Concentration (mg/g) ^a	Ingestion rate (g/day) ^b	Exposure Frequency (day/year) ^c	Exposure duration (years)	Fraction Ingested from Sourced	Averaging Time	Average Body Weight (kg) ^d	Oral Reference Dose (mg/kg- day)	RdD Source	Hazard Quotient
		CS	IR	EF	ED	FI	AT	BW	RfD	--	HQ
Antimony	3 to 5	0.00055	0.1	129	3	100%	1395	17.5	4.00E-04	IRIS	0.002
	6 to 10	0.00055	0.1	129	5	100%	2325	27.5	4.00E-04	IRIS	0.001
	11 to 15	0.00055	0.1	129	5	100%	2325	47.5	4.00E-04	IRIS	0.001
	16 to 18	0.00055	0.1	129	3	100%	1395	65	4.00E-04	IRIS	0.001
	19 to 70	0.00055	0.05	129	14	100%	6510	70	4.00E-04	IRIS	0.0003
										Maximum	0.002

Notes:

- a Based on results from gastric ingestion study by OEHA
- b Based on EPA recommendation for soil ingestion rate (100 mg/d)
- c Exposure frequency is 5 days per week x 4.3 weeks per month x 6 months per year equal to 129 days per year.
This frequency is based on EPA recommendations for outdoor recreational activity and is also used in RIVM and Norwegian oral ingestion risk assessments for ground rubber.
- d EPA Children's Exposure Factor Handbook

AIR QUALITY SURVEY OF SYNTHETIC TURF FIELDS CONTAINING CRUMB RUBBER INFILL

Prepared for

**New York City Department of
Health and Mental Hygiene**

New York, NY

Prepared by

TRC
Windsor, Connecticut

March 2009

AIR QUALITY SURVEY OF SYNTHETIC TURF FIELDS CONTAINING CRUMB RUBBER INFILL

Prepared for
New York City Department of Health and Mental Hygiene
New York, NY

Prepared by
TRC
Windsor, Connecticut

Report Author:
Karen M. Vetrano, Ph.D.

TRC Certified Industrial Hygienist:
Gary Ritter, CIH

TRC Project No. 153896
March 2009

TRC
21 Griffin Road North
Windsor, Connecticut 06095
Telephone 860-298-9692
Facsimile 860-298-6399

*The research was made possible by a grant from the New York Community Trust and partnership
with the Fund for Public Health in New York.*

TABLE OF CONTENTS

<u>SECTION</u>	<u>PAGE</u>
1.0 INTRODUCTION.....	1-1
2.0 AIR QUALITY SURVEY.....	2-1
2.1 Specific Objectives:.....	2-1
2.2 Air Quality Survey Design.....	2-1
2.3 Data Analysis and Interpretation.....	2-3
2.3.1 Data Analysis.....	2-3
2.3.2 Sampling Results	2-4
2.4 Selection of Constituents of Potential Concern (COPCs).....	2-4
2.5 Discussion	2-5
2.5.1 Sampling Results	2-5
2.5.2 Human Health Risk Assessment.....	2-8
3.0 CONCLUSION	3-1
4.0 REFERENCES.....	4-1

TABLE

Air Sampling and Analytic Methods.....	2-2
--	-----

APPENDICES

A	Sampling Locations
B	Sampling Results
C	Sampling Worksheets

EXECUTIVE SUMMARY

Synthetic turf fields have been installed in many athletic and playing fields throughout New York City (NYC), the United States and the world. Many of the synthetic turf fields contain crumb rubber infill. Crumb rubber consists of recycled, chipped/pulverized, used automobile tires primarily made from styrene butadiene rubber (SBR). Crumb rubber granules contain a variety of chemicals typical in rubber, including semi-volatile organic chemicals (SVOC) such as polycyclic aromatic hydrocarbons (PAH) and volatile organic chemicals (VOC). In addition, crumb rubber may contain some amounts of particulate matter and metals. Recent concern about the potential for exposure to chemicals found in crumb rubber prompted NYC Department of Parks (DPR) to request assistance from the NYC Department of Health and Mental Hygiene (DOHMH). In response to this request, and with a grant awarded by the New York Community Trust, the DOHMH contracted TRC to lead an intensive literature review focusing on the potential exposures and health effects related to synthetic turf fields and to identify gaps in what is known. The findings from the review were released in a report prepared by TRC titled “A Review of the Potential Health and Safety Risks from Synthetic Turf Fields Containing Crumb Rubber Infill” (DOHMH 2008). While potential health effects due to heat exposures were identified, an increased risk for human health effects as a result of ingestion, dermal or inhalation exposure to crumb rubber contaminants of potential concern (COPC) was not identified by the review. The review, however, did identify certain knowledge gaps associated with exposure to synthetic turf fields and specifically recommended that air quality related to crumb rubber fields be assessed in the breathing zones of children.

To address the recommendation in the report, with the grant awarded by the New York Community Trust, DOHMH contracted TRC to conduct an air quality survey (AQS). The purpose of the AQS was to investigate the potential release of contaminants from crumb rubber synthetic turf fields and the subsequent potential exposures in the breathing zones of young children to those airborne contaminants. Although there is potential for ingestion and dermal contact of the crumb rubber infill itself, inhalation exposure would be expected to be a primary route of exposure to any emissions from the synthetic turf.

The AQS consisted of air sampling for a suite of SVOCs (PAHs and benzothiazole), VOCs, metals and particulate matter (PM_{2.5}) at two outdoor crumb rubber athletic fields in NYC; Thomas Jefferson Park (East Harlem, Manhattan) and Mullaly Park (Bronx). These

COPCs were selected based upon studies showing that SBR crumb rubber contains these classes of COPCs (DOHMH 2008). These studies were either direct analyses of the crumb rubber or air quality studies conducted in indoor soccer halls. In the AQS, stationary samplers placed on turf fields were used to take measurements in the breathing zone of young children (three feet above ground surface). Air samples were collected under simulated playing conditions such as a practice soccer game and walking/running around the samplers. Stationary background samples were collected upwind of the field at the same time as the corresponding active field samples. A grass field also located at Mullaly Park was sampled in a manner similar to the synthetic turf fields for comparison purposes. Air sampling was conducted under summer conditions (August 2009) in the late morning to afternoon hours to represent potentially the highest concentrations of VOCs released due to the heating of the fields by the sun.. The AQS results represent the conditions of the day and time when sampling was performed.

The results of the AQS air measurements indicate the following:

- Of 69 VOCs tested, eight VOCs were detected in the air measurements. Although VOCs were detected in the air, there was little evidence of harmful levels at the two sampled synthetic turf fields. Also, there was no consistent pattern to indicate that detected VOCs were associated with the synthetic turf. Similar concentrations were found in the background samples from the comparison grass field and upwind locations.
- For the SVOCs:
 - ◆ None of the 17 PAHs tested were detected in any of the ambient air samples.
 - ◆ Benzothiazole, which is considered a chemical “marker” for synthetic rubber (DOHMH 2008) was not detected in any of the air samples, including background samples.
- Of 10 metals tested, two were detected in the ambient air samples. Only one of these metals, however, was detected in the ambient air samples collected from the synthetic turf fields. Similar concentrations were found in both the grass field and upwind samples.
- Ranges of particulate matter (PM_{2.5}) air concentrations from both turf fields were within the background levels found at the comparison grass field and upwind locations.

An analysis of the air in the breathing zones of children above synthetic turf fields did not show appreciable levels from COPCs contained in the crumb rubber. Therefore, a risk assessment related to actual exposure to children was not warranted from the inhalation route of

exposure. Results from one of the bulk crumb rubber samples collected as part of this project identified an elevated lead level in the synthetic turf field at Thomas Jefferson Park.*

* DPR is currently replacing the field and continuing to investigate the source of the lead contamination. Using protocols developed by DOHMH, DPR has since tested the remaining synthetic turf installations throughout NYC for lead and has not found a lead hazard. Results will be posted on the DPR website at www.nyc.gov/parks when available.

1.0 INTRODUCTION

Synthetic turf fields have been installed in many athletic and playing fields throughout New York City (NYC), the United States and the world. The NYC Department of Parks and Recreation (DPR) provides more than 800 athletic fields around the City for New Yorkers to get exercise and enjoy the outdoors. Of these athletic fields, 94 are made of synthetic turf (89 crumb rubber infill fields and 5 carpet-style turf fields without crumb rubber infill). In addition to the athletic fields, there are 17 play areas (14 with crumb rubber infill and 3 carpet-style).

This project focused on synthetic turf fields with crumb rubber infill. The infill-type synthetic turf fields in NYC parks contain several layers, including:

- A bottom layer composed of geo-textile.
- Middle layers composed of broken stone with plastic perforated pipe for drainage and rubber padding for shock absorbance.
- A top layer composed of carpet with soft, flexible plastic grass.
- Crumb rubber infill made from recycled tires added to the 'grass' layer to provide extra padding, serve as a ballast to hold the carpet down, and keep the grass upright. Sand is sometimes mixed with the crumb rubber.

The crumb rubber infill consists of recycled, chipped/pulverized, used automobile tires primarily made from styrene butadiene rubber (SBR). The tire crumbs are roughly the size of grains of coarse sand and generally are spread two to three inches thick over the field surface and packed between ribbons of green polyethyelene fibers used to simulate grass. Crumb rubber granules contain a variety of chemicals typical in rubber, including semi-volatile organic chemicals (SVOC) such as polycyclic aromatic hydrocarbons (PAH) and volatile organic chemicals (VOC). These chemicals may be released into the breathing zones of users, especially on hot days when turf surface temperatures may be elevated. In addition, crumb rubber may contain some amounts of particulate matter and metals. These particles may become airborne during play and sports activities. Crumb rubber may also be further reduced in size and concentration by mechanical abrasion and wear that comes with use of the fields.

In May 2008, DOHMH released a report prepared by TRC, "A Review of the Potential Health and Safety Risks from Synthetic Turf Fields Containing Crumb Rubber Infill." This report identified several gaps in the current knowledge about potential exposures to COPCs from the crumb rubber in synthetic turf fields. These include:

- Outdoor air concentrations of COPCs on both newly installed and older synthetic turf fields. Most of the data generated have been from indoor synthetic turf facilities.
- Background air concentrations of COPCs in NYC. Many of the COPCs found in crumb rubber are also present in the urban environment, but there is little available data on background levels of these COPCs.

The report made the following recommendations:

- Field operators should measure air concentrations of COPCs and particulate matter above outdoor fields to give more representative data related to use of playing fields in urban parks. Measurements taken on a hot, calm (no wind) day would represent a worst case scenario.
- When conducting air studies over fields with crumb rubber, air measurements should also be taken simultaneously at nearby off-field sites, as well as on natural grass and/or asphalt fields, to provide comparative data on exposures related to urban environments.

The AQS addresses the recommendation to investigate the potential airborne release of contaminants from crumb rubber synthetic turf fields and the subsequent potential exposures in the breathing zones of young children to those airborne contaminants. The AQS also addresses the recommendation to investigate natural air measurements on a grass field for comparison data. Although there is potential for ingestion and dermal contact of the crumb rubber infill itself, inhalation exposure would be expected to be a primary route of exposure to any emissions from the synthetic turf.

2.0 AIR QUALITY SURVEY

2.1 Specific Objectives:

This air quality survey (AQS) had the following objectives:

- 1) Characterize the concentrations of certain SVOCs (PAHs, Benzothiazole), VOCs, metals and particulate matter (PM_{2.5}) at selected crumb rubber fields in NYC, including the Mullaly Park field, a "newer" (<1 year old) synthetic turf field and the Thomas Jefferson Park field, an "older" synthetic turf field (> 3 years old); as well as a grass field (also at Mullaly Park) for comparison purposes. The different ages of the turf fields would potentially provide information relating to the aging effect of the crumb rubber.
- 2) Evaluate the suitability of toxicology data to assess the health risks associated with concentrations measured during the AQS.
- 3) Evaluate the level of potential risk for cancer and non-cancer health effects for those analytes found to be above background levels or of toxicological concern.

2.2 Air Quality Survey Design

The AQS consisted of air sampling for a suite of 18 SVOCs (17 PAHs and benzothiazole), 69 VOCs, 10 metals and particulate matter (PM_{2.5}) at two outdoor crumb rubber athletic fields in NYC, Thomas Jefferson Park (East Harlem, Manhattan) and Mullaly Park (Bronx), and at a comparison grass field in Mullaly Park. These COPCs were selected based on studies showing that SBR crumb rubber contains these classes of COPCs (DOHMH 2008). These studies were either direct analysis of the crumb rubber or air quality studies conducted in indoor soccer halls. In addition, a bulk sample of crumb rubber infill was also collected from each turf field and analyzed for 77 organic compounds (VOCs and some SVOCs) and eight metals for the purpose of matching substances identified in the air samples with the constituents found in the bulk crumb rubber.

AQS background and field air samples were collected using stationary samplers. Field air samples were collected from the crumb rubber fields (Refer to Appendix A for sample locations). Background samples, which consisted of air samples collected at upwind locations adjacent to the crumb rubber fields, were collected at the same time as the corresponding active field air samples. In addition, a grass field also located at Mullaly Park, away from the synthetic turf field, was sampled in a manner similar to the synthetic turf fields for comparison purposes. The stationary background and field samples were collected at 3 feet above ground level to

simulate the breathing zone of a young child (USEPA 2008). Samples were collected under simulated playing conditions during a practice soccer game and walking/running around the samplers. Sampling was conducted under summer conditions in the late morning to afternoon hours in order to capture potentially the highest concentrations of VOCs released due to the heating of the fields. Samples were collected over four separate days. Table 1 provides details about the sampling and analytical methods.

Table 1. Air Sampling and Analytic Methods.

Analytes	Sampling Method/Analytical Method	Sampling Equipment	# of Samples (per Turf field)	# of Upwind Samples (per Turf field)*
VOCs	EPA TO-15	1-liter SUMMA canister placed on field (and upwind of field) at 3 ft above ground	4	2
SVOCs/PAHs	NIOSH 5506	Sampling pump with sorptive media placed on field (and upwind of field) at 3 ft above ground	4	2
Metals	OSHA ID 125	Hi-vol sampler placed on field, (and upwind of field) at 3 ft above ground	4	2
SVOCs/ Benzothiazole	NIOSH 2550	Low flow pumps with Teflon filter and adsorbing media	4	2
PM _{2.5}	Continuous Sampling	Dustrak Model 8520 (TSI Instruments)**	4	2

* Upwind samples were not collected for the grass field.

** The Dustrak uses laser photometry to measure particles from 0.1 to 10 um in diameter. A 2.5 size selective inlet nozzle with an omni-directional probe to reduce wind impact was used. Instrument was zeroed before sampling.

The VOC sampling time was pre-set for 1-hour using SUMMA canisters with a flow controlled inlet. Sampling times for all other substances were 120 minutes in duration. The sampling times were chosen based on the anticipated amount of time a child would spend on the field in any given day. NYC Department of Parks and Recreation schedules field use for 1 to 2 hours per permit. Fields permits are in high demand and permit times are limited accordingly.

Field parameters such as the site name, type and age of field, field description, sampling start and end times, date, sampling location on and off the field, ambient and surface temperatures, relative humidity, wind direction/speed and weather conditions were documented. Variables that could provide insight during data interpretation, such as high traffic conditions in adjacent roads were also documented. Field worksheets with the above data can be found in Appendix C.

2.3 Data Analysis and Interpretation

Data were logged into Excel spreadsheets by uniquely coded sample numbers so that all measurements for a single sample appear on the same line. The raw sample data and summary statistics tables can be found in Appendix B. Summary statistics were prepared for each parameter (number of detects, minimum and maximum detected concentration, minimum and maximum detection limit and arithmetic average).

2.3.1 Data Analysis

In order to organize the data into a form manageable and appropriate for risk assessment, data usability was evaluated following USEPA's protocol given in *Guidance for Data Usability in Risk Assessment* (USEPA 1992). The following steps were followed during the data evaluation process as described by USEPA (1989):

- 1) Gather and Sort All Data by Medium (*i.e.*, air and bulk sample).
- 2) Evaluate Methods of Analysis - Analytical methods were evaluated to determine which ones are appropriate for use in the quantitative risk assessment. In doing so, the specificity of the results, the sensitivity of the analytical methods, and the use of adequate quality assurance/quality control (QA/QC) procedures are considered.
- 3) Evaluate the Sample Quantitation Limits (SQL) - For the purpose of the evaluation, all non-detects were evaluated, not simply omitted.
- 4) Evaluate the Data Qualifiers and Codes - Data validation qualifiers were also assessed during the data evaluation process. As indicated in USEPA guidance (USEPA 1989), unqualified data and data qualified with a "J" qualifier are treated as detected concentrations. Data qualified with "UJ" or "U" qualifiers are treated as non-detectable concentrations. As described above, non-detects will be assigned a value equal to the SQL. Data for constituents not detected in any medium or rejected data (qualified with an "R") were not included in the quantitative Human Health Risk Assessment.
- 5) Evaluate Blank Data - Field, trip and laboratory blanks were used to segregate actual site contamination from cross contamination associated with field or laboratory procedures. As indicated in USEPA guidance (1989), sample results are considered positive only if concentrations exceeded ten times the concentration of a common laboratory contaminant in a blank, or five times the concentration of a chemical that is not considered a common laboratory contaminant. Definitions of common laboratory contaminants are provided in USEPA guidance (1989). If less than five or ten times the blank concentration, the constituent will be treated as non-detected in that sample.

- 6) Evaluate Background Data – Site-specific upwind (background) locations were sampled. These site background samples were used as a screening method to evaluate whether constituents detected from within the study area are non-site related.
- 7) Develop Data Sets by Medium - Tables were designed to provide summary statistics (*i.e.*, frequency and range of detects) for constituents detected in air. Full summary statistic tables are provided in Appendix B.

2.3.2 Sampling Results

Summary tables in Appendix B present the findings from the air sampling and bulk crumb rubber analysis. Air sampling was conducted at Mullaly Park's "newer" (<1 yr) synthetic turf field and at Thomas Jefferson Park's "older" (>3 yr) synthetic turf field. Background air samples were collected from upwind locations at Mullaly and Thomas Jefferson Parks and a grass baseball field at Mullaly Park. A bulk sample of crumb rubber infill was collected from each of the turf fields. The data presented represent those compounds that were detected at these fields and at their corresponding background locations given the conditions on the day when sampling was performed. These samples were collected on hot summer days with ambient temperatures ranging from approximately 79°F to 94°F (Appendix B). The surface temperatures on these days ranged from approximately 80°F to 129°F. Of the 18 SVOCs (17 PAHs and benzothiazole), 69 VOCs and 10 metals tested, a total of eight VOCs and two metals were detected in the air measurements as discussed in detail below. Ranges of PM_{2.5} air concentrations from both turf fields were within background levels. Results from one of the bulk crumb rubber samples collected as part of this project identified an elevated lead level in the synthetic turf field at Thomas Jefferson Park.

2.4 Selection of Constituents of Potential Concern (COPCs)

A selection process was used to identify and target site-related COPCs that were likely to contribute significantly to the estimates of risk. Constituents were omitted from the list of COPCs if the:

- Constituent was not detected in any sample;
- Detected air concentrations were present at levels less than the NYS DEC's DAR-1 annual guideline levels (NYSDEC 2007). These screening values are considered conservative screening measures as they assume long-term exposure;

- Detected air concentrations fell within the range measured in the background locations or appear to be from a source unrelated to the synthetic turf.

2.5 Discussion

2.5.1 Sampling Results

Air

Volatile Organic Compounds (VOCs)

Of 69 VOCs tested, eight were detected in the air measurements (Appendix B). Of these eight, three of the VOCs (2-butanone, chloroform and n-hexane) were unique to the synthetic turf samples (*i.e.* not detected in the upwind background locations or at the Mullaly Park grass field). Detected concentrations of 2-butanone and n-hexane were well below the respective guideline values of 5,000 $\mu\text{g}/\text{m}^3$ and 700 $\mu\text{g}/\text{m}^3$, respectively. The detected concentration of chloroform (1 out of 4 samples from Thomas Jefferson Park), however, exceeded its guideline (2.9 $\mu\text{g}/\text{m}^3$ vs. 0.043 $\mu\text{g}/\text{m}^3$). Chloroform has been associated with crumb rubber through direct analysis of the rubber (see DOHMH 2008); however, it was not detected in the analysis of crumb rubber from the Thomas Jefferson Park synthetic turf field. In addition, it has not been detected in air emissions from indoor synthetic turf fields (see DOHMH 2008) suggesting a source other than the crumb rubber for the chloroform. Small amounts of chloroform are formed when chlorine is added to water. A chlorinated swimming pool is located adjacent to the field in which the chloroform was detected, and is a likely source of the compound. Although this reading exceeds the NYS DEC annual guideline concentration, it is far below the short-term guideline concentration for chloroform, 150 $\mu\text{g}/\text{m}^3$.

Five of the 69 VOCs (acetone, chloromethane, ethanol, toluene and methylene chloride) were detected both in the synthetic turf field samples as well as in the upwind background samples and/or the Mullaly Park grass field. Detected concentrations of acetone, chloromethane, ethanol, and toluene did not exceed the respective guideline values of 28,000 $\mu\text{g}/\text{m}^3$, 90 $\mu\text{g}/\text{m}^3$, 45,000 $\mu\text{g}/\text{m}^3$ and 5,000 $\mu\text{g}/\text{m}^3$, respectively. Detected concentrations of methylene chloride from both Mullaly Park (synthetic turf field) and Thomas Jefferson Park, as well as one upwind background sample, exceeded the guideline of 2.1 $\mu\text{g}/\text{m}^3$. The maximum detected concentration was at Thomas Jefferson Park at a concentration of 9 $\mu\text{g}/\text{m}^3$. Methylene chloride is a common laboratory contaminant. It was detected in a majority of the samples including background locations at consistent concentrations. Although the SUMMA canister methodology used for the sample collection does not allow for a blank comparison, the consistency of the methylene

chloride concentrations suggests the presence of laboratory contamination. Methylene chloride has been associated with crumb rubber through direct analysis of the rubber (see DOHMH 2008), however, it was not detected in the analysis of the crumb rubber sample from the Mullaly Park synthetic turf field or the Thomas Jefferson Park synthetic turf field. In addition, it has not been detected in air emissions from indoor synthetic turf fields (see DOHMH 2008).

In addition to the 69 VOCs that were detected as a result of using a standardized analytical method, seven VOC TICs (Tentatively Identified Compounds) were detected in the air measurements. TICs are analytes that the laboratory instrument can detect, but unlike the panel of 69 VOCs the TIC results cannot be verified by the analytic method. Consequently, the TIC findings are merely estimated levels that were detected as part of the analysis. Of the seven VOC TICs detected, four VOC TICs (isobutane, pentane, 2-methyl-1,3-butadiene, and 2-methylbutane) were unique to the synthetic turf fields (*i.e.* not detected in the upwind background locations or at the Mullaly Park grass field). Detected concentrations of 2-methylbutane, isobutane and pentane were well below the respective guideline values of 42,000 $\mu\text{g}/\text{m}^3$, 57,000 $\mu\text{g}/\text{m}^3$ and 4,200 $\mu\text{g}/\text{m}^3$, respectively. 2-Methyl-1,3-butadiene does not have a screening criterion. 2-Methyl-1,3-butadiene also known as isoprene is a common hydrocarbon in animals and plants. It is also found in naturally occurring rubbers. Since 2-methyl-1,3-butadiene was only identified in one sample as a tentatively identified compound and it was not detected in the bulk rubber sample it is not considered a constituent of potential concern. Acetaldehyde was detected in one of the synthetic turf field air measurements as well as in a sample collected from the grass field. Though the measured concentration from the turf field (1.8 $\mu\text{g}/\text{m}^3$) exceeded the respective guideline value of 0.45 $\mu\text{g}/\text{m}^3$, the level was close to the background measured concentration collected from the grass field (1.1 $\mu\text{g}/\text{m}^3$). Acetaldehyde, being that it is a product of combustion including automobile exhaust, is ubiquitous in an urban environment. Although this reading exceeds the NYS DEC annual guideline concentration, it is far below the short-term guideline concentration for acetaldehyde, 4,500 $\mu\text{g}/\text{m}^3$.

Semi-Volatile Organic Compounds (SVOCs)

None of the 18 SVOCs (17 PAHs and benzothiazole) tested were detected in any of the ambient air samples.

Metals

Of 10 metals tested, two (chromium, zinc) were detected in the ambient air samples. However, only one of these metals (chromium) was detected in the ambient air samples obtained from the synthetic turf fields. Similar concentrations were found in the background samples.

Detected concentrations of chromium from the Mullaly Park and Thomas Jefferson Park synthetic turf fields, the Mullaly Park grass field and the two upwind samples all exceeded the guideline value of $1.2 \mu\text{g}/\text{m}^3$. The concentrations of chromium, however, were consistent among all five sample locations. In addition, chromium was detected in a blank sample at $0.65 \mu\text{g}/\text{m}^3$. As indicated in USEPA guidance (1989), sample results are considered positive only if concentrations exceeded ten times the concentration of a common laboratory contaminant in a blank, or five times the concentration of a chemical that is not considered a common laboratory contaminant. Since the detected concentrations of chromium are less than five times the concentration in the blank, the detections are not considered to be positive for chromium. Zinc was detected in a single ambient air sample from the Mullaly Park Grass field at a concentration of $83 \mu\text{g}/\text{m}^3$. This concentration was above the screening criteria of $45 \mu\text{g}/\text{m}^3$. However, it was not detected in any of the samples from the synthetic turf fields or the upwind background samples. The screening levels are 'protective' of long-term, generally continuous exposures. Exposure during the limited time (2 hours per day) spent at any of the playing fields is not expected to be a concern for health effects.

Particulate Matter (PM_{2.5})

Air concentrations of PM_{2.5} at the synthetic turf fields ranged from $0.003 \text{ mg}/\text{m}^3$ to $0.048 \text{ mg}/\text{m}^3$. Background air concentrations of PM_{2.5}, which include the sampling at Mullaly Park's grass field and the specific background samples taken at Mullaly Park and Thomas Jefferson Park, ranged from 0.003 to $0.05 \text{ mg}/\text{m}^3$. The range of PM_{2.5} measurements taken at the synthetic turf fields are within those measured for the grass playing field and the upwind background locations. The primary source of PM_{2.5} is fossil fuel combustion from stationary sources, such as oil-fired power plants, and mobile sources, such as diesel vehicles. Certain industrial sources, e.g. smelting, and large wildfires, also emit fine particulate matter. Since ranges of PM_{2.5} air concentrations from both parks' turf fields are within background levels, and due to the urbanized location of the parks, it is concluded the PM_{2.5} levels from the synthetic turf fields were not distinguishable from background. The consistent measurement of PM_{2.5} at Thomas

Jefferson Park, which is located adjacent to a highway, on a day when the wind was blowing steadily from the direction of the roadway, supports this conclusion.

Bulk Crumb Rubber

A bulk crumb rubber sample was collected from each of the two synthetic turf fields for the purpose of matching substances identified in the air samples with the constituents found in the bulk crumb rubber. The crumb rubber samples were analyzed for 77 organic compounds (VOCs and some SVOCs) and eight metals. Of the organic compounds tested, only one (naphthalene) was detected in the crumb rubber sample collected from the Thomas Jefferson Park synthetic turf field. Detected concentration of the naphthalene was well below the soil cleanup objective level of 100 mg/kg (Appendix B). Concentrations of metals, other than lead and zinc, were well below the respective guideline values. The lead and zinc level for the crumb rubber sample collected from the Thomas Jefferson Park exceeded the respective guideline values of 400 mg/kg and 10,000 mg/kg, respectively. NYS DEC caps the soil cleanup objective values for metals at 10,000 mg/kg. This is not based on health concerns. As zinc is a known component of tires and crumb rubber, a level of 13,100 mg/kg zinc is not at all unexpected. The elevated level of lead detected in the bulk crumb rubber sample from the Thomas Jefferson Park synthetic turf field requires further investigation.

2.5.2 Human Health Risk Assessment

The data does not support the need to conduct a human health risk assessment from the inhalation route of exposure.

3.0 CONCLUSION

The purpose of this AQS was to investigate the potential release of contaminants from crumb rubber synthetic turf fields and the subsequent potential exposures in the breathing zones of young children to those airborne contaminants. Very few constituents were detected in the air samples taken above the fields, and fewer still exceeded the screening levels. Of the 18 SVOCs (17 PAHs and benzothiazole), 69 VOCs and 10 metals tested, a total of eight VOCs and two metals were detected in the air measurements. Of these, only three VOCs were found unique to the synthetic turf fields (*i.e.* not detected in the upwind background locations or at the Mullaly Park grass field); only one of which (chloroform) exceeded the screening level. There was no consistent pattern to indicate that the constituents were associated with the synthetic turf. Regardless, the screening levels are highly conservative and “protective” of long-term, generally continuous exposures, and such continuous and long-term exposures are unlikely to occur at synthetic turf fields. Ranges of PM_{2.5} air concentrations from both synthetic turf fields were within background levels. Overall, none of the detected air measurements were found to be at a level that is likely to cause adverse health effects from typical exposures that occur at synthetic turf fields.

In summary, an analysis of the air in the breathing zones of children above synthetic turf fields do not show appreciable impacts from COPCs contained in the crumb rubber. Therefore, a risk assessment was not warranted from the inhalation route of exposure. The bulk crumb rubber samples collected as part of this project, however, resulted in the detection of an elevated lead level.[†]

[†] DPR is currently replacing the field and continuing to investigate the source of the lead contamination. Using protocols developed by DOHMH, DPR has since tested the remaining synthetic turf installations throughout NYC for lead and has not found a lead hazard. Results will be posted on the DPR website at www.nyc.gov/parks when available.

4.0 REFERENCES

- New York City Department of Health and Mental Hygiene (DOHMH). 2008.** A Review of the Potential Health and Safety Risks From Synthetic Turf Fields Containing Crumb Rubber Infill. Prepared by TRC for DOHMH. May.
- New York State Department of Environmental Conservation (NYS DEC). 2007.** Guidelines for the Control of Toxic Ambient Air Contaminants. DAR-1. Division of Air Resources. September 10.
- New York State Department of Environmental Conservation (NYS DEC) and New York State Department of Health (NYS DOH). 2006.** New York State Brownfield Cleanup Program Development of Soil Cleanup Objectives Technical Support Document, New York State Department of Environmental Conservation and New York State Department of Health. September.
- USEPA. 1989.** Risk Assessment Guidance for Superfund Volume I: Human Health Evaluation Manual (Part A) Interim Final. Washington, DC: US Environmental Protection Agency Office of Emergency and Remedial Response. EPA/540/1-89/002. December.
- USEPA. 1992.** Guidance for Data Usability in Risk Assessment, Final. US Environmental Protection Agency Office of Emergency and Remedial Response. Washington DC. Publication 9285.7-09A. April.
- USEPA. 2008.** Child-Specific Exposures Factor Handbook. EPA/600/R-06/096F.

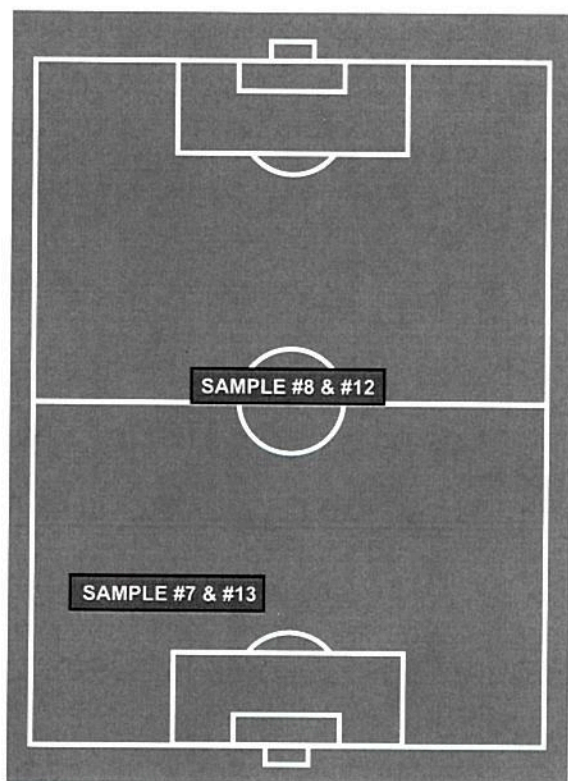
APPENDIX A

SAMPLING LOCATIONS



WIND
DIRECTION

BACKGROUND
SAMPLE #6 & 11



MULLARY PARK

NOT TO SCALE



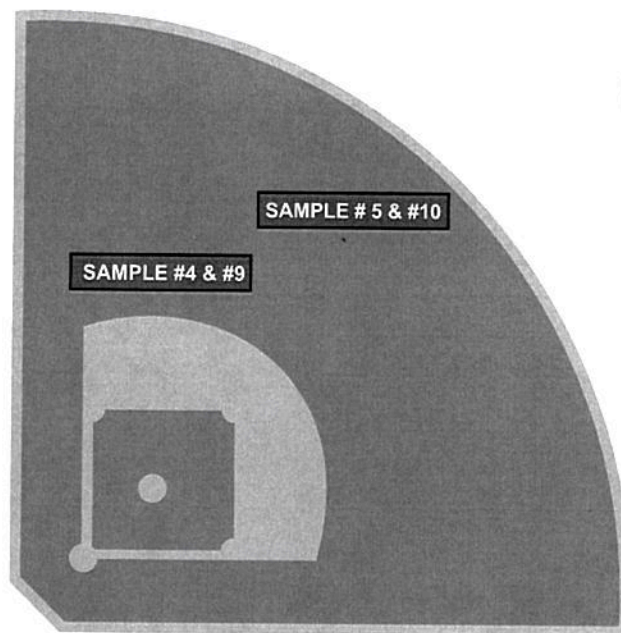
21 Griffin Rd. North
Windsor, CT 06095
(860) 298-9692

MULLARY PARK
NEW YORK, NEW YORK

APPENDIX A-1
SAMPLING LOCATIONS -
TURF FIELD

Date: 09/08

Project No. 153896.0010.0000



NOT TO SCALE



21 Griffin Rd. North
Windsor, CT 06095
(860) 298-9692

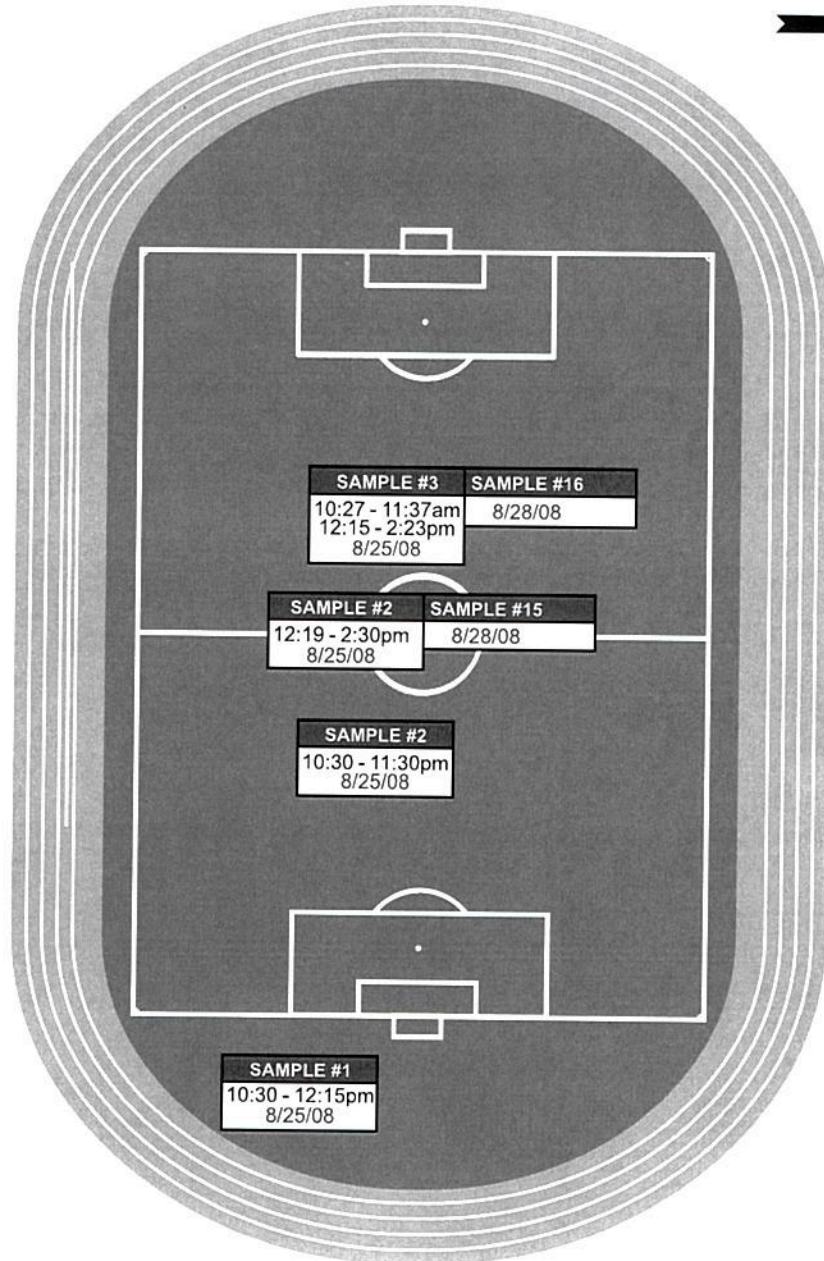
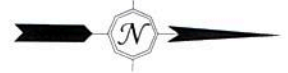
MULLARY PARK
NEW YORK, NEW YORK

**APPENDIX A-2
SAMPLING LOCATIONS -
GRASS FIELD**

Date: 09/08

Project No. 153896.0010.0000

THOMAS JEFFERSON



BACKGROUND
SAMPLE #1
12:29 - 2:34pm
8/25/08

BACKGROUND
SAMPLE #14
8/28/08

Playground

NOT TO SCALE



WIND DIRECTION
AT 10:30 A.M. NE



21 Griffin Rd. North
Windsor, CT 06095
(860) 298-9692

THOMAS JEFFERSON
NEW YORK, NEW YORK

APPENDIX A-3 SAMPLING LOCATIONS - TURF FIELD

Date: 09/08

Project No. 153896.0010.0000

APPENDIX B

SAMPLING RESULTS

Table B-1
Summary Air Sampling Results for Detected Analytes

Analytes	Synthetic Turf Fields (Concentration Range)	Background - Grass/Upwind (Concentration Range)	Annual Air Guideline [†]	Short-Term Air Guideline [†]
	(µg/m ³)	(µg/m ³)	(µg/m ³)	(µg/m ³)
Volatile Organic Compounds or VOCs (Of 69 VOCs tested, eight were detected)				
2-Butanone (MEK)	ND – 3	ND	5,000	13,000
Acetone	9.3 – 51	ND - 11	28,000	180,000
Chloroform	ND - 2.9*	ND	0.043	150
Chloromethane	ND - 1.1	ND - 1.1	90	22,000
Ethanol	6.2 – 22	5.1 - 8.9	45,000	NA
n-Hexane	ND - 2.1	ND	700	NA
Methylene Chloride	ND - 9*	ND - 6.9*	2.1	14,000
Toluene	ND - 2.7	ND - 2	5,000	37,000
Metals (Of 10 metals tested, two were detected)				
Chromium	0.87 - 1.4*	ND - 1.8*	1.2	NA
Zinc	ND	ND - 83	45	NA
Particulate Matter or PM				
PM 2.5	0.003 - 0.048	0.003 - 0.05	15	NA
Semi-Volatile Organic Compounds or SVOCs (Of 18 SVOCs tested which included 17 PAHs and benzothiazole, none were detected in any of the ambient air samples)				
*Measurement exceeded guideline value.				
ND = Not detected above the laboratory reporting limit.				
NA = Not available.				
[†] NYS DEC 2007. DAR 1 Tables – Short-term and annual air guideline levels.				

* Sample ID	MPT7	MPT8	MPT12	MPT13	MPT12a	MPT12b	Lab Blank
Sample Name	Mullaly Park Turf	Mullaly Park Turf	Mullaly Park Turf Dup. Comb.	Mullaly Park Turf	Mullaly Park Turf	Mullaly Park Turf	
Sample Date	08/26/08	08/26/08	08/27/08	08/27/08	08/27/08	08/27/08	
* Sample Type	Air	Air	Air	Air	Air	Air	Air
Lab ID	280801650-14	280801650-13	280801650-7 280801650-3	280801650-6	280801650-7	280801650-3	
Duplicate			Yes				
CONSTITUENTS							
VOCs							
4 Methyl-2-pentanone	2 U	2 U	2 U	2 U	2 U	2 U	
Acetone	11	16	9.25	9.8	9.9	8.6	
Acetonitrile	0.84 U	0.84 U	0.84 U	0.84 U	0.84 U	0.84 U	
Acrylonitrile	1.1 U	1.1 U	1.1 U	1.1 U	1.1 U	1.1 U	
Benzene	1.6 U	1.6 U	1.6 U	1.6 U	1.6 U	1.6 U	
Benzyl Chloride	3.7 U	3.7 U	3.7 U	3.7 U	3.7 U	3.7 U	
Bromodichloromethane	3.3 U	3.3 U	3.3 U	3.3 U	3.3 U	3.3 U	
Bromoethane	2.2 U	2.2 U	2.2 U	2.2 U	2.2 U	2.2 U	
Bromoethene	2.2 U	2.2 U	2.2 U	2.2 U	2.2 U	2.2 U	
Bromoform	5.2 U	5.2 U	5.2 U	5.2 U	5.2 U	5.2 U	
Bromomethane	1.9 U	1.9 U	1.9 U	1.9 U	1.9 U	1.9 U	
Butadiene, 1,3-	1.1 U	1.1 U	1.1 U	1.1 U	1.1 U	1.1 U	
Butanone, 2- (MEK)	1.7	1.6	1.5 U	1.5 U	1.5 U	1.5 U	
Carbon disulfide	1.6 U	1.6 U	1.6 U	1.6 U	1.6 U	1.6 U	
Carbon Tetrachloride	3.1 U	3.1 U	3.1 U	3.1 U	3.1 U	3.1 U	
Chlorobenzene	2.3 U	2.3 U	2.3 U	2.3 U	2.3 U	2.3 U	
Chloroethane	1.3 U	1.3 U	1.3 U	1.3 U	1.3 U	1.3 U	
Chloroform	2.4 U	2.4 U	2.4 U	2.4 U	2.4 U	2.4 U	
Chloromethane	1 U	1	1.1	1.1	1 U	1.1	
Chloropropene, 3-	1.6 U	1.6 U	1.6 U	1.6 U	1.6 U	1.6 U	
Chlorotoluene, 2-	2.6 U	2.6 U	2.6 U	2.6 U	2.6 U	2.6 U	
Cyclohexane	1.7 U	1.7 U	1.7 U	1.7 U	1.7 U	1.7 U	
Dibromochloromethane	4.3 U	4.3 U	4.3 U	4.3 U	4.3 U	4.3 U	
Dibromoethane, 1,2-	3.8 U	3.8 U	3.8 U	3.8 U	3.8 U	3.8 U	
Dichlorobenzene, 1,2-	3 U	3 U	3 U	3 U	3 U	3 U	
Dichlorobenzene, 1,3-	3 U	3 U	3 U	3 U	3 U	3 U	
Dichlorobenzene, 1,4-	3 U	3 U	3 U	3 U	3 U	3 U	
Dichloroethane, 1,1-	2 U	2 U	2 U	2 U	2 U	2 U	
Dichloroethane, 1,2-	2 U	2 U	2 U	2 U	2 U	2 U	
Dichloroethene, 1,1-	2 U	2 U	2 U	2 U	2 U	2 U	
Dichloroethene, cis- 1,2-	2 U	2 U	2 U	2 U	2 U	2 U	
Dichloroethene, trans-1,2-	2 U	2 U	2 U	2 U	2 U	2 U	
Dichloropropane, 1,2-	2.3 U	2.3 U	2.3 U	2.3 U	2.3 U	2.3 U	
Dichloropropene, cis-1,3-	2.3 U	2.3 U	2.3 U	2.3 U	2.3 U	2.3 U	
Dichloropropene, trans-1,3-	2.3 U	2.3 U	2.3 U	2.3 U	2.3 U	2.3 U	
Dioxane, 1,4-	1.8 U	1.8 U	1.8 U	1.8 U	1.8 U	1.8 U	
Ethanol	7	20	6.2	7.6	6.2	6.2	
Ethyl acetate	1.8 U	1.8 U	1.8 U	1.8 U	1.8 U	1.8 U	
Ethylbenzene	2.2 U	2.2 U	2.2 U	2.2 U	2.2 U	2.2 U	
Ethyltoluene, 4-	2.5 U	2.5 U	2.5 U	2.5 U	2.5 U	2.5 U	
Freon 11 (Trichlorofluoromethane)	2.8 U	2.8 U	2.8 U	2.8 U	2.8 U	2.8 U	
Freon 113 (1,1,2 Trichlorotrifluoroethane)	3.8 U	3.8 U	3.8 U	3.8 U	3.8 U	3.8 U	
Freon 114 (1,2 Dichlorotetrafluoroethane)	3.5 U	3.5 U	3.5 U	3.5 U	3.5 U	3.5 U	
Freon 12 (Dichlorodifluoromethane)	2.5 U	2.5 U	2.5 U	2.5 U	2.5 U	2.5 U	
Heptane, n-	2 U	2 U	2 U	2 U	2 U	2 U	
Hexachloro-1,3-butadiene	5.3 U	5.3 U	5.3 U	5.3 U	5.3 U	5.3 U	
Hexane, n-	1.8 U	1.8 U	1.8 U	1.8 U	1.8 U	1.8 U	
Hexanone, 2-	2 U	2 U	2 U	2 U	2 U	2 U	
Isopropyl alcohol	3.7 U	3.7 U	3.7 U	3.7 U	3.7 U	3.7 U	
Methylene chloride	5.2 U	6.6	6	5.2 U	5.2 U	6	
Methyl-tert-butyl ether	1.8 U	1.8 U	1.8 U	1.8 U	1.8 U	1.8 U	
Propylene	1.7 U	1.7 U	1.7 U	1.7 U	1.7 U	1.7 U	
Styrene	2.1 U	2.1 U	2.1 U	2.1 U	2.1 U	2.1 U	
Tertiary butyl alcohol	1.5 U	1.5 U	1.5 U	1.5 U	1.5 U	1.5 U	
Tetrachloroethane, 1,1,2,2-	3.4 U	3.4 U	3.4 U	3.4 U	3.4 U	3.4 U	
Tetrachloroethene	3.4 U	3.4 U	3.4 U	3.4 U	3.4 U	3.4 U	
Tetrahydrofuran	1.5 U	1.5 U	1.5 U	1.5 U	1.5 U	1.5 U	
Toluene	1.9 U	2	1.9 U	1.9 U	1.9 U	1.9 U	
Trichlorobenzene, 1,2,4-	3.7 U	3.7 U	3.7 U	3.7 U	3.7 U	3.7 U	
Trichloroethane, 1,1,1-	2.7 U	2.7 U	2.7 U	2.7 U	2.7 U	2.7 U	
Trichloroethane, 1,1,2-	2.7 U	2.7 U	2.7 U	2.7 U	2.7 U	2.7 U	
Trichloroethene	2.7 U	2.7 U	2.7 U	2.7 U	2.7 U	2.7 U	
Trimethylbenzene, 1,2,4-	2.5 U	2.5 U	2.5 U	2.5 U	2.5 U	2.5 U	
Trimethylbenzene, 1,3,5-	2.5 U	2.5 U	2.5 U	2.5 U	2.5 U	2.5 U	
Trimethylpentane, 2,2,4-	2.3 U	2.3 U	2.3 U	2.3 U	2.3 U	2.3 U	
Vinyl acetate	1.8 U	1.8 U	1.8 U	1.8 U	1.8 U	1.8 U	
Vinyl chloride	1.3 U	1.3 U	1.3 U	1.3 U	1.3 U	1.3 U	
Xylene (ortho)	2.2 U	2.2 U	2.2 U	2.2 U	2.2 U	2.2 U	
Xylene (para & meta)	2.2 U	2.2 U	2.2 U	2.2 U	2.2 U	2.2 U	

MPT12a	MPT12b	Lab Blank
Mullaly Park Turf	Mullaly Park Turf	
08/27/08	08/27/08	
Air	Air	Air
280801650-7	280801650-3	
		0.2 U
		0.2 U
		0.2 U
		0.1 U
		0.1 U
		0.1 U
		0.1 U
		0.1 U
		0.1 U
		0.1 U
		0.2 U
		0.1 U
		0.2 U
		0.1 U
		0.1 U
		0.2 U
		0.65
		1 U
		5 U
		0.5 U
		0.75 U
		1 U
		0.5 U
		5 U
		1 U

Appendix B-1b
Thomas Jefferson Park - Synthetic Turf Field
New York, NY

* Sample ID :	TJ2	TJ3	TJ15	TJ16
Sample Name:	Thomas Jefferson Turf	Thomas Jefferson Turf	Thomas Jefferson Turf	Thomas Jefferson Turf
Sample Date :	08/25/08	08/25/08	08/28/08	08/28/08
* Sample Type :	Air	Air	Air	Air
Lab ID :	280801650-20	280801650-17	280801650-8	280801650-2
Duplicate :				
CONSTITUENTS				
VOCs				
4-Methyl-2-pentanone	2 U	2 U	2 U	2 U
Acetone	11	20	11	51
Acetonitrile	0.84 U	0.84 U	0.84 U	0.84 U
Acrylonitrile	1.1 U	1.1 U	1.1 U	1.1 U
Benzene	1.6 U	1.6 U	1.6 U	1.6 U
Benzyl Chloride	3.7 U	3.7 U	3.7 U	3.7 U
Bromodichloromethane	3.3 U	3.3 U	3.3 U	3.3 U
Bromoethane	2.2 U	2.2 U	2.2 U	2.2 U
Bromoethene	2.2 U	2.2 U	2.2 U	2.2 U
Bromoform	5.2 U	5.2 U	5.2 U	5.2 U
Bromomethane	1.9 U	1.9 U	1.9 U	1.9 U
Butadiene, 1,3-	1.1 U	1.1 U	1.1 U	1.1 U
Butanone, 2- (MEK)	2.4	3	1.6	1.5
Carbon disulfide	1.6 U	1.6 U	1.6 U	1.6 U
Carbon Tetrachloride	3.1 U	3.1 U	3.1 U	3.1 U
Chlorobenzene	2.3 U	2.3 U	2.3 U	2.3 U
Chloroethane	1.3 U	1.3 U	1.3 U	1.3 U
Chloroform	2.4 U	2.4 U	2.4 U	2.9
Chloromethane	1.1	1 U	1 U	1 U
Chloropropene, 3-	1.6 U	1.6 U	1.6 U	1.6 U
Chlorotoluene, 2-	2.6 U	2.6 U	2.6 U	2.6 U
Cyclohexane	1.7 U	1.7 U	1.7 U	1.7 U
Dibromochloromethane	4.3 U	4.3 U	4.3 U	4.3 U
Dibromoethane, 1,2-	3.8 U	3.8 U	3.8 U	3.8 U
Dichlorobenzene, 1,2-	3 U	3 U	3 U	3 U
Dichlorobenzene, 1,3-	3 U	3 U	3 U	3 U
Dichlorobenzene, 1,4-	3 U	3 U	3 U	3 U
Dichloroethane, 1,1-	2 U	2 U	2 U	2 U
Dichloroethane, 1,2-	2 U	2 U	2 U	2 U
Dichloroethene, 1,1-	2 U	2 U	2 U	2 U
Dichloroethene, cis-1,2-	2 U	2 U	2 U	2 U
Dichloroethene, trans-1,2-	2 U	2 U	2 U	2 U
Dichloropropane, 1,2-	2.3 U	2.3 U	2.3 U	2.3 U
Dichloropropene, cis-1,3-	2.3 U	2.3 U	2.3 U	2.3 U
Dichloropropene, trans-1,3-	2.3 U	2.3 U	2.3 U	2.3 U
Dioxane, 1,4-	1.8 U	1.8 U	1.8 U	1.8 U
Ethanol	6.8	12	7.6	22
Ethyl acetate	1.8 U	1.8 U	1.8 U	1.8 U
Ethylbenzene	2.2 U	2.2 U	2.2 U	2.2 U
Ethyltoluene, 4-	2.5 U	2.5 U	2.5 U	2.5 U
Freon 11 (Trichlorofluoromethane)	2.8 U	2.8 U	2.8 U	2.8 U
Freon 113 (1,1,2 Trichlorotrifluoroethane)	3.8 U	3.8 U	3.8 U	3.8 U
Freon 114 (1,2 Dichlorotetrafluoroethane)	3.5 U	3.5 U	3.5 U	3.5 U
Freon 12 (Dichlorodifluoromethane)	2.5 U	2.5 U	2.5 U	2.5 U
Heptane, n-	2 U	2 U	2 U	2 U
Hexachloro-1,3-butadiene	5.3 U	5.3 U	5.3 U	5.3 U
Hexane, n-	1.8 U	1.8 U	2.1	1.8 U
Hexanone, 2-	2 U	2 U	2 U	2 U
Isopropyl alcohol	3.7 U	3.7 U	3.7 U	3.7 U
Methylene chloride	5.2 U	5.2 U	5.2 U	9
Methyl-tert-butyl ether	1.8 U	1.8 U	1.8 U	1.8 U
Propylene	1.7 U	1.7 U	1.7 U	1.7 U
Styrene	2.1 U	2.1 U	2.1 U	2.1 U
Tertiary butyl alcohol	1.5 U	1.5 U	1.5 U	1.5 U
Tetrachloroethane, 1,1,2,2-	3.4 U	3.4 U	3.4 U	3.4 U
Tetrachloroethene	3.4 U	3.4 U	3.4 U	3.4 U
Tetrahydrofuran	1.5 U	1.5 U	1.5 U	1.5 U
Toluene	1.9 U	1.9	1.9 U	2.7
Trichlorobenzene, 1,2,4-	3.7 U	3.7 U	3.7 U	3.7 U
Trichloroethane, 1,1,1-	2.7 U	2.7 U	2.7 U	2.7 U
Trichloroethane, 1,1,2-	2.7 U	2.7 U	2.7 U	2.7 U
Trichloroethene	2.7 U	2.7 U	2.7 U	2.7 U
Trimethylbenzene, 1,2,4-	2.5 U	2.5 U	2.5 U	2.5 U
Trimethylbenzene, 1,3,5-	2.5 U	2.5 U	2.5 U	2.5 U
Trimethylpentane, 2,2,4-	2.3 U	2.3 U	2.3 U	2.3 U
Vinyl acetate	1.8 U	1.8 U	1.8 U	1.8 U
Vinyl chloride	1.3 U	1.3 U	1.3 U	1.3 U
Xylene (ortho)	2.2 U	2.2 U	2.2 U	2.2 U
Xylene (para & meta)	2.2 U	2.2 U	2.2 U	2.2 U

Appendix B-1b
Thomas Jefferson Park - Synthetic Turf Field
New York, NY

* Sample ID :	T12	T13	T15	T16
Sample Name :	Thomas Jefferson Turf	Thomas Jefferson Turf	Thomas Jefferson Turf	Thomas Jefferson Turf
Sample Date :	08/25/08	08/25/08	08/28/08	08/28/08
* Sample Type :	Air	Air	Air	Air
Lab ID :	280801650-20	280801650-17	280801650-8	280801650-2
Duplicate :				
VOC TICs				
Acetaldehyde		1.80 J		J
Butane, 2-methyl				2.95 J
Isobutane				2.38 J
Methyl-1,3-Butadiene, 2-			2.79 J	
Pentane				8.85
SVOCs				
Acenaphthene	0.51 U	0.49 U	0.75 U	0.77 U
Acenaphthylene	0.51 U	0.49 U	0.75 U	0.77 U
Anthracene	0.51 U	0.49 U	0.75 U	0.77 U
Benzo(a)anthracene	0.25 U	0.25 U	0.38 U	0.39 U
Benzo(a)pyrene	0.25 U	0.25 U	0.38 U	0.39 U
Benzo(b)fluoranthene	0.25 U	0.25 U	0.38 U	0.39 U
Benzo(e)pyrene	0.25 U	0.25 U	0.38 U	0.39 U
Benzo(g,h,i)perylene	0.25 U	0.25 U	0.38 U	0.39 U
Benzo(k)fluoranthene	0.25 U	0.25 U	0.38 U	0.39 U
Chrysene	0.25 U	0.25 U	0.38 U	0.39 U
Dibenzo(a,h)anthracene	0.25 U	0.25 U	0.38 U	0.39 U
Fluoranthene	0.25 U	0.25 U	0.38 U	0.39 U
Fluorene	0.51 U	0.49 U	0.75 U	0.77 U
Indeno(1,2,3-c,d)pyrene	0.25 U	0.25 U	0.38 U	0.39 U
Naphthalene	0.51 U	0.49 U	0.75 U	0.77 U
Phenanthrene	0.25 U	0.25 U	0.38 U	0.39 U
Pyrene	0.25 U	0.25 U	0.38 U	0.39 U
Benzothiazole	0.25 U	0.25 U	0.37 U	0.38 U
Metals				
Cadmium	0.25 U	0.25 U	0.37 U	0.39 U
Chromium	1.1	0.87	1.4	1.1
Copper	1.2 U	1.3 U	1.8 U	1.9 U
Iron	6.2 U	6.3 U	9.1 U	9.7 U
Lead	0.62 U	0.63 U	0.91 U	0.97 U
Manganese	0.93 U	0.94 U	1.4 U	1.5 U
Nickel	1.2 U	1.3 U	1.8 U	1.9 U
Silver	0.62 U	0.63 U	0.91 U	0.97 U
Tin	6.2 U	6.3 U	9.1 U	9.7 U
Zinc	1.2 U	1.3 U	1.8 U	1.9 U

* Sample ID :	MPF4	MPF5	MPF9	MPF10
Sample Name :	Mullaly Park Grass Field	Mullaly Park Grass Field	Mullaly Park Grass Field	Mullaly Park Grass Field
Sample Date :	08/26/08	08/26/08	08/27/08	08/27/08
* Sample Type :	Air	Air	Air	Air
Lab ID :	280801650-21	280801650-15	280801650-4	280801650-19
Duplicate :				
CONSTITUENTS				
VOCs				
4-Methyl-2-pentanone	2 U	2 U	2 U	2 U
Acetone	9.5	7.1 U	9.2	7.3
Acetonitrile	0.84 U	0.84 U	0.84 U	0.84 U
Acrylonitrile	1.1 U	1.1 U	1.1 U	1.1 U
Benzene	1.6 U	1.6 U	1.6 U	1.6 U
Benzyl Chloride	3.7 U	3.7 U	3.7 U	3.7 U
Bromodichloromethane	3.3 U	3.3 U	3.3 U	3.3 U
Bromoethane	2.2 U	2.2 U	2.2 U	2.2 U
Bromoethene	2.2 U	2.2 U	2.2 U	2.2 U
Bromoform	5.2 U	5.2 U	5.2 U	5.2 U
Bromomethane	1.9 U	1.9 U	1.9 U	1.9 U
Butadiene, 1,3-	1.1 U	1.1 U	1.1 U	1.1 U
Butanone, 2- (MEK)	1.5 U	1.5 U	1.5 U	1.5 U
Carbon disulfide	1.6 U	1.6 U	1.6 U	1.6 U
Carbon Tetrachloride	3.1 U	3.1 U	3.1 U	3.1 U
Chlorobenzene	2.3 U	2.3 U	2.3 U	2.3 U
Chloroethane	1.3 U	1.3 U	1.3 U	1.3 U
Chloroform	2.4 U	2.4 U	2.4 U	2.4 U
Chloromethane	1 U	1.1	1 U	1 U
Chloropropene, 3-	1.6 U	1.6 U	1.6 U	1.6 U
Chlorotoluene, 2-	2.6 U	2.6 U	2.6 U	2.6 U
Cyclohexane	1.7 U	1.7 U	1.7 U	1.7 U
Dibromochloromethane	4.3 U	4.3 U	4.3 U	4.3 U
Dibromoethane, 1,2-	3.8 U	3.8 U	3.8 U	3.8 U
Dichlorobenzene, 1,2-	3 U	3 U	3 U	3 U
Dichlorobenzene, 1,3-	3 U	3 U	3 U	3 U
Dichlorobenzene, 1,4-	3 U	3 U	3 U	3 U
Dichloroethane, 1,1-	2 U	2 U	2 U	2 U
Dichloroethane, 1,2-	2 U	2 U	2 U	2 U
Dichloroethene, 1,1-	2 U	2 U	2 U	2 U
Dichloroethene, cis- 1,2-	2 U	2 U	2 U	2 U
Dichloroethene, trans-1,2-	2 U	2 U	2 U	2 U
Dichloropropane, 1,2-	2.3 U	2.3 U	2.3 U	2.3 U
Dichloropropene, cis-1,3-	2.3 U	2.3 U	2.3 U	2.3 U
Dichloropropene, trans-1,3-	2.3 U	2.3 U	2.3 U	2.3 U
Dioxane, 1,4-	1.8 U	1.8 U	1.8 U	1.8 U
Ethanol	5.8	5.1	7.6	6.4
Ethyl acetate	1.8 U	1.8 U	1.8 U	1.8 U
Ethylbenzene	2.2 U	2.2 U	2.2 U	2.2 U
Ethyltoluene, 4-	2.5 U	2.5 U	2.5 U	2.5 U
Freon 11 (Trichlorofluoromethane)	2.8 U	2.8 U	2.8 U	2.8 U
Freon 113 (1,1,2 Trichlorotrifluoroethane)	3.8 U	3.8 U	3.8 U	3.8 U
Freon 114 (1,2 Dichlorotetrafluoroethane)	3.5 U	3.5 U	3.5 U	3.5 U
Freon 12 (Dichlorodifluoromethane)	2.5 U	2.5 U	2.5 U	2.5 U
Heptane, n-	2 U	2 U	2 U	2 U
Hexachloro-1,3-butadiene	5.3 U	5.3 U	5.3 U	5.3 U
Hexane, n-	1.8 U	1.8 U	1.8 U	1.8 U
Hexanone, 2-	2 U	2 U	2 U	2 U
Isopropyl alcohol	3.7 U	3.7 U	3.7 U	3.7 U
Methylene chloride	5.2 U	5.2 U	5.2 U	5.2 U
Methyl-tert-butyl ether	1.8 U	1.8 U	1.8 U	1.8 U
Propylene	1.7 U	1.7 U	1.7 U	1.7 U
Styrene	2.1 U	2.1 U	2.1 U	2.1 U
Tertiary butyl alcohol	1.5 U	1.5 U	1.5 U	1.5 U
Tetrachloroethane, 1,1,2,2-	3.4 U	3.4 U	3.4 U	3.4 U
Tetrachloroethene	3.4 U	3.4 U	3.4 U	3.4 U
Tetrahydrofuran	1.5 U	1.5 U	1.5 U	1.5 U
Toluene	1.9 U	1.9 U	1.9 U	1.9 U
Trichlorobenzene, 1,2,4-	3.7 U	3.7 U	3.7 U	3.7 U
Trichloroethane, 1,1,1-	2.7 U	2.7 U	2.7 U	2.7 U
Trichloroethane, 1,1,2-	2.7 U	2.7 U	2.7 U	2.7 U
Trichloroethene	2.7 U	2.7 U	2.7 U	2.7 U
Trimethylbenzene, 1,2,4-	2.5 U	2.5 U	2.5 U	2.5 U
Trimethylbenzene, 1,3,5-	2.5 U	2.5 U	2.5 U	2.5 U
Trimethylpentane, 2,2,4-	2.3 U	2.3 U	2.3 U	2.3 U
Vinyl acetate	1.8 U	1.8 U	1.8 U	1.8 U
Vinyl chloride	1.3 U	1.3 U	1.3 U	1.3 U
Xylene (ortho)	2.2 U	2.2 U	2.2 U	2.2 U
Xylene (para & meta)	2.2 U	2.2 U	2.2 U	2.2 U

Appendix B-1c
Mullaly Park - Grass Field
New York, NY

* Sample ID :	MPF4	MPF5	MPF9	MPF10
Sample Name :	Mullaly Park Grass Field	Mullaly Park Grass Field	Mullaly Park Grass Field	Mullaly Park Grass Field
Sample Date :	08/26/08	08/26/08	08/27/08	08/27/08
* Sample Type :	Air	Air	Air	Air
Lab ID :	280801650-21	280801650-15	280801650-4	280801650-19
Duplicate :				
VOC TICS				
Acetaldehyde		1.80 J		
Hexanal	4.09 J			
Nonanal	5.81 J			
SVOCs				
Acenaphthene	0.8 U	0.81 U	0.82 U	0.8 U
Acenaphthylene	0.8 U	0.81 U	0.82 U	0.8 U
Anthracene	0.8 U	0.81 U	0.82 U	0.8 U
Benzo(a)anthracene	0.4 U	0.4 U	0.41 U	0.4 U
Benzo(a)pyrene	0.4 U	0.4 U	0.41 U	0.4 U
Benzo(b)fluoranthene	0.4 U	0.4 U	0.41 U	0.4 U
Benzo(e)pyrene	0.4 U	0.4 U	0.41 U	0.4 U
Benzo(g,h,i)perylene	0.4 U	0.4 U	0.41 U	0.4 U
Benzo(k)fluoranthene	0.4 U	0.4 U	0.41 U	0.4 U
Chrysene	0.4 U	0.4 U	0.41 U	0.4 U
Dibenzo(a,h)anthracene	0.4 U	0.4 U	0.41 U	0.4 U
Fluoranthene	0.4 U	0.4 U	0.41 U	0.4 U
Fluorene	0.8 U	0.81 U	0.82 U	0.8 U
Indeno(1,2,3-c,d)pyrene	0.4 U	0.4 U	0.41 U	0.4 U
Naphthalene	0.8 U	0.81 U	0.82 U	0.8 U
Phenanthrene	0.4 U	0.4 U	0.41 U	0.4 U
Pyrene	0.4 U	0.4 U	0.41 U	0.4 U
Benzothiazole	0.42 U	0.41 U	0.42 U	0.39 U
Metals				
Cadmium	0.41 U	0.43 U	0.41 U	0.41 U
Chromium	1.1	1.7	1 U	1 U
Copper	2.1 U	2.1 U	2.1 U	2 U
Iron	10 U	11 U	10 U	10 U
Lead	1 U	1.1 U	1 U	1 U
Manganese	1.6 U	1.6 U	1.5 U	1.5 U
Nickel	2.1 U	2.1 U	2.1 U	2 U
Silver	1 U	1.1 U	1 U	1 U
Tin	10 U	11 U	10 U	10 U
Zinc	2.1 U	2.1 U	83	2 U

Appendix B-1d
Mullaly Park - Background
New York, NY

* Sample ID :	MPT6	MPT11
Sample Name :	Mullaly Park Turf Background	Mullaly Park Turf Field Background
Sample Date :	08/26/08	08/27/08
* Sample Type :	Air	Air
Lab ID :	280801650-16	280801650-5
Duplicate :		
CONSTITUENTS		
VOCs		
4-Methyl-2-pentanone	2 U	2 U
Acetone	11	10
Acetonitrile	0.84 U	0.84 U
Acrylonitrile	1.1 U	1.1 U
Benzene	1.6 U	1.6 U
Benzyl Chloride	3.7 U	3.7 U
Bromodichloromethane	3.3 U	3.3 U
Bromoethane	2.2 U	2.2 U
Bromoethene	2.2 U	2.2 U
Bromoform	5.2 U	5.2 U
Bromomethane	1.9 U	1.9 U
Butadiene, 1,3-	1.1 U	1.1 U
Butanone, 2- (MEK)	1.5 U	1.5 U
Carbon disulfide	1.6 U	1.6 U
Carbon Tetrachloride	3.1 U	3.1 U
Chlorobenzene	2.3 U	2.3 U
Chloroethane	1.3 U	1.3 U
Chloroform	2.4 U	2.4 U
Chloromethane	1 U	1 U
Chloropropene, 3-	1.6 U	1.6 U
Chlorotoluene, 2-	2.6 U	2.6 U
Cyclohexane	1.7 U	1.7 U
Dibromochloromethane	4.3 U	4.3 U
Dibromoethane, 1,2-	3.8 U	3.8 U
Dichlorobenzene, 1,2-	3 U	3 U
Dichlorobenzene, 1,3-	3 U	3 U
Dichlorobenzene, 1,4-	3 U	3 U
Dichloroethane, 1,1-	2 U	2 U
Dichloroethane, 1,2-	2 U	2 U
Dichloroethene, 1,1-	2 U	2 U
Dichloroethene, cis-1,2-	2 U	2 U
Dichloroethene, trans-1,2-	2 U	2 U
Dichloropropane, 1,2-	2.3 U	2.3 U
Dichloropropene, cis-1,3-	2.3 U	2.3 U
Dichloropropene, trans-1,3-	2.3 U	2.3 U
Dioxane, 1,4-	1.8 U	1.8 U
Ethanol	8.9	6.7
Ethyl acetate	1.8 U	1.8 U
Ethylbenzene	2.2 U	2.2 U
Ethyltoluene, 4-	2.5 U	2.5 U
Freon 11 (Trichlorofluoromethane)	2.8 U	2.8 U
Freon 113 (1,1,2 Trichlorotrifluoroethane)	3.8 U	3.8 U
Freon 114 (1,2 Dichlorotetrafluoroethane)	3.5 U	3.5 U
Freon 12 (Dichlorodifluoromethane)	2.5 U	2.5 U
Heptane, n-	2 U	2 U
Hexachloro-1,3-butadiene	5.3 U	5.3 U
Hexane, n-	1.8 U	1.8 U
Hexanone, 2-	2 U	2 U
Isopropyl alcohol	3.7 U	3.7 U
Methylene chloride	5.2 U	6.9
Methyl-tert-butyl ether	1.8 U	1.8 U
Propylene	1.7 U	1.7 U
Styrene	2.1 U	2.1 U
Tertiary butyl alcohol	1.5 U	1.5 U
Tetrachloroethane, 1,1,2,2-	3.4 U	3.4 U
Tetrachloroethene	3.4 U	3.4 U
Tetrahydrofuran	1.5 U	1.5 U
Toluene	2	1.9 U
Trichlorobenzene, 1,2,4-	3.7 U	3.7 U
Trichloroethane, 1,1,1-	2.7 U	2.7 U
Trichloroethane, 1,1,2-	2.7 U	2.7 U
Trichloroethene	2.7 U	2.7 U
Trimethylbenzene, 1,2,4-	2.5 U	2.5 U
Trimethylbenzene, 1,3,5-	2.5 U	2.5 U
Trimethylpentane, 2,2,4-	2.3 U	2.3 U
Vinyl acetate	1.8 U	1.8 U
Vinyl chloride	1.3 U	1.3 U
Xylene (ortho)	2.2 U	2.2 U
Xylene (para & meta)	2.2 U	2.2 U

Appendix B-1d
Mullaly Park - Background
New York, NY

* Sample ID :	MPT6	MPT11
Sample Name :	Mullaly Park Turf Background	Mullaly Park Turf Field Background
Sample Date :	08/26/08	08/27/08
* Sample Type :	Air	Air
Lab ID :	280801650-16	280801650-5
Duplicate :		
SVOCs		
Acenaphthene	0.81 U	0.8 U
Acenaphthylene	0.81 U	0.8 U
Anthracene	0.81 U	0.8 U
Benzo(a)anthracene	0.4 U	0.4 U
Benzo(a)pyrene	0.4 U	0.4 U
Benzo(b)fluoranthene	0.4 U	0.4 U
Benzo(e)pyrene	0.4 U	0.4 U
Benzo(g,h,i)perylene	0.4 U	0.4 U
Benzo(k)fluoranthene	0.4 U	0.4 U
Chrysene	0.4 U	0.4 U
Dibenzo(a,h)anthracene	0.4 U	0.4 U
Fluoranthene	0.4 U	0.4 U
Fluorene	0.81 U	0.8 U
Indeno(1,2,3-c,d)pyrene	0.4 U	0.4 U
Naphthalene	0.81 U	0.8 U
Phenanthrene	0.4 U	0.4 U
Pyrene	0.4 U	0.4 U
Benzothiazole	0.4 U	0.39 U
Metals		
Cadmium	0.39 U	0.39 U
Chromium	1.8	1.2
Copper	1.9 U	1.9 U
Iron	9.7 U	9.7 U
Lead	0.97 U	0.97 U
Manganese	1.4 U	1.4 U
Nickel	1.9 U	1.9 U
Silver	0.97 U	0.97 U
Tin	9.7 U	9.7 U
Zinc	1.9 U	1.9 U

Appendix B-1e
Thomas Jefferson Park - Background
New York, NY

* Sample ID :	TJ1	TJ14
Sample Name :	Thomas Jefferson Turf Background	Thomas Jefferson Turf Background
Sample Date :	08/25/08	08/28/08
* Sample Type :	Air	Air
Lab ID :	280801650-18	280801650-1
Duplicate :		
CONSTITUENTS		
VOCs		
4-Methyl-2-pentanone	2 U	2 U
Acetone	8.9	9
Acetonitrile	0.84 U	0.84 U
Acrylonitrile	1.1 U	1.1 U
Benzene	1.6 U	1.6 U
Benzyl Chloride	3.7 U	3.7 U
Bromodichloromethane	3.3 U	3.3 U
Bromoethane	2.2 U	2.2 U
Bromoethene	2.2 U	2.2 U
Bromoform	5.2 U	5.2 U
Bromomethane	1.9 U	1.9 U
Butadiene, 1,3-	1.1 U	1.1 U
Butanone, 2- (MEK)	1.5 U	1.5 U
Carbon disulfide	1.6 U	1.6 U
Carbon Tetrachloride	3.1 U	3.1 U
Chlorobenzene	2.3 U	2.3 U
Chloroethane	1.3 U	1.3 U
Chloroform	2.4 U	2.4 U
Chloromethane	1 U	1
Chloropropene, 3-	1.6 U	1.6 U
Chlorotoluene, 2-	2.6 U	2.6 U
Cyclohexane	1.7 U	1.7 U
Dibromochloromethane	4.3 U	4.3 U
Dibromoethane, 1,2-	3.8 U	3.8 U
Dichlorobenzene, 1,2-	3 U	3 U
Dichlorobenzene, 1,3-	3 U	3 U
Dichlorobenzene, 1,4-	3 U	3 U
Dichloroethane, 1,1-	2 U	2 U
Dichloroethane, 1,2-	2 U	2 U
Dichloroethene, 1,1-	2 U	2 U
Dichloroethene, cis-1,2-	2 U	2 U
Dichloroethene, trans-1,2-	2 U	2 U
Dichloropropane, 1,2-	2.3 U	2.3 U
Dichloropropene, cis-1,3-	2.3 U	2.3 U
Dichloropropene, trans-1,3-	2.3 U	2.3 U
Dioxane, 1,4-	1.8 U	1.8 U
Ethanol	6.2	8
Ethyl acetate	1.8 U	1.8 U
Ethylbenzene	2.2 U	2.2 U
Ethyltoluene, 4-	2.5 U	2.5 U
Freon 11 (Trichlorofluoromethane)	2.8 U	2.8 U
Freon 113 (1,1,2 Trichlorotrifluoroethane)	3.8 U	3.8 U
Freon 114 (1,2 Dichlorotetrafluoroethane)	3.5 U	3.5 U
Freon 12 (Dichlorodifluoromethane)	2.5 U	2.5 U
Heptane, n-	2 U	2 U
Hexachloro-1,3-butadiene	5.3 U	5.3 U
Hexane, n-	1.8 U	1.8 U
Hexanone, 2-	2 U	2 U
Isopropyl alcohol	3.7 U	3.7 U
Methylene chloride	5.2 U	5.2 U
Methyl-tert-butyl ether	1.8 U	1.8 U
Propylene	1.7 U	1.7 U
Styrene	2.1 U	2.1 U
Tertiary butyl alcohol	1.5 U	1.5 U
Tetrachloroethane, 1,1,2,2-	3.4 U	3.4 U
Tetrachloroethene	3.4 U	3.4 U
Tetrahydrofuran	1.5 U	1.5 U
Toluene	1.9 U	1.9 U
Trichlorobenzene, 1,2,4-	3.7 U	3.7 U
Trichloroethane, 1,1,1-	2.7 U	2.7 U
Trichloroethane, 1,1,2-	2.7 U	2.7 U
Trichloroethene	2.7 U	2.7 U
Trimethylbenzene, 1,2,4-	2.5 U	2.5 U
Trimethylbenzene, 1,3,5-	2.5 U	2.5 U
Trimethylpentane, 2,2,4-	2.3 U	2.3 U
Vinyl acetate	1.8 U	1.8 U
Vinyl chloride	1.3 U	1.3 U
Xylene (ortho)	2.2 U	2.2 U
Xylene (para & meta)	2.2 U	2.2 U

Appendix B-1e
Thomas Jefferson Park - Background
New York, NY

* Sample ID :	TJ1	TJ14
Sample Name:	Thomas Jefferson Turf Background	Thomas Jefferson Turf Background
Sample Date :	08/25/08	08/28/08
* Sample Type :	Air	Air
Lab ID :	280801650-18	280801650-1
Duplicate :		
VOC TICs		
Nonanal	5.81 U	
SVOCs		
Acenaphthene	0.51 U	0.77 U
Acenaphthylene	0.51 U	0.77 U
Anthracene	0.51 U	0.77 U
Benzo(a)anthracene	0.26 U	0.39 U
Benzo(a)pyrene	0.26 U	0.39 U
Benzo(b)fluoranthene	0.26 U	0.39 U
Benzo(e)pyrene	0.26 U	0.39 U
Benzo(g,h,i)perylene	0.26 U	0.39 U
Benzo(k)fluoranthene	0.26 U	0.39 U
Chrysene	0.26 U	0.39 U
Dibenzo(a,h)anthracene	0.26 U	0.39 U
Fluoranthene	0.26 U	0.39 U
Fluorene	0.51 U	0.77 U
Indeno(1,2,3-c,d)pyrene	0.26 U	0.39 U
Naphthalene	0.51 U	0.77 U
Phenanthrene	0.26 U	0.39 U
Pyrene	0.26 U	0.39 U
Benzothiazole	0.26 U	0.38 U
Metals		
Cadmium	0.33 U	0.38 U
Chromium	0.96	1.1
Copper	1.6 U	1.9 U
Iron	8.1 U	9.5 U
Lead	0.81 U	0.95 U
Manganese	1.2 U	1.4 U
Nickel	1.6 U	1.9 U
Silver	0.81 U	0.95 U
Tin	8.1 U	9.5 U
Zinc	1.6 U	1.9 U

Table B-2
Summary Crumb Rubber Results for Detected Analytes

Analytes	Synthetic Turf Fields (Concentration Range)	Soil Cleanup Objective†
	(mg/kg)	(mg/kg)
Volatile and Semi-Volatile Organic Compounds (VOCs and SVOCs) (Of 77 organic compounds tested, one was detected)		
Naphthalene	0.216	100
Metals (Of eight metals tested, six were detected)		
Arsenic	0.768	16
Barium	0.96 - 4.87	400
Cadmium	0.23 - 1.3	4.3
Chromium	0.888	110
Lead	5.9 - 409*	400
Zinc	1,810 - 13,100*	10,000
*Measurement exceeded NYS DEC soil cleanup objective.		
†NYS DEC, 2006. 6NYCRR Part 375-6-8. Soil Cleanup Objective Tables for restricted residential land uses. http://www.dec.ny.gov/regs/15507.html#15517 .		

Appendix B-2a
Mullaly Park - Bulk Crumb Rubber
New York, NY

* Sample ID * Sample Depth Sample Date * Sample Type Lab ID	Mullaly Park 08/26/08 Crumb Rubber SA-83958-2	Blank 8090301-BLK1
CONSTITUENTS		
VOCs and SVOCs (ug/kg)		
Acetone	1340 U	10 U
Acrylonitrile	67.1 U	1 U
Benzene	134 U	1 U
Bromobenzene	134 U	1 U
Bromochloromethane	134 U	1 U
Bromodichloromethane	67.1 U	1 U
Bromoform	134 U	1 U
Bromomethane	268 U	2 U
Butanone, 2- (MEK)	1340 U	10 U
Butylbenzene, sec-	134 U	1 U
Butylbenzene, tert-	134 U	1 U
Butylbenzene, n-	134 U	1 U
Carbon disulfide	671 U	5 U
Carbon tetrachloride	134 U	1 U
Chlorobenzene	134 U	1 U
Chloroethane	268 U	2 U
Chloroform	134 U	1 U
Chloromethane	268 U	2 U
Chlorotoluene, 2-	134 U	1 U
Chlorotoluene, 4-	134 U	1 U
Dibromo-3-chloropropane, 1,2-	268 U	2 U
Dibromochloromethane	67.1 U	1 U
Dibromoethane, 1,2-	67.1 U	1 U
Dibromomethane	134 U	1 U
Dichlorobenzene, 1,2-	134 U	1 U
Dichlorobenzene, 1,3-	134 U	1 U
Dichlorobenzene, 1,4-	134 U	1 U
Dichlorodifluoromethane (Freon 12)	268 U	2 U
Dichloroethane, 1,2-	134 U	1 U
Dichloroethane, 1,1-	134 U	1 U
Dichloroethene, 1,1-	134 U	1 U
Dichloroethene, cis-1,2-	134 U	1 U
Dichloroethene, trans-1,2-	134 U	1 U
Dichloropropane, 1,2-	134 U	1 U
Dichloropropane, 1,3-	134 U	1 U
Dichloropropane, 2,2-	134 U	1 U
Dichloropropene, 1,1-	134 U	1 U
Dichloropropene, cis-1,3-	67.1 U	1 U
Dichloropropene, trans-1,3-	67.1 U	1 U
Di-isopropyl ether	134 U	1 U
Dioxane, 1,4-	2680 U	20 U
Ethanol	53700 U	400 U
Ethyl ether	134 U	1 U
Ethyl tert-butyl ether	134 U	1 U
Ethylbenzene	134 U	1 U
Hexachlorobutadiene	67.1 U	1 U
Hexanone, 2- (MBK)	1340 U	10 U
Isopropylbenzene	134 U	1 U
Isopropyltoluene, 4-	134 U	1 U
Methyl tert-butyl ether	134 U	1 U
Methyl-2-pentanone, 4- (MIBK)	1340 U	10 U
Methylene Chloride	671 U	10 U
Naphthalene	134 U	1 U
Propylbenzene, n-	134 U	1 U
Styrene	134 U	1 U
Tert-amyl methyl ether	134 U	1 U
Tert-Butanol/butyl alcohol	1340 U	10 U
Tetrachloroethane, 1,1,1,2-	134 U	1 U
Tetrachloroethane, 1,1,2,2-	67.1 U	1 U
Tetrachloroethene	134 U	1 U
Tetrahydrofuran	1340 U	10 U
Toluene	134 U	1 U
trans-1,4-Dichloro-2-butene	671 U	5 U
Trichlorobenzene, 1,2,3-	134 U	1 U
Trichlorobenzene, 1,2,4-	134 U	1 U
Trichlorobenzene, 1,2,5-	134 U	1 U
Trichloroethane, 1,1,1-	134 U	1 U
Trichloroethane, 1,1,2-	134 U	1 U
Trichloroethene	134 U	1 U
Trichlorofluoromethane	134 U	1 U
Trichloropropane, 1,2,3-	134 U	1 U
Trichlorotrifluoroethane, 1,1,2-	134 U	1 U
Trimethylbenzene, 1,2,4-	134 U	1 U
Trimethylbenzene, 1,3,5-	134 U	1 U
Vinyl chloride	134 U	1 U
Xylene, m,p-	268 U	2 U
Xylene, o-	134 U	1 U

Appendix B-2a
Mullaly Park - Bulk Crumb Rubber
New York, NY

* Sample ID :	Mullaly Park	Blank
* Sample Depth :	08/26/08	
* Sample Date :	Crumb Rubber	
* Sample Type :	SA-83958-2	8090301-BLK1
Lab ID :		
Total Metals (mg/kg)		
Arsenic	0.225 U	
Barium	0.956	
Cadmium	0.231	
Chromium	0.15 U	
Lead	5.95	
Selenium	0.225 U	
Silver	0.225 U	
Zinc	1810	

Appendix B-2b
Thomas Jefferson Park - Bulk Crumb Rubber
New York, NY

* Sample ID * Sample Depth Sample Date * Sample Type Lab ID	Thomas Jefferson 0-2' 08/25/08 Crumb Rubber SA-83958-2	Blank 8090301-BLK1
CONSTITUENTS		
VOCs and SVOCs (ug/kg)		
Acetone	2060 U	10 U
Acrylonitrile	206 U	1 U
Benzene	206 U	1 U
Bromobenzene	206 U	1 U
Bromochloromethane	206 U	1 U
Bromodichloromethane	206 U	1 U
Bromoform	206 U	1 U
Bromomethane	412 U	2 U
Butanone, 2- (MEK)	2060 U	10 U
Butylbenzene, sec-	206 U	1 U
Butylbenzene, tert-	206 U	1 U
Butylbenzene, n-	206 U	1 U
Carbon disulfide	1030 U	5 U
Carbon tetrachloride	206 U	1 U
Chlorobenzene	206 U	1 U
Chloroethane	412 U	2 U
Chloroform	206 U	1 U
Chloromethane	412 U	2 U
Chlorotoluene, 2-	206 U	1 U
Chlorotoluene, 4-	206 U	1 U
Dibromo-3-chloropropane, 1,2-	412 U	2 U
Dibromochloromethane	206 U	1 U
Dibromoethane, 1,2-	206 U	1 U
Dibromomethane	206 U	1 U
Dichlorobenzene, 1,2-	206 U	1 U
Dichlorobenzene, 1,3-	206 U	1 U
Dichlorobenzene, 1,4-	206 U	1 U
Dichlorodifluoromethane (Freon 12)	412 U	2 U
Dichloroethane, 1,2-	206 U	1 U
Dichloroethane, 1,1-	206 U	1 U
Dichloroethene, 1,1-	206 U	1 U
Dichloroethene, cis-1,2-	206 U	1 U
Dichloroethene, trans-1,2-	206 U	1 U
Dichloropropane, 1,2-	206 U	1 U
Dichloropropane, 1,3-	206 U	1 U
Dichloropropane, 2,2-	206 U	1 U
Dichloropropene, 1,1-	206 U	1 U
Dichloropropene, cis-1,3-	206 U	1 U
Dichloropropene, trans-1,3-	206 U	1 U
Di-isopropyl ether	206 U	1 U
Dioxane, 1,4-	4120 U	20 U
Ethanol	82500 U	400 U
Ethyl ether	206 U	1 U
Ethyl tert-butyl ether	206 U	1 U
Ethylbenzene	206 U	1 U
Hexachlorobutadiene	206 U	1 U
Hexanone, 2- (MBK)	2060 U	10 U
Isopropylbenzene	206 U	1 U
Isopropyltoluene, 4-	206 U	1 U
Methyl tert-butyl ether	206 U	1 U
Methyl-2-pentanone, 4- (MIBK)	2060 U	10 U
Methylene Chloride	2060 U	10 U
Naphthalene	216	1 U
Propylbenzene, n-	206 U	1 U
Styrene	206 U	1 U
Tert-amyl methyl ether	206 U	1 U
Tert-Butanol/butyl alcohol	2060 U	10 U
Tetrachloroethane, 1,1,1,2-	206 U	1 U
Tetrachloroethane, 1,1,2,2-	206 U	1 U
Tetrachloroethene	206 U	1 U
Tetrahydrofuran	2060 U	10 U
Toluene	206 U	1 U
trans-1,4-Dichloro-2-butene	1030 U	5 U
Trichlorobenzene, 1,2,3-	206 U	1 U
Trichlorobenzene, 1,2,4-	206 U	1 U
Trichlorobenzene, 1,2,5-	206 U	1 U
Trichloroethane, 1,1,1-	206 U	1 U
Trichloroethane, 1,1,2-	206 U	1 U
Trichloroethene	206 U	1 U
Trichlorofluoromethane	206 U	1 U
Trichloropropane, 1,2,3-	206 U	1 U
Trichlorotrifluoroethane, 1,1,2-	206 U	1 U
Trimethylbenzene, 1,2,4-	206 U	1 U
Trimethylbenzene, 1,3,5-	206 U	1 U
Vinyl chloride	206 U	1 U
Xylene, m,p-	412 U	2 U
Xylene, o-	206 U	1 U

Appendix B-2b
Thomas Jefferson Park - Bulk Crumb Rubber
New York, NY

• Sample ID : • Sample Depth : • Sample Date : • Sample Type : Lab ID :	Thomas Jefferson 0-2" 08/25/08 Crumb Rubber SA-83958-2	Blank 8090301-BLK1
Total Metals (mg/kg) Arsenic Barium Cadmium Chromium Lead Selenium Silver Zinc	 0.768 J 4.87 1.3 0.888 J 409 1.5 U 1.5 U 13,100	

Table B-3
Summary of Temperature Measurements

	Ambient Air Temp Range (°F)	Surface Temp Range (°F)
Synthetic Turf Fields		
Mullaly Park Synthetic Turf Field	83 - 87.4	96.7 - 120.8
Thomas Jefferson Park Synthetic Turf Field	79.1 - 84.5	91.9 - 129.1
Background Grass/Upwind		
Mullaly Park Background Grass Field	79.1 - 93.8	87.5 - 110.7
Mullaly Park Upwind Background (Grass)	79.1 - 93.8	91.2 - 110.2
Thomas Jefferson Park Upwind Background (Grass)	79.1 - 84.5	80.5 - 106.6

Temperature readings were obtained with a Kestral 4500 Pocket Weather Tracker every 10 minutes over an approximate 90 minute period.

Table B-3a
 Surface Temperatures
 Mullaly Park - 08/26/08
 New York, NY

Station 4 - Grass Field		Station 5 - Grass Field		Station 6 - Background		Station 7 - Turf Field		Station 8 - Turf - Field	
Time	Temperature °F	Time	Temperature °F	Time	Temperature °F	Time	Temperature °F	Time	Temperature °F
12:40	94.4	12:35	87.5	3:13	104.6	3:10	113.2	3:06	120.3
12:50	107.5	12:52	102.7	3:26	108.5	3:21	120.8	3:24	114.5
1:00	98.6	1:04	97.5	3:37	110.2	3:32	119.9	3:35	115.8
1:10	98.4	1:13	99.8	3:47	98.8	3:43	105.7	3:45	105.5
1:22	96.7	1:25	97.7	3:56	96.8	4:02	107.7	4:04	120.4
1:33	105.2	1:37	102.3	4:14	102.7	4:13	96.7	4:11	113.1
1:42	104.6	1:45	99.8	4:30	106.1	4:27	103.1	4:28	108.1
1:54	101.3	2:01	104.7			4:40	105.3		
2:15	110.7	2:16	100.5						
	101.9		99.2		104.0		109.1		114.0

Table B-3b
 Surface Temperatures
 Mullaly Park - 08/27/08
 New York, NY

Station 9 - Grass Field		Station 10 - Grass Field		Station 11 - Background		Station 12 - Turf Field		Station 13 - Turf - Field	
Time	Temperature °F	Time	Temperature °F	Time	Temperature °F	Time	Temperature °F	Time	Temperature °F
12:15	102.3	12:15	98.3	2:40	105.7	2:35	110.2	2:33	104.2
12:33	103.8	12:35	98.5	3:05	107.2	3:03	112.5	3:01	108.7
12:43	100.2	12:45	98.1	3:25	104.4	3:22	111.8	3:20	109.7
1:00	107.1	1:03	102.9	3:49	101.9	3:46	103.6	3:43	107.4
1:17	108.4	1:19	99.1	4:14	99.7	4:11	101.5	4:08	107.2
1:45	98.9	1:48	92.5	4:22	91.2	4:21	99.8	4:19	100.1
Average	103.5		98.2		101.7		106.6		106.2

Table B-3c
Surface Temperatures
Thomas Jefferson Park - 08/25/08
New York, NY

Station 1 - Background		Station 2 - Turf Field		Station 3 - Turf - Field	
Time	Temperature °F	Time	Temperature °F	Time	Temperature °F
11:20	88.3	11:00	92.9	11:15	104.2
11:34	86.5	11:18	100.8	11:30	97.5
12:47	92.7	11:35	94.6	12:35	114.8
1:07	88.2	12:45	112.3	12:43	112.5
1:20	92.2	1:02	107.5	12:59	110
1:39	91.7	1:17	108.1	1:15	102.3
1:55	88.9	1:35	107.8	1:33	113.4
2:01	106.6	1:52	96.4	1:49	97
2:20	92.2	2:04	106.8	2:07	91.9
		2:20	111.7	2:15	102.8
Average	91.9		103.0		104.8

Table B-3d
 Surface Temperatures
 Thomas Jefferson Park - 08/28/08
 New York, NY

Station 14 - Background		Station 15 - Turf Field		Station 16 - Turf - Field	
Time	Temperature °F	Time	Temperature °F	Time	Temperature °F
12:22	104.3	12:18	122.5	12:16	129.1
12:41	95.6	12:37	100.6	12:39	122
1:10	80.5	1:03	121.8	1:00	121.7
1:25	94.2	1:21	120.9	1:19	112.4
1:50	104.6	1:46	125.6	1:44	124.6
Average	95.8		118.3		122.0

APPENDIX C

SAMPLING WORKSHEETS

Site Name Mulhary Park Date 8-26-08

Type of Field Turf Age of Field 21 yrs

Field Description

Ambient Temperature at start 87.4 ending 83.0

RH start 23.4 ending 26.3

Winds speed start 2.2 wind speed ending 2.0

Wind direction start SW wind direction ending SW

Weather conditions Sunny, clear sky

Data Interpretation

Sample 7 DT R638 14:39 / Sample 8 - R5751

Sample Location ID	Pump #	Start Time	Start Calibration	End Time	End Calibration
MPT6B1	4	2:50	2.017	4:52	2.050
MPT6C	6	2:50	2.051	4:52	2.026
MPT6D	15	2:50	4.068	4:52	4.422
MPT7B1	2	2:45	2.020	4:43	2.081
MPT7C1	1	2:45	2.054	4:43	2.087
MPT7D	7	2:45	4.015	4:43	4.188
MPT8C1	5	2:48	2.078	4:47	1.958
MPT8B1	3	2:48	2.047	4:47	1.969
MPT8D	10	2:48	4.078	4:47	4.110

SW
corner
of
field

center
field

Site Name Mullary Field Park Date 8-27-08

Type of Field Turf Field Age of Field 41

Field Description

Ambient Temperature at start 84.6 ending 83.7

RH start 26.8 ending 30.6

Winds speed start 2.2 wind speed ending 2.1

Wind direction start SE wind direction ending S

Weather conditions Sunny, clear sky

Data Interpretation

Sample Location ID	Pump #	Start Time	Start Calibration	End Time	End Calibration
MPT11B1	5	2:26	2.020	4:33	1.915
MPT11C1	6	2:26	2.032	4:33	2.047
MPT11D	99	2:26	4.035	4:33	4.118
MPT12D	10	2:25	4.073	4:30	4.058 4.033
MPT12B1	2	2:25	2.055	4:30	2.064
MPT12C1	1	2:25	2.034	4:30	2.901
MPT13D	8	2:22	4.066	4:28	4.089
MPT13C1	3	2:22	2.039	4:28	1.923
MPT13B1	4	2:22	2.029	4:28	2.006

DT
R6392

R6038

5751

Site Name Mullary Park Date 8-26-08

Type of Field Grass Baseball Field Age of Field _____

Field Description Baseball field east side of turf soccer fields and west of the train tracks

Ambient Temperature at start 79.1° ending 93.8°

RH start 39.7 ending 26.5

Winds speed start 2.6 mph wind speed ending 1.0

Wind direction start N wind direction ending E

Weather conditions Sunny, clear sky

Data Interpretation _____

Station 4 Dusttrak R6038

Station 5 DT R5751

Sample Location ID	Pump #	Start Time	Start Calibration	End Time	End Calibration
MPF4B1	1	12:22	2.087	2:22	2.068
MPF4C1	2	12:22	2.018	2:22	1.942
MPF4D	7	12:22	4.098	2:22	3.942
MPF5B1	3	12:24	2.045	2:26	2.008 2.006
MPF5C1	5	12:24	2.063	2:26	2.006 1.949
MPF5D	10	12:24	4.036	2:26	3.665

West
out
field

East
out
field

Site Name Mullany Park Date 8-27-08

Type of Field Grass Field Age of Field _____

Field Description

Baseball field next to train tracks. ~ 10 men playing baseball + 4 individuals playing soccer/football around equipment

Ambient Temperature at start 84.2 ending 84.5

RH start 36.5 ending 27.5

Winds speed start 2.8 wind speed ending 3.0

Wind direction start ~~SE~~ E wind direction ending NE

Weather conditions sunny, clear sky

Data Interpretation usual train traffic, construction on S side - New Yankee stadium
no people in the park

Sample Location ID	Pump #	Start Time	Start Calibration	End Time	End Calibration
MPF9B1	1	12:00	2.077	2:01	1.945
MPF9C1	2	12:00	2.008	2:01	2.00 1.946
MPF9D	10	12:00	4.072	2:01	3.941
MPF10C1	4	12:00	2.072	2:05	2.011
MPF10B1	3	12:00	2.084	2:05	1.907
MPF10D	8	12:00	4.072	2:05	3.804

DT
R5751

DT-
R6038

Site Name Thomas Jefferson Date 8-25-08

Type of Field Turf Age of Field 73

Field Description :

~ 10 people played on the field during monitoring, green turf field
no observable odors during testing hours, background sample in the park bt FDR+ Field

Ambient Temperature at start 79.1 ending 84.5

RH start 64.7% ending _____

Winds speed start 0.0 - 1.0 wind speed ending 2.6

Wind direction start NE wind direction ending _____

Weather conditions overcast, cloudy → rain → sunny → overcast

Data Interpretation traffic on the (W), playground on the (E)

Sample Location ID	Pump #	Start Time Start Cali	Start Calibration Stop time Cali	End Time 2nd start cali	End Calibration End-time - Cali
TJ1C	2	10:34 2.010	11:43 2.002	12:29 2.014	2:34 1.965
TJ1B	3	10:34 2.005	11:43 2.001	12:28 2.021	2:34 1.999
TJ1D	9	10:34 4.013	11:43 3.926	—	2:35 3.926
TJ2C	4	10:29 2.043	11:37 2.038	12:20 2.050	2:29 - 1.992
TJ2B	5	10:29 2.000	11:37 1.987	12:19 2.005	2:30 - 1.941
TJ2D	8	10:29 4.044	11:37 4.029	12:22 4.066	2:28 - 4.015
TJ3B	1	10:27 2.108	11:37 2.005	12:15 2.061	2:23 - 1.975
TJ3C	6	10:27 2.053	11:37 2.082	12:17 1.945	2:25 - 1.962
TJ3D	7	10:27 4.072	11:37 4.038	12:18 4.034	2:26 - 3.957
TJ1D	99	12:27 4.020	1:49 3.915	—	2:38 - 3.915

Site Name Thomas Jefferson field Date 8-28-08

Type of Field Turf field Age of Field 73

Field Description

Ambient Temperature at start 81.1 ending 81.9

RH start 42.5 ending 20.7

Winds speed start 1.8 wind speed ending 2.7

Wind direction start E wind direction ending E

Weather conditions Sunny, clear sky

Data Interpretation Kids play on w side of field

Sample Location ID	Pump #	Start Time	Start Calibration	End Time	End Calibration
TJ14B1	5	11:40	4.029	1:54	1.848
TJ14C1	6	11:40	2.009	1:54	1.964
TJ14D	10	11:40	2.025	1:54	3.839
TJ15C1	1	11:48	2.070	2:00	2.016
TJ15B1	2	11:48	2.052	2:00	1.977
TJ15D	7	11:48	4.061	2:00	4.232
TJ16B1	3	11:55	2.053	2:05	1.941
TJ16C1	4	11:55	2.053	2:05	1.954
TJ16D	8	11:55	4.065	2:05	3.872

Tur 79°

DT
6392

T123
DT
6038

5757